Non-Specific Membrane-Matrix Interactions Influence Diffusivity of Lipid Vesicles in Hydrogel Environments

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Abstract: The diffusion of extracellular vesicles and liposomes in vivo is affected by different tissue environmental conditions and is of great interest in the development of liposome-based therapeutics and drug-delivery systems. Here, we use a bottom-up biomimetic approach to study how steric and electrostatic interactions influence the diffusivity of synthetic large unilamellar vesicles in hydrogel environments. Single-particle tracking of these extracellular vesicle-like particles in agarose hydrogels as an extracellular matrix model shows that membrane deformability and surface charge affect the hydrogel pore spaces that vesicles have access to, which determines overall diffusivity. Moreover, we show that passivation of vesicles with PEGylated lipids, as often used in drug delivery systems enhances diffusivity, but that this effect cannot be fully explained with electrostatic interactions alone. Finally, we compare our experimental findings with existing computational and theoretical work in the field to help explain the non-specific interactions between diffusing particles and gel matrix environments.

Key words: **Extracellular vesicles, liposomes, hydrogels, PEGylation, diffusion, agarose**

Large unilamellar vesicles (LUVs), or liposomes, are phospholipid structures 100-1000 nm in diameter that are often used as a minimal model of cell-derived extracellular vesicles (EVs). Their application in drug delivery systems takes advantage of the structure and function of their in vivo counterparts, as facilitators of intercellular transport[1,2] to shield their payloads from the external tissue environment and mediate their transport to and uptake by target cells.[3,4] Despite the rising interest in using such lipid nanoparticle systems, for example to deliver anti-cancer therapeutics[5,6] or as carriers of immunogenic materials in vaccine formulations,[7] there is little data on how nanoparticle mobility and transport within tissues is affected by different tissue environmental conditions and the membrane material properties.

A 2009 study by Lieleg et al.[8] showed that hydrogel materials derived from extracellular matrix (ECM) can act as an electrostatic filter, sequestering charged nanoparticles while allowing neutral particles to pass through unimpeded. Later, Yu et al.[9] and Lenzini et al.[10] found that the deformability of lipid vesicles, modulated by the lipid composition of the membrane and by the presence of water-permissive channel proteins, respectively, can influence their access through hydrogel pore spaces and thus their movement and transport. Other work, including studies on rigid polymeric nanoparticles diffusing in polymer solutions,[11,12] in colloidal mucin suspensions,[13] and in hydrogels[14–17] have also investigated the various ways in which charge and steric interactions affect the dynamics of non-deforming particles. Clearly, particle diffusion is affected by a diverse range of biophysical factors and of particular interest is the way these factors might interact. For example, recent theoretical work suggests that particle diffusability is a balancing act between particle deformability and particle-matrix adhesion.[18] When particles or vesicles are subjected to surface modifications, such as PEGylation[8,13,19–23] or the inclusion of more complex molecules, their surface interactions with the surrounding medium must also be considered. It is not difficult to imagine, then, that the combined effects of such interactions can give rise the complex distribution patterns of vesicles observed in vivo.[24]

To systematically study how different material properties of lipid vesicles and ECM-like hydrogels can influence vesicle diffusion, we use single-particle tracking[25] to study the diffusion of synthetic LUVs embedded in agarose. Agarose is a polysaccharide polymer from red algae that undergoes thermo-reversible gelation *via* non-covalent hydrogen bonding.[26,27] The polymer contains negatively charged sulfate and pyruvate residues and its chain length can vary, thus affecting the gel mechanics.[26,28] While much simpler in chemical composition than the diverse molecules found in human ECM, agarose provides greater control over material properties. Stiffness and porosity of agarose gels, for example, can be found in a comparable range to human tissues such as brain or cartilage,

and can be controlled with concentration and gelation conditions.[29,30] Agarose is also a relevant material used in a number of different biomedical applications,[29,31,32] including in three-dimensional cell culture platforms[33,34] and as components of composite materials for tissue engineering.[29,35–37] Most importantly, agarose is bio-inert,[33] allowing for the investigation of non-specific steric and electrostatic interactions without the influence of specific biochemical interactions that may be present with reconstituted ECM materials or mucin suspensions. By using this biomimetic system, we aim to better understand the biophysical mechanisms that govern the diffusion of extracellular-like vesicles or drug carriers in tissue-like materials. We also aim to directly compare deformable vesicles with similarly-sized rigid nanoparticles and particles with surface modifications to better tease apart how particle deformability and surface interactions contribute to overall particle diffusion and dynamics. Altogether, these results could one day lead to more efficient targeting and delivery of lipid nanoparticle-based therapeutics and vaccine delivery.[3,5–7,38]

Results and Discussion

Mobility of embedded LUVs. LUVs embedded inside agarose gels were imaged with epifluorescence microscopy and analyzed with single-particle tracking[25] to obtain their diffusion coefficients (Fig. 1A,B). The mobile fraction was determined from the distribution of $\log_{10}$ diffusion coefficients. The peak observed at a value of -15 for 100 nm-extruded LUVs composed of pure 1,2-dioleoyl-sn-glycero-3-phosphocholine (DOPC) embedded in 1% agarose (Fig. 1B) lies below the mobility cut-off of -14 (see Supplemental Information, SI Section S1) and thus corresponds to fully-immobilized particles. Particle immobilization and interactions in general with a gel matrix have previously been described in terms of electrostatic effects (at least for polystyrene particles).[39] In essence, while some particles can become fully entrapped by the gel matrix, other particles will be able to diffuse unhindered within the matrix voids as if they were in liquid water. Given close enough proximity to a wall or surface, particles can transiently bind and unbind with the gel matrix, reducing their mean squared displacement and thus "effective" diffusion coefficient. It is possible in our case with flexible lipid vesicles that these interactions can also be steric, with transient trapping and freeing of particles due to thermal fluctuations. The obtained effective diffusion coefficient can thus be used as a measure of the frequency and strength of membrane-matrix interactions, including steric ones. Since virtually all particles lie below -12 (Fig. 1B), the value given by the Stokes-Einstein equation for an ideal 100nm spherical particle diffusing in liquid water, this implies that all particles in the system are interacting with the gel matrix, sterically or otherwise.

Analyzing individual particle trajectories reveals different diffusive behaviours. Freely diffusing particles in liquid can be observed covering large areas while immobilized particles in gels remain stationary. Most LUVs embedded in gels undergo anomalous diffusion or subdiffusive behaviour, whereby particles diffuse within the confines of a matrix pore, resulting in a characteristic mean squared displacement (MSD) curve with a plateau at long time lags (Fig. 1E). These particles can sometimes "hop" between pore spaces, similar to what has been described in polystyrene nanoparticles diffusing in liquid polymer solutions[12] and in hydrogel matrices.[14] Examples can be seen in Figs. 1 and 2. We note that particles appear to also become transiently trapped at certain locations within presumed pore spaces. These particles dwell at these locations for multiple consecutive image frames for periods of 0.25 s, up to several seconds long, sometimes alternating between several trapping points before being freed. This appears to occur at size scales smaller than the apparent pore sizes mapped out by the rest of the particle's trajectory or by neighbouring particles.

Collective LUV diffusion from gel infiltration. We next looked at the ability of DOPC LUVs to infiltrate pre-formed agarose hydrogel disks to obtain an independent measure of particle mobility based on population dynamics (Fig. 1C, SI Section S2). While many LUVs ended up adhering to and getting stuck at the edge of the agarose disks, some LUVs were observed to infiltrate into them over 100 hours of imaging. Equilibrium in the density gradient was not reached in the time frame tested. By approximating the LUV density gradient with a finite difference, we calculate a population-wide diffusion coefficient using Fick's First Law of Diffusion to be $3.92 \pm 0.52 \times 10^{-14}\ m^2 s^{-1}$ (see Fig. 1D and SI Section S2), corresponding to a $\log_{10}$ value of -13.41. This falls between the mobility cut-off and the value for the Stokes-Einstein particle, agreeing well with the results from the single-particle tracking of embedded LUVs (Fig. 1A, B).

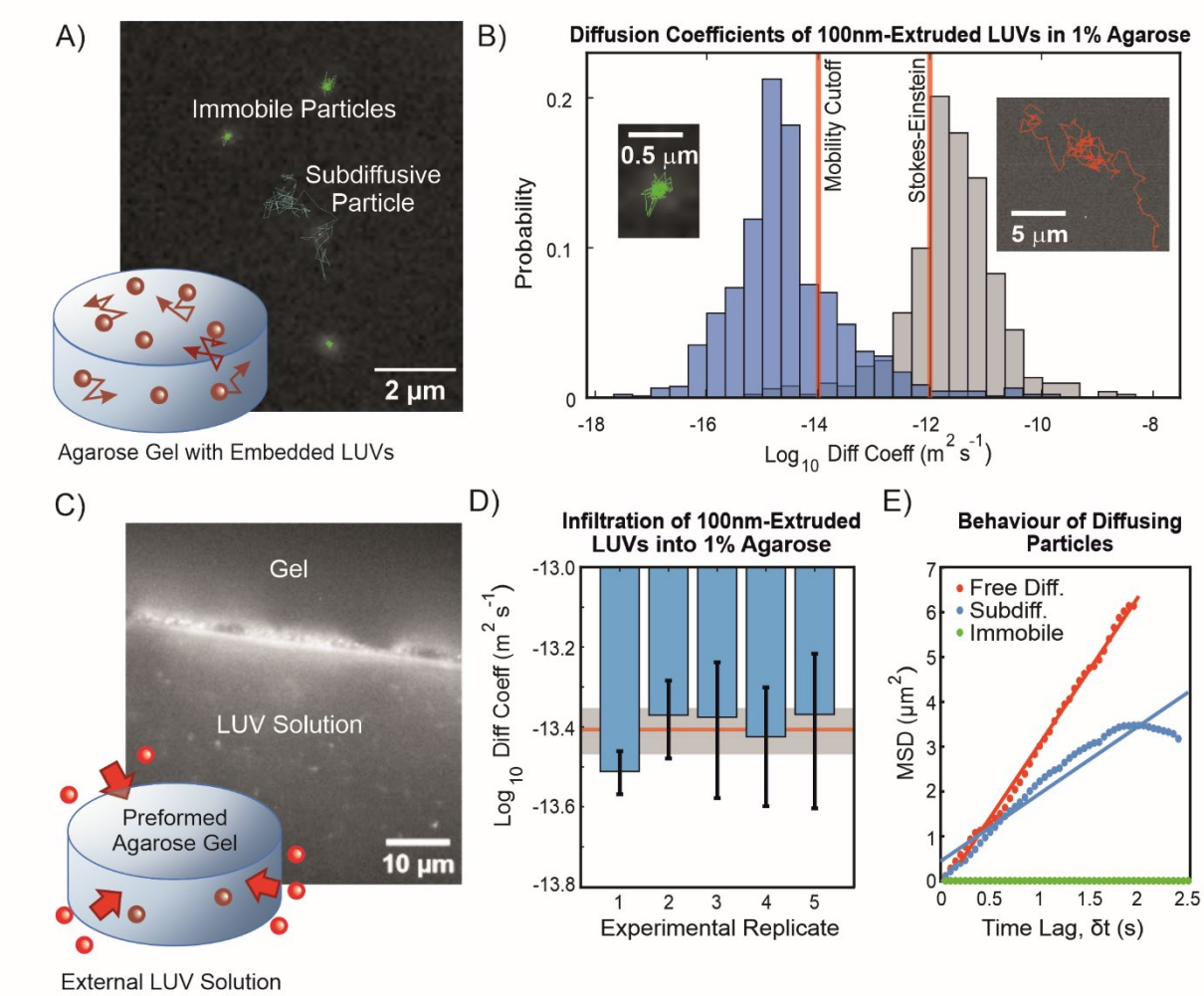

Figure 1, Particle mobility assessed in two ways: diffusion of DOPC LUVs embedded in gels and infiltration of LUVs into gels. A) The diffusion of 100nm-extruded DOPC LUVs embedded in 1% agarose gels is analyzed with single particle tracking. Particle paths are indicated as an overlay on a representative time frame. Here, green trajectories represent immobilized LUVs and a "subdiffusive" trajectory is shown in blue. This particular LUV was observed to "jump" between two hydrogel pore regions, resulting in this biphasic trajectory separated by a particularly large displacement. B) The base-ten logarithms of the diffusion coefficients are shown in histogram form for 100nm-extruded LUVs embedded in 1% agarose (blue) and in agarose-free liquid PBS (grey). The values of -14 and -12 are indicated by red vertical lines, the former being the lower limit of detection of particle motion for the experimental setup and the latter being the theoretical value determined from the Stokes-Einstein equation for an ideal 100nm-diameter spherical particle diffusing in liquid water. The insets show fragments of two trajectories corresponding to an immobilized/confined LUV in 1% agarose (green, left) and a freely diffusing one in agarose-free PBS (red, right). C) A measure of the collective particle diffusion is determined by incubating preformed agarose gel disks with an external solution of LUVs and monitoring their infiltration into the gel. The epifluorescence microscopy image shows the edge of an agarose gel disk, where LUVs can be seen adsorbed onto the surface. Individual LUVs in the solution are seen as tiny spots. D) The particle densities inside and outside the gel over time can be related to the diffusion coefficient (see SI Section S2). The average $\log_{10}$ diffusion coefficients computed from five independent experimental replicates are presented here with error bars showing standard deviation. The ensemble average across all experimental replicates was found to be $3.92 \pm 0.52 \times 10\text{-}14 m^2 s^{-1}$ (indicated as a red line with grey shaded area showing standard deviation), corresponding to a $\log_{10}$ value of -13.41. E) In both experiments, LUVs can be observed to exhibit different diffusive behaviours, characterized by different mean squared displacement (MSD) plots. A freely-diffusing LUV is characterized by a linear MSD plot, as shown in red, where the slope is proportional to the diffusion coefficient. The green plot shows a fully immobilized particle in agarose, where the slope is very close to zero. The blue MSD plot shows characteristic subdiffusive or anomalous diffusion behavior, whereby the initial portion at small time lag is relatively linear, but reaches a plateau at longer time lags due to the particle being confined within hydrogel pores. Straight line fits are shown for each MSD data set to show how the "effective" diffusion coefficient would be calculated.

Osmotic deflation of LUVs increases diffusivity. Work by Yu et al.[9] on LUV compositions of different phase transition temperature and by Lenzini et al.[10] on cell-derived EVs has shown that vesicle deformability can affect their diffusion in a hydrogel. Another way to make LUVs more deformable is to deflate them by introducing them into a hypertonic environment. We studied the mobility of DOPC LUVs in agarose gels of differing osmolalities by adding glucose to the hydrogel solution while keeping the initial intravesicular solution constant. Fig. 2A shows that LUV mobility increases with osmolality from isotonic to +12% osmolality, and thus degree of deflation. No significant difference in LUV size was detected with dynamic light scattering (DLS), though size distributions appear to have slightly higher variability in hypertonic solutions (Fig. 3B in the main text; size distributions found in Fig. S3C in the SI). We also did not observe differences in the bulk rheology of agarose gels formed with and without glucose (Fig. S3A), thus the microstructure of the gel is not expected to vary more than what is naturally found in agarose.[26] The lack of statistically significant difference in mobility from +12% to +17% osmolality (Fig. 2A) could be due to a phenomenon similar to what was described by Yu et al.[9] whereby greater deformability ultimately exposes greater surface area that can conform to and interact with the matrix walls, resulting in immobilization. Recent theoretical work[18] also shows that particle diffusibility in a gel matrix is dependent on a balance of particle deformation and adhesive forces in the matrix. Kinetic energy in a hyper-deformable vesicle's collision with a matrix wall could end up being spent on deforming the membrane, such that insufficient energy remains for overcoming matrix-adhesive forces.

When the pore size of the gel is increased from 100±40 nm to 650±450 nm by decreasing the agarose concentration from 1% to 0.5% w/v (Fig. S4), the overall mobility of LUVs increases (Fig. 2D). At this concentration of agarose, the difference between the mobile fractions in different osmolalities is not statistically significant. The effect of deflation thus appears to only be relevant when the average pore diameter is comparable (*i.e.*, on the same order of magnitude) to the diameter of the LUV. This is reasonable, as the LUVs would have greater access to matrix pores regardless of deformability in the 0.5% gel.

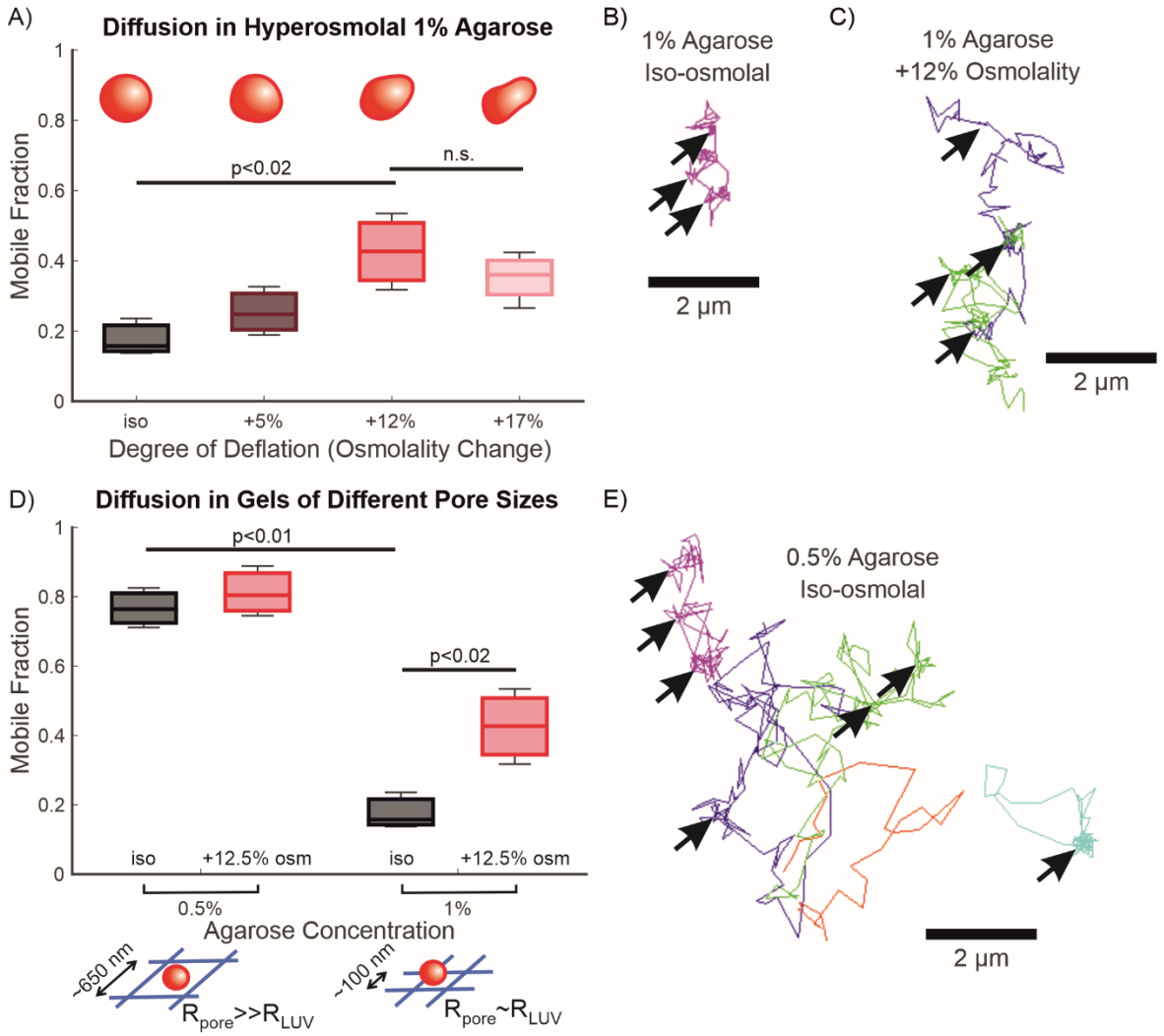

Figure 2, Mobility of DOPC LUVs increases with osmotic deflation and with hydrogel pore size. A) Mobile fractions of LUVs in agarose gels with different osmolalities relative to the LUV interior with cartoon representations of a LUV at varying stages of deflation. B) Example trace of a particle trajectory in 1% iso-osmolal agarose with arrows indicating apparent particle "trapping" regions. This trajectory may not be representative of the majority of particles in this condition, as it is specifically a particle with relatively high mobility to illustrate trapping behaviour, hence the apparent pore size traced by the particle trajectory may not correspond to the average size determined with turbidimetry. C) Representative particle trajectories in 1% agarose with 12% osmolality increase compared to the intravesicular solution. Different colours represent different particles and arrows indicate apparent trapping points. D) Mobile fractions of LUVs in agarose gels of differing concentration and osmolal strength with cartoon representations comparing the relative sizes of pores and LUVs. Bars on graphs represent statistically significant differences, as determined with one-way ANOVA and pairwise Tukey-Kramer post-hoc analysis at the p-values indicated. Lack of statistically significant difference is indicated with n.s. E) Representative particle trajectories in 0.5% iso-osmolal agarose with arrows indicating apparent trapping points. Different colours represent separate particles.

LUV surface charge affects mobility. Lieleg et al.[8] reported that Matrigel, a complex mixture of cell-derived ECM materials exhibits electrostatic filtering behavior on diffusing particles. This has also been shown with polymer solutions and in computer simulations.[11] To determine whether agarose has similar characteristics, we produced negatively charged LUVs from a 2:1 molar ratio mixture of DOPC/DOPS (1,2-dioleoyl-sn-glycero-3-phospho-L-serine) as well as positively charged LUVs composed of 2:1 DOPC/DOTAP (1,2-dioleoyl-3-trimethylammonium-propane) to embed in agarose (Fig. 3A). The addition of DOPS does not significantly change LUV size (Fig. 3B), but the DOPC/DOTAP particles appear significantly larger than other tested particles, possibly due to aggregation. Figure 3C shows that the negatively charged DOPC/DOPS LUVs have greater mobility overall compared to the positively charged DOPC/DOTAP and pure zwitterionic DOPC LUVs (see also Fig. S5). Increased osmolality (i.e. LUV deflation) does not significantly improve the mobility of DOPC/DOPS LUVs, but there still appears to be an upward trend. DOPC/DOTAP LUVs remain immobile in hypertonic conditions, as well as in much lower agarose concentrations (Fig. 4A-C,E). At 0.2% w/v, agarose behaves like a liquid, being able to flow. The lack of mobility at this concentration would suggest that this interaction is not merely steric, but electrostatic, causing the positively charged LUVs to stick to the agarose polymer bundles.

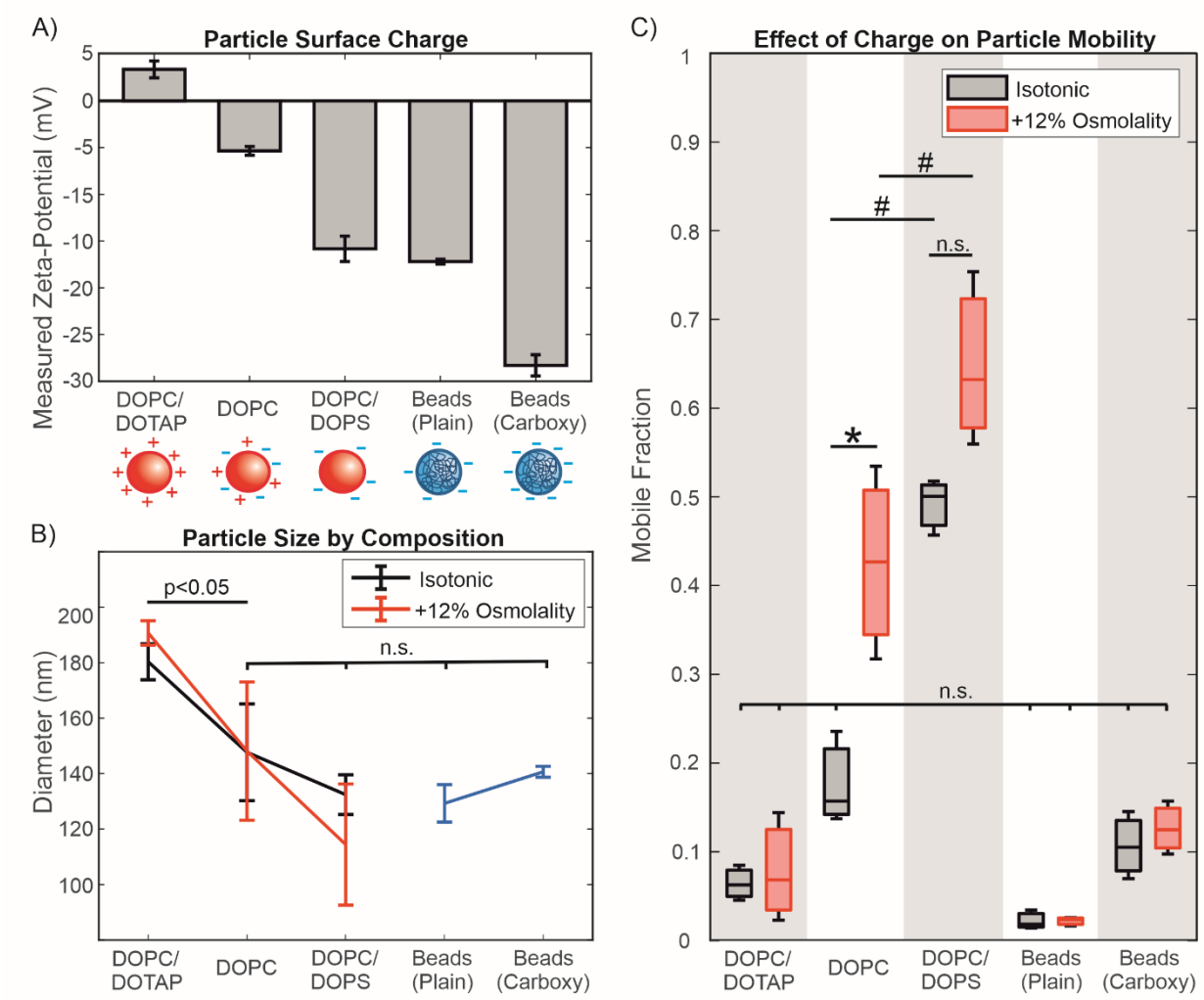

Figure 3, Effect of charge and composition on particle mobility in 1% agarose gels. The lipid ratios in DOPC/DOTAP and DOPC/DOPS LUVs correspond to 2:1. A) Zeta-potential of LUVs and polystyrene beads (either plain or with surface carboxylation), as measured in PBS; the sketches below roughly illustrate their surface charge. B) Diameters of LUVs and particles, as measured with dynamic light scattering. Average diameters of DOPC/DOPS and DOPC LUVs, as well as the polystyrene beads (blue) are not significantly different, in both isotonic (black) and hypertonic (red) conditions, as determined by 2-way ANOVA with pairwise Tukey-Kramer post-hoc analysis ($p>0.05$). Representative size distributions of particles can be found in Fig. S5 in the SI. Sizes of DOPC/DOTAP LUVs are significantly larger than those of the other particles ($p<0.05$). C) Mobile fractions of different particles in isotonic (black; 290mOsm/kg) and +12% hypertonic (red; 320mOsm/kg) buffer conditions; see Fig. S5 for diffusivity data. Statistical significance was determined with 2-way ANOVA with pairwise Tukey-Kramer post-hoc analysis. # represents a significant difference across membrane compositions ($p<0.01$). Statistically significant differences ($p<0.02$) at different osmolalities is represented with * while lack of a significant difference ($p>0.02$) is denoted with n.s.

As a control, we compared LUV mobility to that of polystyrene beads. Despite having a similar size and negative surface charge to our DOPC/DOPS LUVs, plain polystyrene beads are fully immobilized in the hydrogel compared to the highly mobile DOPC/DOPS LUVs. Beads with surface carboxylation appear to have slightly higher mobility than plain beads, but not at a statistically significant level. Neither types of beads are affected by increased osmolarity. One explanation for this relates to the fact that the polystyrene beads are rigid while the LUVs are deformable and capable of squeezing through gel matrix pores that would otherwise be too small to pass through. The increased surface charge of the DOPC/DOPS compared to DOPC LUVs should also result in a slightly stiffer membrane,[40] although this effect could be minimized by the high salt concentration. In 0.5% agarose gels, plain polystyrene beads remain immobile (mobile fraction = 0.07 ± 0.05) despite the much larger pore size while carboxylated beads become much more mobile (mobile fraction = 0.87 ± 0.08; see Fig. S5 in SI). It is possible that the enhanced negative charge of the carboxylated beads overcomes specific attractive interactions present between the agarose gel and the polystyrene beads. Another possibility relates back to the theoretical work of Yu et al.,[18] who showed that highly rigid particles require very low attractive forces to maintain diffusibility. These results reveal an interesting intersection of different factors affecting the diffusion of particles through a gel matrix. In particular, we note that the use of polystyrene nanoparticles to model the diffusion of LUVs, EVs, and other soft particles may not be accurate due to the differences in their overall deformability, even if their surface properties are matched.

PEGylation of LUVs increases their mobility. Nanoparticles are often "passivated" with PEG, a hydrophilic polymer used to prevent the adsorption of proteins on surfaces and hinder or slow down immune reactivity. It has been claimed that this passivation effect is due to the neutral charge of PEG masking the underlying surface,[8,19,22] though other works suggest the involvement of steric or entropic effects of the PEG chains.[13,41] We questioned whether this effect could restore the mobility of our DOPC/DOTAP LUVs (Fig. 4D). For this to work, the layer of PEG chains would need to be thicker than the Debye length of the charges on the membrane. The inclusion DSPE-mPEG1K ((1,2-distearoyl-sn-glycero-3-phosphoethanolamine-N-[methoxy(polyethylene glycol)-1000]), a phospholipid coupled to a 1000 Da PEG chain at 10 mol% should make a PEG layer thick enough to screen out most electrostatic effects. In the high ionic strength buffer environment (PBS) that was tested, the Debye length, the distance over which an electric charge exerts an influence is <1 nm.[42] Meanwhile, the three-dimensional Flory radius of the polymer, R_F is given by

$$R_F \approx a_m n_p^{3/5}$$

where a_m is the size of the monomer unit ($a_m \approx 0.39$nm for PEG, as used by Marsh et al.[43]) and n_p is the number of monomers in the polymer (~23 for PEG1000). The mean-field theory equilibrium length of the polymer chain, L^{MF}, describing the average height of the polymer brush layer is also given by Marsh et al. as

$$L^{MF} \approx n_p a_m^{5/3} \left(X_p / A_l \right)^{1/3}$$

where X_p is the molar fraction of PEGylated lipid and A_l is the area per lipid molecule of the membrane, taken to be ~0.6nm^2 for a lipid in the fluid phase. Both R_F and L^{MF} are approximately 2.6 nm for PEG1000 at 10mol% coverage in a fluid-phase lipid membrane, and thus greater than the Debye length. Therefore, the PEG layer should be thick enough to block electrostatic interactions with the underlying phospholipid surface. This indeed appears to restore the mobility of the DOPC/DOTAP LUVs in 0.5% w/v agarose to the same level as pure DOPC LUVs (and DOPC+PEG LUVs; Fig. 4F). The mobility of the passivated DOPC/DOTAP+PEG LUVs in 1% agarose is much improved compared to that of the non-passivated DOPC and DOPC/DOTAP LUVs, but is lower than the passivated DOPC+PEG LUVs. One possible explanation for this is that LUVs need to deform to fit through the smaller pores of the 1% agarose. This could force the PEG layer to compress or deform out of the way, exposing the underlying positive charge. This is possible, as previous work has shown that PEG is highly compliant.[20,23]

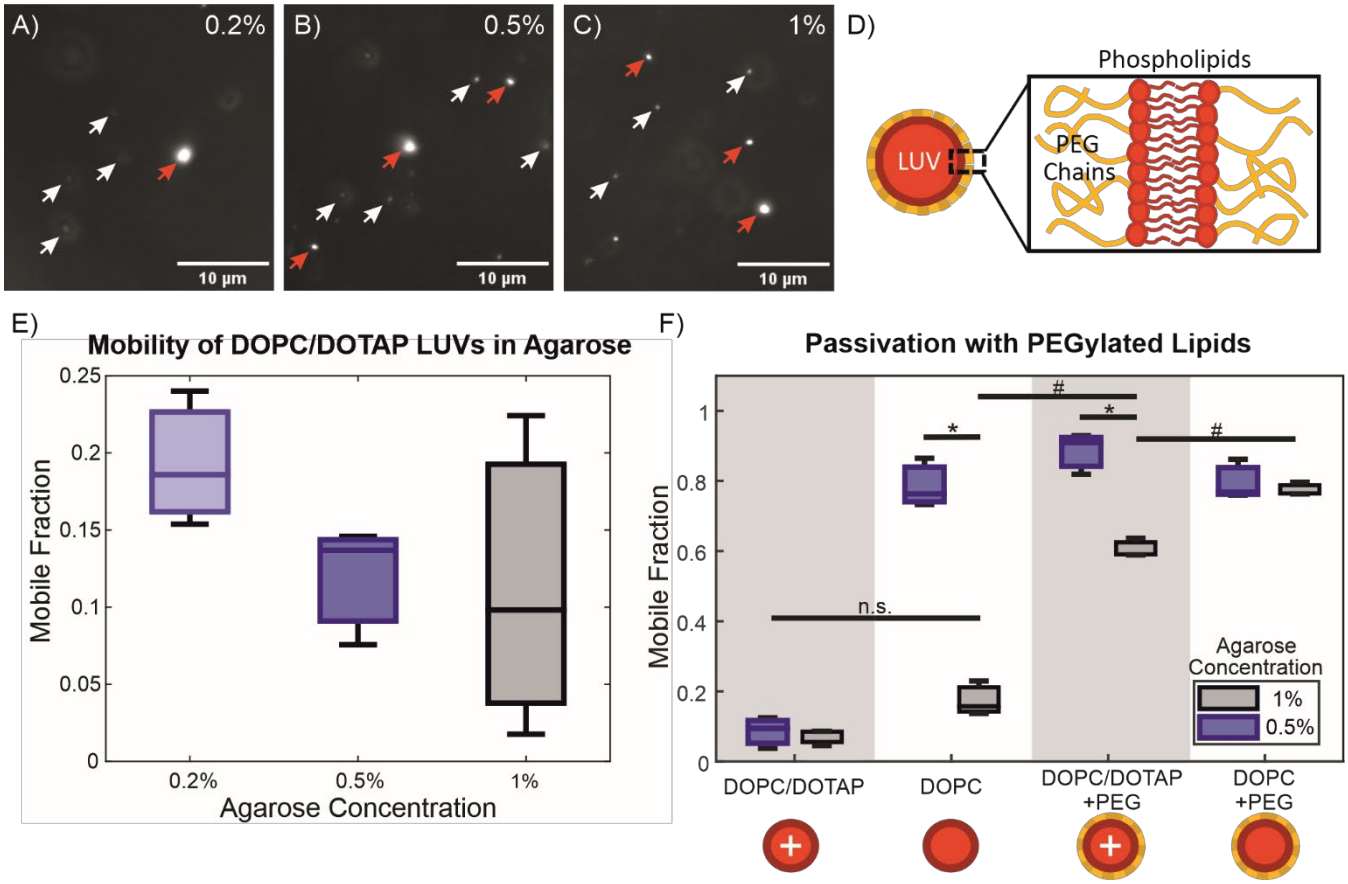

Figure 4, Recovery of DOPC/DOTAP LUV mobility with PEGylated lipids. A-C) Representative images of 100 nm-extruded 2:1 DOPC/DOTAP LUVs in 0.2%, 0.5%, and 1% w/v agarose, respectively. Red arrows indicate abnormal particles or possible particle aggregates while white arrows indicate single particles for comparison. D) Schematic diagram of a LUV membrane containing PEGylated phospholipids. E) Mobile fractions of 2:1 DOPC/DOTAP LUVs in agarose gels/solutions of differing concentration. Agarose at 0.2% concentration behaves macroscopically as a liquid, but likely has some weak molecular network linkages. F) Comparison of mobile fractions of 2:1 DOPC/DOTAP LUVs (DOPC/DOTAP), pure DOPC LUVs (DOPC), 2:1 DOPC/DOTAP LUVs +10mol% DSPE-mPEG1K (DOPC/DOTAP + PEG), and DOPC LUVs +10mol% DSPE-mPEG1K (DOPC + PEG). # indicates statistically significant differences when comparing mobilities across LUV compositions in the same gel concentration. * indicates statistically significant differences when comparing mobilities of LUVs in different agarose gel concentrations. Mobility of DOPC/DOTAP LUVs in both 1% and 0.5% agarose, as well as that of DOPC LUVs in 1% agarose are not significantly different (n.s.). Statistical significance is determined with two-way ANOVA with pairwise Tukey-Kramer post-hoc analysis for multiple comparisons.

While PEGylation improved the mobility of DOPC/DOTAP LUVs, it also improved that of DOPC LUVs despite previous results of higher mobility in agarose with greater negative charge. To investigate further how PEG affects LUV mobility, we tested other PEG chain sizes (1000, 2000, 5000 Da) at two concentrations (1 mol% and 10 mol%; Fig. 5A-C). Figure 5D,E shows that all PEG chain sizes improve LUV mobility in 1% agarose, while LUVs with PEG2000 and PEG5000 appear to be more mobile than those with PEG1000 in 0.5% agarose gels (see histograms and size distributions in Fig. S6, SI). LUV mobility is higher with 10mol% PEGylated lipid, where the PEG chains are in a polymer brush conformation[20,23,43] compared to 1 mol% polymer, where they are in mushroom conformation.[20,23,43] The difference in PEG chain conformations could help explain why PEG1000 appears ineffective in 0.5% agarose, as the mushroom-to-brush transition occurs at a higher concentration and is less well-defined. These results suggest that the effect that PEG has on particle mobility is not entirely electrostatic, as previously claimed.[8]

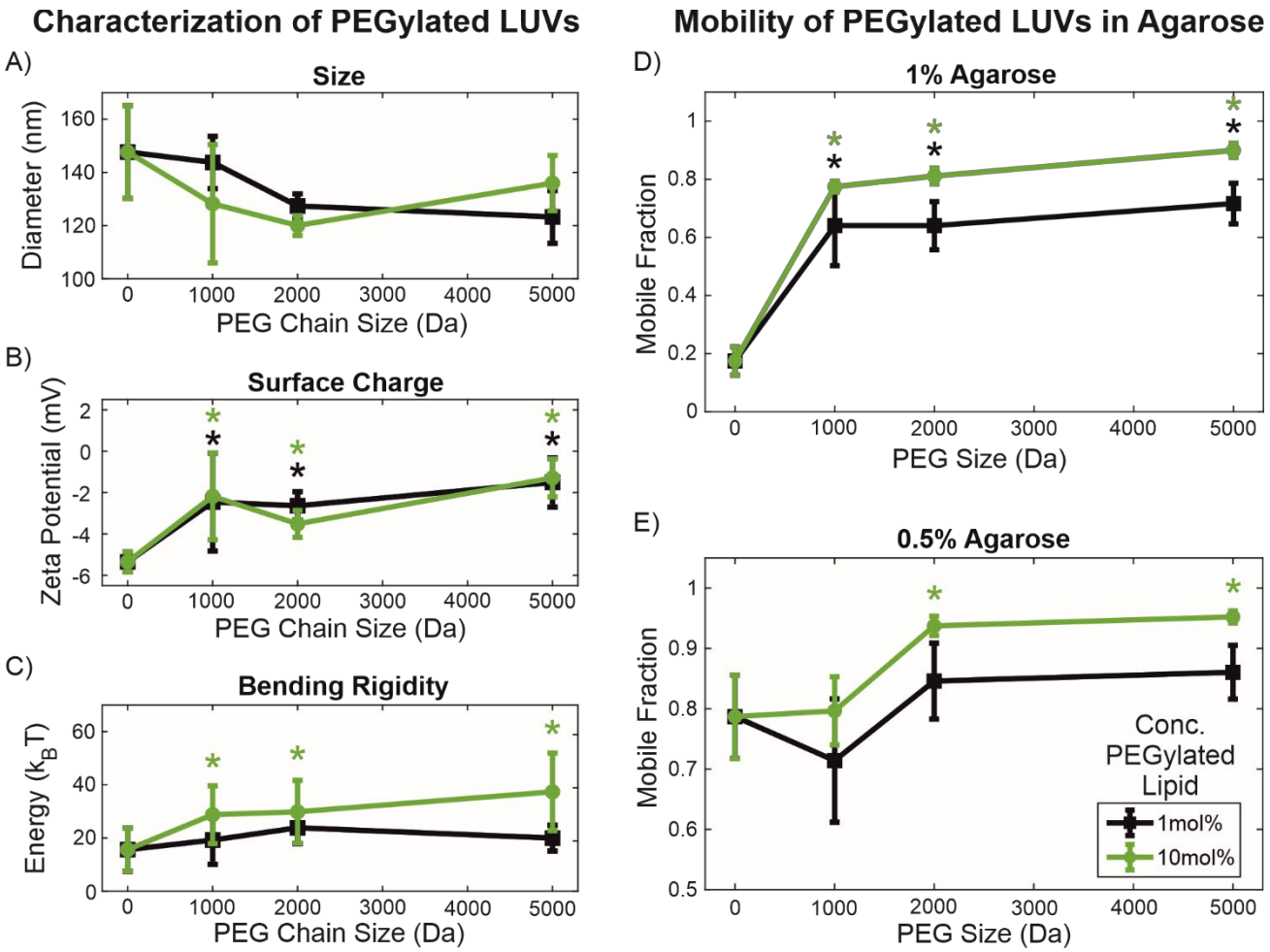

Figure 5, Effect of degree and density of PEGylation on LUV characteristics and mobility. Green data represents 10 mol% concentration of PEGylated lipid in membranes while black data represents 1 mol% concentration. A) Influence of PEG chain size and concentration on LUV size, as measured with DLS. No statistically significant differences were found ($p>0.05$) and size distributions appear similar (Fig. S6B in SI). B) Influence of PEG chain size and concentration on surface charge, measured as the zeta-potential. Bare DOPC LUVs are significantly more electronegative compared to PEGylated LUVs, regardless of PEG size and concentration ($p<0.01$, green and black *), demonstrating screening by PEG. C) Influence of PEG chain size and concentration on membrane bending rigidity, measured with fluctuation analysis on giant unilamellar vesicles (GUVs) and presented in k_BT energy units (k_B being the Boltzmann constant and T being temperature). Membranes with 10 mol% PEGylated lipid are significantly stiffer than those with 1 mol% PEGylated lipid. Bare DOPC membranes are significantly less stiff than PEGylated membranes only at 10 mol% concentration ($p<0.01$). No statistically significant differences were found between different PEG chain sizes. Statistically significant differences compared to bare DOPC membranes were determined with 2-way ANOVA with pairwise Tukey-Kramer post-hoc analysis, as indicated with *. D) Mobilities of all PEGylated particles in 1% agarose are significantly greater than those of bare DOPC LUVs at both concentrations of PEGylated lipid ($p<0.01$, green and black *). Mobilities of LUVs with 10 mol% PEGylated lipid are significantly greater than those of LUVs with 1 mol% PEGylated lipid ($p<0.01$). E) Mobilities of LUVs with 10mol% PEG2000 and PEG5000 are significantly greater than those of bare LUVs and LUVs with 10mol% PEG1000 ($p<0.01$). No significant differences were found at 1 mol% PEGylated lipid. Statistically significant differences compared to bare DOPC membranes are indicated with *, as determined by 3-way ANOVA with pair-wise Tukey-Kramer post-hoc analysis.

Fluctuation analysis on GUVs composed of the same lipid mixtures as the tested LUVs shows a slight stiffening of lipid membranes with the presence of PEGylated lipids (Fig. 5C). This behavior is consistent with predictions for membrane stiffening by anchored polymers,[44] and experiments on membranes with biopolymer adsorption[45] as well as microemulsions.[46] However, the stiffening observed here appears too small to fully explain the dramatic increase in mobility in 1% agarose for both concentrations of PEGylated lipid tested. One possibility would be that the PEG forms a soft lubricating layer that facilitates the movement of LUVs through matrix pores. This would be supported by the higher mobility observed with larger PEG chains and at higher PEG coverage (in the brush regime), which would form a thicker layer. The underlying mechanism for this could be explained by the entropic repulsive force described in the computational work of Li and Shi,[41] whereby the compression of the PEG layer during a collision with the matrix wall would produce a strong repulsive force, preventing entrapment of the LUV in the matrix. It has also previously been reported that surface PEGylation improves the mobility of nanoparticles in mucin gels by sterically preventing the adsorption of colloidal mucin onto the particle surface.[13] These results, though similar, are likely due to a different mechanism, as agarose is not known to exist in a colloidal phase capable of adsorbing onto the particle surface and should be fully incorporated into the matrix scaffold.[26,28] A particularly interesting question

for future consideration would be whether the glycocalyx, the diverse array of polymeric sugar molecules expressed on the surfaces of cells and cell-derived EVs might play analogous roles in vivo.

Conclusions

Here, we have demonstrated the applicability of agarose as a bio-inert and non-adhesive model for investigating the non-specific steric and electrostatic interactions of vesicles with a 3D polymer matrix. The use of a bottom-up biomimetic approach to study lipid vesicle diffusion in hydrogel materials has shown that different biophysical factors contribute to particle dynamics. Even without specific biochemical interactions via cell adhesion molecules and ECM proteins, complex diffusive behaviour can arise through the combined effects of non-specific steric and electrostatic interactions. Such factors might contribute to preferential infiltration of some vesicle populations into tissues with certain ECM architectures, forming a basis for organotropic "homing" of particles.

We have presented experimental data that support and bring together existing computational and theoretical work on particle diffusion through different ECM-like environments. We show that polystyrene beads are a poor model for studying the diffusion of liposomes and that the effect of PEGylation on particle diffusivity is not due to electrostatic effects, as previously claimed, nor through the prevention of adsorption of soluble or colloidal materials on the particle surface, as no such species exist in our system. The design of lipid nanocarriers and engineered liposome-based therapeutics can thus take advantage of non-specific membrane-matrix interactions to achieve improved penetration and retention in target tissues. Further research on ECM-derived materials and more complex lipid vesicle systems is needed to better understand the diverse molecular interactions that govern the movement of vesicles in the tissue microenvironment.

Experimental Section

LUV production and characterization. Lipid stocks dissolved in chloroform (Avanti Polar Lipids) were used to prepare mixtures containing 4mM DOPC (1,2-dioleoyl-sn-glycero-3-phosphocholine) as a base solution. Negatively charged LUVs were made with a 2:1 molar ratio mixture of DOPC and DOPS (1,2-dioleoyl-sn-glycero-3-phospho-L-serine), while positively charged LUVs were made with a 2:1 molar ratio mixture of DOPC and DOTAP (1,2-dioleoyl-3-trimethylammonium-propane). PEGylated lipids were used to passivate vesicles for diffusion in hydrogels. At room temperature, all lipids are above their main phase transition temperatures and no demixing in the membrane is expected to occur. Additions of 1 mol% or 10 mol% DSPE-mPEG1K, DSPE-mPEG2K, or DSPE-mPEG5K (1,2-distearoyl-sn-glycero-3-phosphoethanolamine-N-[methoxy(polyethylene glycol)-1000], -2000], and -5000], respectively) were added to base solutions of DOPC or DOPC/DOTAP. Fluorescent visualization was facilitated by the addition of 0.2mol% $DiIC_{18}(5)$ (Thermo Fisher Scientific; 1,1'-Dioctadecyl-3,3,3',3'-Tetramethylindodicarbocyanine, 4-Chlorobenzenesulfonate Salt).

LUVs were produced by first spreading a thin layer of a lipid mixture inside a glass vial and drying under vacuum for 1.5 hours. Next, the lipid film was hydrated with phosphate buffered saline (PBS; tablets for 200 mL solutions from Sigma-Aldrich) and vortexed for 30 minutes to produce multilamellar lipid structures. The resulting solution was then extruded with a Mini Extruder (Avanti Polar Lipids), 21 passes each through a 200 nm and then 100 nm polycarbonate Nuclepore Track-Etched Membrane (Sigma-Aldrich).

Size distribution and zeta-potential of particles were measured using a Malvern Instruments Nano-ZS Zetasizer equipped with a 632.8 nm 4mW HeNe laser to ensure sample consistency. Samples in disposable folded capillary cells (DTS1070; Malvern Panalytical) were analyzed with dynamic light scattering (DLS) at a scattering angle of 173° to determine size distribution before determination of zeta-potential. All particles were measured in high-salt buffer conditions resulting in electrostatic screening, so zeta-potential values are used to illustrate relative differences in surface charge rather than absolute charge.

Fluctuation analysis. To probe how the bending rigidity of lipid membranes changes with the presence of PEGylated lipids, we used fluctuation analysis on giant unilamellar vesicles (GUVs).[47,48] GUVs were made using the gel-assisted swelling method[49,50] (see Section S7 in the SI). Briefly, 20 µL 5% w/v solution of polyvinyl alcohol (PVA; fully hydrolyzed, MW = 145000Da; Merck Group) in water with 50 mM sucrose was spread onto a 2 cm by 5 cm area corresponding to the dimensions of a rectangular, 2 mm-thick Teflon spacer and allowed to dry completely in an oven at 50°C. Next, a thin 15 µL layer of 4 mM lipid mixture dissolved in chloroform was spread on top of the PVA layer and dried in a vacuum for 1.5 hours. The slide was then assembled into a sandwich with another glass slide and a Teflon spacer in the middle, held together with binder clips (Fig. S7). The lipid layer was hydrated for 30 minutes with 2 mL PBS + 50 mM sucrose (345 mOsm/kg). The sucrose was necessary to help with the swelling process and to generate a sugar gradient that would later aid in visualizing the GUVs. GUVs were harvested and diluted 1:1 in a solution of PBS + 100 mM glucose (394 mOsm/kg) to slightly deflate the GUVs and were visualized under phase contrast with a 40× objective on a Zeiss AXIO Observer.D1 microscope. Image sequences of 3000 frames were recorded with a pco.edge sCMOS camera at 25 frames per second (fps) with 200 µs exposure. Fluctuation analysis software[48] computed the bending rigidity based on the Fourier decomposition of thermally-driven membrane fluctuations into spherical modes.

Preparation and characterization of agarose gels. Stock solutions of 2% w/v low gelling-temperature agarose (BioReagent, for molecular biology; Sigma) were made by dispersing agarose powder in PBS and microwaving at 350 W power in 5-8s intervals until dissolved. Stocks were stored at 4°C and could be re-melted at 95°C using the same microwaving method. The

molten agarose remained liquid down to 35°C. Gels of 1%, 0.5%, and 0.2% concentration were formed by melting stock gels and mixing with warm PBS (35°C) directly on glass slides for imaging, kept warm on a hotplate set to 35°C, or directly on a heated rheometer stage in the case of rheology measurements. Molten gels were taken off heating apparatus to cool to ambient temperature to induce gelation. Gel osmolality, which influences degree of vesicle deflation, was varied with the addition of glucose as opposed to salts to maintain the ionic strength of the solution, avoiding electrostatic screening effects. Solution osmolality prior to the addition of agarose was adjusted with a freezing-point osmometer (Osmomat 3000, Gonotec, Germany). A list of tested gel formulations can be found in Table 2 in the SI (Section S3).

Bulk rheology of agarose hydrogels was studied in shear mode using an Anton-Paar MCR301 rheometer with 12 mm cone-plate (CP12) geometry. Gels were mixed directly on the rheometer stage heated to 35°C, then cooled below 20°C to allow the sample to start to set while the probe was lowered to the measurement position on the sample. The gel was left for 5 minutes to fully set before testing up to 1% rotational strain from 1-10 Hz.

Average gel pore size was estimated using a turbidimetric assay described by Aymard et al.[30] and Narayanan et al.[27] Briefly, molten agarose was added to disposable 2.5 mL PMMA cuvettes (Sigma-Aldrich) and allowed to cool to ambient temperature (~22°C) to gel. Absorbance values over 600-900 nm were measured using a Thermo Spectronic Helios Gamma UV-Vis Spectrophotometer (Thermo Fisher). This was compared to analytical data from Aymard et al.[30] (see Section S4, SI).

Quantifying particle mobility. LUVs were embedded in agarose gels by mixing extruded LUV solutions with molten agarose directly on a glass microscopy slide within a rubber spacer (see Section S1, SI). A glass coverslip was placed on top, such that the agarose droplet wetted both glass surfaces, forming a disk. The imaging chamber was set aside at room temperature for 5 minutes to set. For control experiments with embedded polystyrene beads, working mixtures of Fluoresbrite YG 0.1 µm-diameter Microbeads (Polysciences, Inc) and FluoSpheres carboxylate-modified 0.1 µm-diameter red (580/605) polystyrene beads (Invitrogen) were made by diluting bead suspensions 1:100 in PBS before being mixed into gels, replacing the LUV solution at the same volume.

Samples were imaged with a pco.edge sCMOS camera mounted to a Zeiss AXIO Observer.D1 microscope with a 63× water immersion objective in epifluorescence mode with appropriate excitation and emission filters. Image sequences of length 5s (~100 frames) were captured with 20 fps frame rate and ~45 ms exposure in a 100 µm × 100 µm region of interest (ROI). Three ROIs were recorded per sample to account for internal heterogeneity. Particle mobility within gels was analyzed with the single-particle tracking plugin for FIJI developed by Sbalzarini and Koumoutsakos.[25] A sequence length of 100 frames was chosen because longer sequences resulted in decreased signal-to-noise ratio from photobleaching, leading to increased false positives in particle detection. Histograms of the $\log_{10}$ diffusion coefficients (in m^2s^{-1}) obtained from the plugin were used to determine the mobile fraction using -14 as the cut-off. This value corresponds to the lower limit of detection of particle movement based on imaging parameters (see Section S1, SI).

We also studied the infiltration of 100 nm-extruded DOPC LUVs into preformed 1% w/v agarose hydrogels to obtain a collective diffusion coefficient. The prior imaging chamber setup was done with slight adjustments (see Fig. S2, SI). Briefly, agarose gel disks were formed without LUVs in an imaging chamber and allowed to set before a solution of LUVs (12 µM lipid) was pipetted into the chamber. Three 100 µm × 100 µm ROIs each in the gel interior and the exterior solution were imaged per time point, per sample over 100 hours. The number of particles in each ROI inside and outside the gel was counted as a function of time. Assuming slow diffusion and obtaining density gradients using finite differences, the diffusion coefficient was determined as follows: each ROI is a rectangular box with dimension $h = 100\mu m$ and $L = 100\mu m$. The flux per unit area, J flowing into the ROI in the gel interior is given by

$$J = \frac{1}{L}\frac{dN_{in}}{dt},$$

where N_{in} is the number of LUVs in the gel interior and t is time. Due to the slow diffusion, we assume N_{in} to vary linearly in time, hence

$$\frac{dN_{in}}{dt} = \frac{N_{in}}{t}.$$

Fick's first law of diffusion connects the density $\varphi = \frac{N}{hL}$ with the flux by introducing a diffusion constant, D *via*

$$J = -D\frac{d\varphi}{dx} = -\frac{D}{hL}\frac{dN}{dx}$$

Finally, we approximate the density gradient using a finite difference

$$\frac{dN}{dx} = \frac{N_{in} - N_{out}}{x},$$

where $x = 300\mu m$ is the distance between the ROI of the gel interior to the edge of the gel and N_{out} is the number of particles in the ROIs in the exterior LUV solution. Combining these relations and solving for the diffusion constant gives

$$D = -\frac{N_{in}hx}{(N_{in} - N_{out})t}.$$

Because of the slow diffusion, all initial timepoints where $N_{in} = 0$ are excluded. Diffusion coefficients were calculated at each timepoint measured (4-13 data points per replicate over 100 hours of imaging, see Fig. S2D in SI), then averaged for each replicate.

Statistical analysis. Histograms of $\log_{10}$ diffusion coefficients are normalized to show probability and represent pooled data from three ROIs within an individual gel. Each ROI corresponds to 100-300 diffusing particles. The variability in the number of identified particle tracks in different gels arises from differences in particle mobility. Statistics on mobile fractions were calculated with n=3 gels, presented as standard box plots. All other data are presented as mean with standard deviation. Statistical significance was determined with N-way ANOVA (as indicated) with Tukey-Kramer tests for multiple comparisons at the significance levels indicated, computed using MATLAB.

Acknowledgements

N.W. Tam would like to acknowledge funding from the International Max Planck Research School on Multi-Scale Biosystems, as well as Alexander Becker, a summer intern supported by the German Academic Exchange Service (DAAD) RISE research internship program, for his help with experiments. A. Cipitria would like to acknowledge funding from the DFG Emmy Noether grant (CI 203/2-1), from IKERBASQUE Basque Foundation for Science, and from the Spanish Ministry of Science and Innovation (PID2021-123013OB-I00).

Supporting Information

Supplementary materials, including list of abbreviations used, details on methodology, theoretical derivations, and additional vesicle and hydrogel characterization data

Conflict of Interest Disclosure

The authors declare no conflict of interest.

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N.W. Tam did experiments and wrote the manuscript. O. Schullian analyzed experimental data and provided theoretical derivations. R. Dimova and A. Cipitria contributed to project supervision and editing of the manuscript.

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