Protein Dynamics Govern the Oxyferrous State Lifetime of an Artificial Oxygen Transport Protein

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ABSTRACT. It has long been known that the alteration of protein side chains which occlude or expose the heme cofactor to water can greatly affect the stability of the oxyferrous heme state. Here we demonstrate that the rate of dynamically-driven water penetration into the core of an artificial oxygen transport protein also correlates with oxyferrous state lifetime by reducing global dynamics, without altering the structure of the active site, via the simple linking of the two monomers in a homodimeric artificial oxygen transport protein using a glycine-rich loop. The tethering of these two helices does not significantly affect the active site structure, pentacoordinate heme binding affinity, reduction potential, or gaseous ligand affinity. It does, however, significantly reduce the hydration of the protein core as demonstrated by resonance Raman spectroscopy, backbone amide hydrogen exchange, and pKa shifts in buried histidine side chains. This further destabilizes the charge-buried entatic state and nearly triples the oxyferrous state lifetime. These data are the first direct evidence that dynamically-driven water penetration is a rate-limiting step in the oxidation of these complexes. It furthermore demonstrates that structural rigidity which limits water penetration is a critical design feature in metalloenzyme construction and provides an explanation for both the failures and successes of earlier attempts to create oxygen-binding proteins.

Secondary Abstract

Simple designed proteins offer a great opportunity to reveal the engineering parameters that underlay complex natural protein function. Here we use a relatively small perturbation - making a homodimeric oxygen transport protein into a single chain form - to reduce the overall conformational flexibility of a protein fold without detectably changing the structure of its active site. This reduces the ability of water to penetrate into the core of the protein where the oxygen-

binding cofactor lies, extending the lifetime of the oxygen bound state more than threefold. This demonstrates that the control of water penetration is an important aspect of metalloenzyme catalysis, and thus metalloenzyme design.

Introduction

The hallmark of hexacoordinate hemoglobins is that they have two ligands in the oxidized state, but in the reduced state, they exist in a mixed hexa- and pentacoordination state in which one of the ligands – termed the distal ligand - has a weak affinity for the ferrous heme iron (1,2). The partial pentacoordination of the heme cofactor allows for the binding of molecular oxygen at the

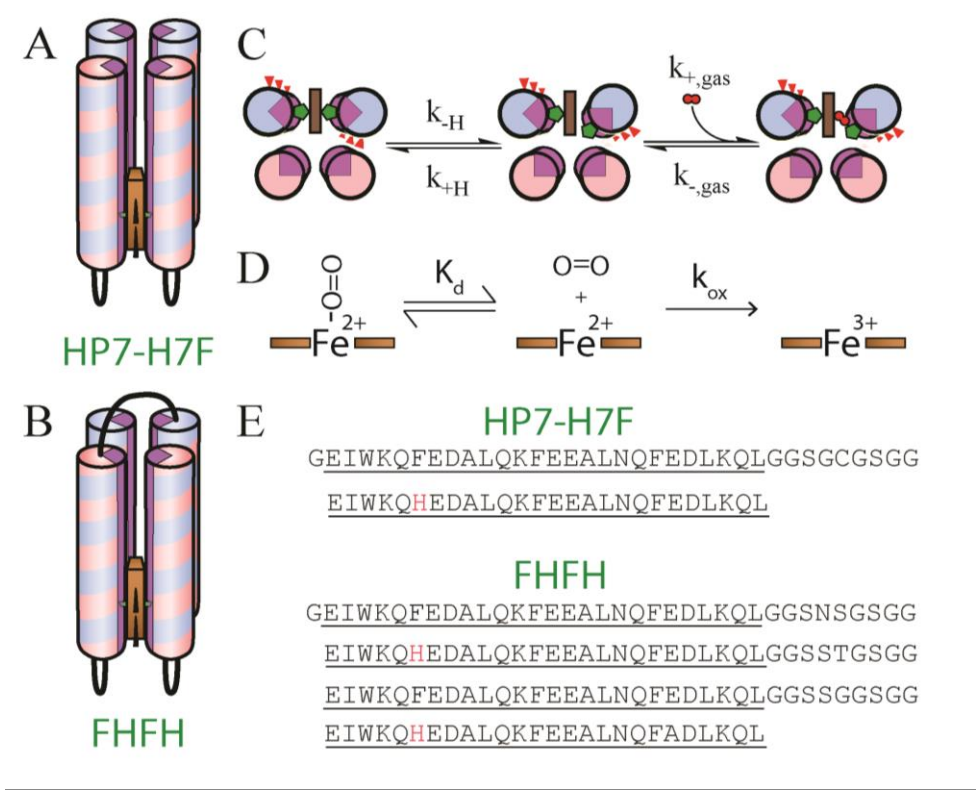

Figure 1. Structure, sequence and oxygen binding mechanism of two artificial oxygen transport proteins. Topology of the homodimeric protein HP7-H7F (A) and the single-chain protein FHFH (B). (C) Mechanism of oxygen binding to the ferrous heme. (D) Equilibrium and rate constants of the oxidation of the oxyferrous state. (E) Sequences of the two proteins.

ferrous iron. In the case of the artificial oxygen transport protein HP7 this destabilization is the result of the coupling of histidine side chain ligation with the burial of three charged glutamate residues on the same helix, creating a high energy 'entatic' conformational state (3-6). Relief of the strain of buried charges by the rotation of the glutamates into solution necessitates the detachment of the adjacent histidine side chain.

Oxygen binding by heme proteins is complicated by the requirement that the heme iron be in the ferrous state for oxygen ligation. When exposed to oxygen, the ferrous heme in oxygen transport proteins can be oxidized by collisional electron exchange when not bound to molecular oxygen, but once formed, the oxyferrous complex has a reduction potential too low to be oxidized by a second molecule of oxygen (Figure 1D) (7). As we have shown (4), the rate of oxidation of the oxyferrous state is thus a function of the total protein concentration, the total oxygen concentration, the oxygen dissociation constant K_d, and the rate constant for the reaction between oxygen and the ferrous heme complex, k_{ox}:

(1) $$Rate = k_{ox} \cdot K_d \cdot \frac{([P]+[O_2]+K_d)-\sqrt{([P]+[O_2]+K_d)^2-4[P][O_2]}}{2}$$

Oxidation of the oxyferrous complex in model systems is greatly accelerated by water (8-10). This is best demonstrated by the fact that imidazole-ligated hemes embedded in non-polar plastics stably form oxyferrous states with several-minute lifetimes, while the same systems dissolved in water rapidly oxidize, forming superoxide (11,12). Similarly, mutations of active site residues in bovine myoglobin which sterically increase or decrease the solvent exposure of the bound heme cofactor decrease and increase, respectively, oxyferrous state lifetimes by over two orders of magnitude, and this was shown to be of a concerted reaction consisting of the

protonation of the bound oxygen by water and the dissociation of neutral superoxide (9). However, water penetration to buried heme sites in proteins driven by sidechain and backbone dynamics has not yet been demonstrated to affect lifetimes. Dynamic effects have been suggested to play a role by computational simulation of the autooxidation of myoglobin (13), and Olson and coworkers have shown that dilute preparations of human hemoglobin dissociate into dimers which oxidize more than 10-fold more quickly than the native tetramer (14). However, neither the conformational distribution nor the water penetration of the latter have been characterized.

The artificial diheme protein HP7 has proven to be a useful model system, or protein "maquette" (15), for the study of natural oxygen transport proteins. In an examination of a simplified version of the protein containing only one heme binding site, we have demonstrated that the mutational removal of all three buried glutamate side chains from the remaining heme-binding helices in HP7-H7F results in stronger distal histidine affinity, which strengthens heme cofactor binding (15,16), but weakens carbon monoxide binding by competition. However, the mutant protein does not stably bind oxygen. This loss of function was attributed to either or both of two observable effects of the mutation: the 22-fold decrease in the rate of detachment of the distal histidine ligand, which results in a longer time period for the bis-histidine bound ferrous heme to be collisionally oxidized by free molecular oxygen before the oxyferrous state can be formed, or the greatly increased degree of water penetration into the protein core, evinced by the fact that hydrogen-deuterium solvent exchange is so fast in the mutant protein that it is complete before NMR data collection can occur. The former works to reduce the yield of oxyferrous state upon mixture of the reduced heme protein with oxygen, while the latter shortens the lifetime of the oxyferrous state after formation by speeding oxidation during transient pentacoordination

subsequent to oxygen binding (see Figure 1D).

Here we demonstrate that changes in protein dynamic motions affect both the rate of water penetration into and hydration of the hydrophobic core of HP-7 by introducing a simple, nonperturbative connecting loop between the two helices which comprise the heme-binding active site. There is a precedent for protein stabilization by fashioning a single-chain protein from a self-assembling helical bundle – when Degrado and coworkers transformed a nonfunctional homotetrameric diporphyrin protein (17) that they had designed into a single-chain tetraporphyrin protein, the overall structure became so stabilized that a number of destabilizing mutations had to me made in order to make the protein flexible enough that it could bind cofactors under non-denaturing conditions (18). Other single-chain four helix heme-binding bundles have been designed (19-21), but their oxygen transport properties were not assessed.

This minimal modification to HP-7 extends the oxyferrous state lifetime without significantly affecting gaseous ligand binding – the first direct evidence that dynamically-driven water penetration is a rate-limiting step in the oxidation of these complexes. This loop reduces water penetration by limiting the number of possible open conformations, greatly reducing the size of the protein conformational ensemble and consequently the population of open states competent for solvent penetration.

Materials and Methods

Chemicals. Hemin was purchased from Fluka (Buchs, Switzerland). Molecular oxygen (O_2) (99.98% purity), carbon monoxide (CO) (99.9%), and molecular nitrogen (N_2) (99.99%) gases were from Matheson Gas (Basking Ridge, NJ). N_2 and CO were scrubbed of residual O_2 by passage through two bubblers filled with a reduced vanadium sulfate solution followed by

another filled with water (22). PD-10 desalting columns were from GE Healthcare (Port Washington, NY). All other solvents and reagents were from either VWR or Sigma.

General biochemistry. FHFH was expressed and purified as described for other heme protein maquettes earlier (23). Optical spectra were collected with a Hewlett-Packard (New York, NY) 8452A Diode array spectrophotometer running the Olis (Bogart, GA) SpectralWorks software and equipped with a Quantum Northwest (Liberty Lake, WA) Peltier temperature controller. Each binding, reduction potential, and stopped-flow experiment was performed at least three times, and reported errors are standard deviations from the mean. All experiments were performed at 20° C in 250mM Boric Acid, 100mM KCl pH 9.0 unless otherwise specified. Hemin stock solutions of approximately 0.5-1.0 mg/ml were prepared in DMSO and used within four hours. Stock solution concentrations were determined using the pyridine hemochrome assay (24).

Holoprotein complexes, containing 1.0 equivalent per homodimer, were prepared as before (5) by five consecutive additions of 0.2 equivalents of hemin in DMSO with at least ten minutes between additions. Proteins were then purified from any unbound heme using PD10 desalting columns. Holoprotein solution concentrations were determined using the experimentally derived Soret extinction coefficients of ε_{414} = 129 $mM^{-1}cm^{-1}$ for HP7-H7F and 122 $mM^{-1}cm^{-1}$ for FHFH.

Heme affinity measurements. For oxidized binding experiments approximately 0.1 molar equivalent aliquots of hemin solution were added via gastight syringe to a stirring 4 ml solution of 2-3 μM protein, with a ten minute equilibration delay between additions. Heme binding was

monitored by loss of the absorption at 385 nm due to free hemin, and the concomitant appearance of a sharp Soret band at 412 nm, corresponding to heme bound to the protein via bis-histidine axial coordination. The number of heme binding sites was quantified for each protein from plots of the Soret maximum at 412 nm vs the number of equivalents added.

Reduced binding experiments were performed on solutions of ~3 μM protein as described previously (5). Briefly, hemin was added in approximately 0.1 molar equivalent aliquots to anaerobic protein solution kept at a potential less than -450 mV using periodic additions of sodium dithionite. K_d values were obtained from plots of the Soret band absorbance measured at 436 nm vs. the concentration of hemin added and fit with the tight binding equation:

$$(2) \qquad \Delta A = \mathrm{A_{start}} + \varepsilon_{\mathrm{unb}}[\mathrm{hem}] + \varepsilon_{\mathrm{bnd}}[\mathrm{prot}] \cdot \left[\frac{\mathrm{K_d}+[\mathrm{hem}]+[\mathrm{prot}]-\sqrt{(K_d+[\mathrm{hem}]+[\mathrm{prot}])^2-4[\mathrm{hem}][\mathrm{prot}]}}{2[\mathrm{prot}]}\right]$$

Where $\varepsilon_{\mathrm{unb}}$ is the molar absorption coefficient of unbound hemin at that wavelength, $\varepsilon_{\mathrm{bnd}}$ is the additional absorbance of bound hemin at that wavelength, [hem] is the total hemin concentration, [prot] is the total protein concentration and K_d is the dissociation constant for the reduced hemin.

Reduction potential determination. Spectroelectrochemical redox titrations were performed as described previously (5) using 10-20 μM solutions of hemoprotein containing >100 μM of the corresponding apoprotein in order to eliminate the possibility of heme dissociation upon reduction. Reported reduction potentials are referenced to a standard hydrogen electrode. All spectroelectrochemical titrations were performed anaerobically using μL additions of freshly prepared sodium dithionite to adjust the solution potential to more negative values and potassium ferricyanide to more positive values. Redox mediators were as used previously (25). Titrations

were analyzed by monitoring the Q band absorbance at 559 nm as the heme protein was reduced or oxidized and data fit with the Nernst equation using an *n*-value of 1.0:

(3) $$\%\boldsymbol{R} = \frac{\mathbf{1}}{\mathbf{10}^{(\boldsymbol{E}-\boldsymbol{E_m})/[(\frac{\boldsymbol{RT}}{\boldsymbol{nF}})]}}$$

where %R is the fraction of reduced heme, E is the solution potential, E_m is the reduction midpoint potential, and n is the number of electrons.

Flash photolysis analysis of distal histidine association. Rate constants for CO and histidine binding to the pentacoordinate state were determined using laser flash photolysis. Transient absorbance data were collected on an Edinburgh Instruments LPK920 laser flash photolysis spectrometer equipped with an Opotek Vibrant 355II tunable laser source. Excitation at 532 nm was used to excite the preformed carbonmonoxyferrous complex causing the detachment of the ligand CO and the rates and magnitudes of CO and histidine association were determined by analyzing multi-exponential rebinding traces taken as a function of CO concentration using the method of Hargrove (26):

(4) $$\gamma_1 + \gamma_2 = k_{-His} + k_{+His} + k_{+CO}[CO]$$

(5) $$\gamma_1 \gamma_2 = k_{-His}k_{+CO}[CO]$$

Where γ_1 and γ_2 are the fitted first and second CO-dependent exponential rates and the kinetic constants are defined as in Equation 5. Protein concentrations were 20-25 μM and

carbonmonoxyferrous complexes were prepared by reducing solutions of the holoproteins with an excess of dithionite under an atmosphere containing 10-100% CO mixed with N_2.

Stopped Flow analysis of gaseous ligand binding. Binding kinetics of O_2 and CO with ferrous protein-heme complexes were determined as before (4) using a Biologic (Lyon, France) SFM 300 stopped flow mixer equipped with a custom-built Olis RSM 1000 spectrometer for multiwavelength detection. Binding was followed spectroscopically over gas concentrations from 2% to 50% saturation at 15°C. Air- or gas-saturated buffer was mixed with N_2-saturated buffer in the first mix, and then this mixture was combined in a second mix with an N_2-saturated ferrous heme protein solution. Protein concentrations were 20 μM and ferrous complexes were prepared by carefully titrating anaerobic solutions of the holoproteins with a slight excess of sodium dithionite as observed by visible spectroscopy followed by anaerobic canular transfer to the stopped-flow loading syringe. Binding kinetic data were fit with Eqn. 6, which assumes that O_2 binding rate, k_{+O2} , is much greater than the sum of the distal histidine association and dissociation rates (27):

$$k_{obs} = \frac{k_{-H} k_{+gas}[gas]}{k_{+H} + k_{-H} + k_{+gas}[gas]} \quad (6)$$

Where k_{obs} is the fitted single exponential binding rate, k_{+H} and k_{-H} are the distal histidine-ferrous heme iron association and dissociation rates, and k_{+gas} is the gaseous ligand association rate constant. At high [gas], $k_{obs} = k_{-H}$.

Kinetic analysis of CO dissociation. The dark dissociation rate of the CO complex was

determined using the ferricyanide trapping method of Moffet *et al* (26): briefly, anaerobic solutions of carbonmonoxyferrous protein complex were mixed with varying concentrations of potassium ferricyanide and the rates of the linked reactions of CO dissociation followed by ferricyanide heme oxidation were followed by monitoring the disappearance of the carbonmonoxyferrous Soret peak at 421 nm. The rates of oxidation *vs.* the concentration of ferricyanide give the inverse CO dissociation rate as given by Eqn. 7:

(7) $$\frac{1}{k} = \frac{k_{on}[CO]}{(k_{ox})(k_{off})}\left(\frac{1}{[FC]}\right) + \frac{1}{k_{off}}$$

Kinetic analysis of O_2 dissociation. The dissociation rate of the O_2 complex was determined using the CO displacement method of Gardner et al (28). Briefly, the O_2 complex was formed as above by mixing reduced protein and O_2-saturated buffer in a stopped-flow spectrophotometer followed by a second mix a100 ms later with a two-fold larger volume of CO-containing buffer. The rate of O_2 replacement by CO was collected using final CO concentrations ranging from 200-600 µM. At high CO concentration, the observed rate of replacement, r_{obs} is given by Eqn. 7:

(8) $$r_{obs} = k_{-O_2}/[1 + \left(\frac{k_{+O_2}[O_2]}{k_{+CO}[CO]}\right)]$$

When $k_{+co}[CO] \gg k_{+O2}[O_2]$, the observed replacement rate constant is directly equal to O_2 dissociation rate constant: $r_{obs} \approx k_{-O_2}$.

Oxyferrous state lifetimes. As before (4), the lifetimes of O_2-bound ferrous protein-heme complexes were determined spectroscopically in stopped-flow mixing experiments using final O_2

concentrations from 16.67% to 66.67% saturation at 16° C. Protein concentrations were 10 μM with 1 μM hemin added to ensure full complexation. Ferrous complexes were prepared by carefully titrating anaerobic solutions with a slight excess of sodium dithionite as observed by visible spectroscopy. Ferrous samples were anaerobically transferred to the stopped-flow loading syringe by canula. Kinetic data were fit with Eqn. 1.

Nuclear magnetic resonance. All NMR experiments were performed at 20ºC using a Varian Inova spectrometer operating at a 600MHz and equipped with a triple resonance cryogenic probe capable of applying pulse field gradients in the z-direction. Chemical shifts are referenced to water at 4.77 ppm for ^{1}H. Data were processed using the program NMRPipe (29) and analyzed using Sparky (30).

Sensitivity enhanced ^{1}H-^{15}N heteronuclear single quantum coherence (HSQC) spectra (31) were collected to assess structural specificity (23) on 50-100 μM holo- and apoprotein samples with sweep widths of 10,000 Hz for ^{1}H and 2000 Hz for ^{15}N utilizing GARP decoupling of ^{15}N during ^{1}H acquisition.

In order to determine the pK_a of histidine side chains, the ^{1}H-^{15}N multiple-bond correlation signals originating in the imidazole side chains were collected as before (5) using a non-sensitivity enhanced HSQC pulse sequence in which the INEPT (insensitive nuclei enhanced by polarization transfer) periods were set to $1/^{1}J_{NH}$, where $^{1}J_{NH}$ is the one-bond ^{1}H-^{15}N coupling constant, in order to attenuate backbone amide signals (32). Spectral widths were 8000 Hz for ^{1}H and 12156 Hz for ^{15}N.

^{15}N-labelled apoprotein samples were 200 μM in 25mM KaD_2PO_4 D_2O buffer with the pD (pH meter reading + 0.4 pH units) initially fixed at 5.0. After each spectrum was taken the

pD was adjusted upwards by the addition of small amounts of KOD dissolved in D_2O. The observed ^{15}N chemical shifts were fit with the Henderson-Hasselbach equation assuming fast chemical exchange between protonation states:

$$(9) \qquad \delta_{obs} = \frac{\Delta\delta_{prot}}{1+10^{[pD-pKa]}} + \delta_0$$

where δ_0 is the neutral chemical shift of ^{15}N, $\Delta\delta_{prot}$ is the change in chemical shift due to protonation, and pK_a is the fitted acid dissociation constant.

Hydrogen-Deuterium exchange. To compare the rates of backbone hydrogen exchange (33), ^{15}N-labelled holo- and apoprotein samples of HP7-H7F and FHFH in 25 mM potassium phosphate buffer pH 6.5 were passed through a PD-10 solvent exchange column pre-equilibrated in 25 mM KD_2PO_4 D_2O buffer pD 6.5 (pH meter reading + 0.4 pH units), immediately placed in a 5mm NMR tube, and equilibrated at 20°C for five minutes in the sample compartment of the NMR spectrometer. Final protein concentrations were approximately 300 μM. The fraction of remaining amide protons was assessed as a function of time by collecting one dimensional ^{15}N-HSQC, or isotope-selective, ^{1}H spectra.

Deep UV resonance raman (dUVRR) spectroscopy. The dUVRR spectra of the FHFH and HP-7-H7F were acquired as described previously (34) using a custom-built instrument (35). Incident laser power at the sample was attenuated to 500 μW to avoid protein degradation, and spectra were monitored for degradation over time using the aromatic ring modes. Spectral calibration was carried out using a standard cyclohexane spectrum (36). All dUVRR spectral preprocessing

was carried out in MATLAB (7.1, MathWorks, Natick, MA) using cosmic ray and water band removal methods described previously (37). A nonlinear least-squares algorithm was used to fit the amide and aromatic bands to mixed Gaussian/Lorentzian peaks, which approximate the Voigt line shape as described previously (38).

As standard reductants such as dithionite have significant molar absorptivity coefficients in the deep-UV, spectra of the reduced forms of each heme protein were acquired using titanium(III) citrate as the reductant. Concentrated solutions of titanium(III) citrate were generated by adding 5 mL of an Ar purged buffered solution of 0.2 M sodium citrate (pH 7) to approximately 75 mg of $TiCl_3$ to produce titanium(III) citrate in an Ar purged crimp top vial. H4 and HP7 protein samples (4.5 µM and 5.3 µM) were prepared in 25 mM borate, 10 mM KCl, 50 mM $NaClO_4$ buffer (pH 9.5) and purged with Ar in similar vials. Air exposed samples were confirmed for reduction and reoxidation by UV-Vis spectroscopy. 5 µL aliquots of titanium(III) citrate solution were added to 1 mL of each protein sample until the absorption band at 560 nm reached a maximum intensity samples remained fully reduced for 15-20 minutes, which defined the maximal time frame for individual sample collection in the dUVRR sample flow cell.

Results

Protein design. Water penetration into the hydrophobic core of a globular protein like HP7-H7F is driven by large-scale or even global dynamic protein motion (33). In order to restrict these global motions, we surmised that a simple glycine-rich connecting loop similar to that we previously used to create the light-activated electron transfer dyad protein HHHF (39) could reduce water penetration by limiting the number of possible open conformations, greatly reducing the size of the protein conformational ensemble and consequently the population of

open states competent for solvent penetration. Thus we modified the homodimeric "candelabra" structure (23) by connecting the second helix of one monomer of HP7-H7F to the first helix of the other, resulting in the single chain protein FHFH (Figure 1A-C).

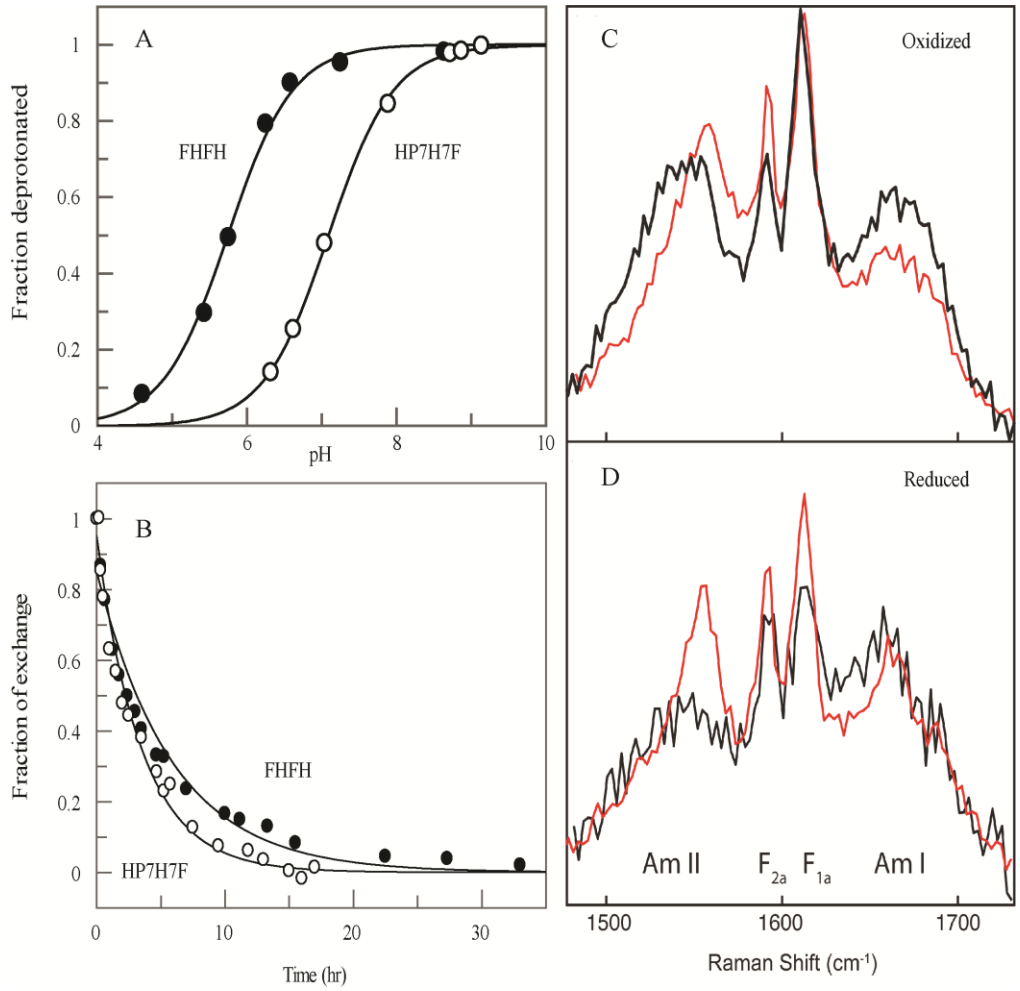

Figure 2**.** Experimental demonstration of reduced hydration in the core of FHFH. (A) NMR determination of the pK_a values of the histidine ligands in apo-HP7-H7F (○) and FHFH (●). Lines drawn are fits of the $^{15}N(\varepsilon 2)$ chemical shifts with equation 9 having pK_a values of 7.3 and 5.9 respectively. (B) Hydrogen-deuterium exchange in the ferric heme complexes of HP7-H7F (○) and FHFH (●). Lines drawn are exponential fits to the integrated data with exchanges rates of 0.17 s-1 and 0.29 s-1 respectively. (C and D) Deep-UV resonance Raman spectra of the oxidized (C) and reduced (D) forms of the HP7-H7F (black) and HFHF (red) are shown for the spectral region containing the backbone Amide I and II mode as well as two hydration-sensitive ring related aromatic modes.

FHFH displays visible absorbance spectra indicative of bis-histidine coordination in both oxidation states (Figure 3A) with Soret and Q-band peak maxima identical to those of heme bound HP7-H7F. It is important to confirm the helical bundle structure of the protein, as some designed helical bundle proteins have been found to form domain-swapped higher oligomeric states (40). Both apo- and holoprotein forms of FHFH are monomeric as analyzed by size exclusion chromatography (not shown). Furthermore, while the apoprotein appears to be helical as detected by circular dichroism, it has ^{15}N-HSQC spectra indicative of molten globular structure. The addition of a single heme cofactor changes the HSQC spectra to a level of chemical shift dispersion indicative of a partial phase transition in which the two helices which ligate the heme become native-like and the two unliganded helices remain molten globular, consistent with what we have previously observed for HP7 (Supplemental Figure 1) (23).

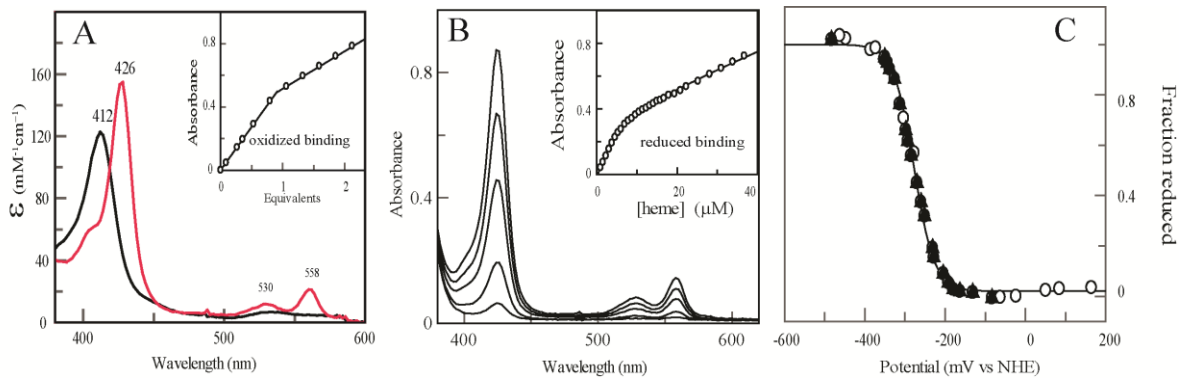

Figure 3. The thermodynamics of heme binding to FHFH. (A) Oxidized (black) and reduced (red) spectra and oxidized heme titration (inset) of heme-bound FHFH. Extinction coefficients were calculated using the intercepts from the end point titrations. (B) Reduced heme binding to 3.0 μM FHFH. Some spectra have been omitted for the sake of clarity. (Inset) Equilibrium binding isotherm derived from the titration in panel B. The line drawn is a fit with Eqn. 2 using a $K_{d,red}$ of 3.4 µM. (C) Equilibrium potentiometric determination of the electron affinity of FHFH. Line drawn is a fit with eq 3 with an E_{mid} of -277 mV vs NHE.

Core hydration. Interior group pKa's: Several different lines of evidence suggest that the

introduction of this loop significantly reduces the hydration of the protein interior: First, the pK_a of the buried ligand histidine side chains in the apoprotein is 1.4 pH units lower in FHFH than in HP7-H7F (see Figure 2A and supplementary Figure 2), corresponding to an increase of 1.9 kcal/mol in the energy required for protonation. While there are many local structural interactions which can shift side chain pK_as in either direction, the most common cause of pK_a shifts which destabilize the charged state of side chains in proteins is a decrease in the local dielectric (41).

Hydrogen/Deuterium exchange. A more rigorous method of comparing interior hydration is that of hydrogen/deuterium exchange (42). In whole-protein experiments, the bulk of backbone amide residues exchange in a few seconds at neutral pH and 25°C. These are the residues either directly exposed to or only minimally protected from solvent. Exchange of core residues, however, is typically greatly slowed – core proton exchange requires an open conformation which exposes these backbone amides to solvent. Such conformations are much higher in energy than lowest energy or 'Native-state' conformations. Exchange rates for these residues are thus the sum of the fractional Böltzmann populations of all of the open states competent for hydrogen exchange multiplied by the exchanges rates of these conformations.

We have previously shown that the core residues of ferric HP7-H7F exchange on a minutes time scale, too rapidly to collect two dimensional data on individual residue protection factors but slowly enough that 1-dimensional ^{15}N-selective ^{1}H NMR data on the exchange of the entire amide proton population can be quantified (4). A similar one-dimensional analysis of ferric FHFH demonstrates that core residue hydrogen exchange is almost two-fold slower (Figure 2B). Thus the addition of the loops restricts water access to the protein core, either by

making a subset of these open conformations impossible to form or too high in energy to populate, effectively reducing the internal water concentration.

Deep UV Resonance Raman. A complimentary method for examining the hydration of protein interiors is dUVRR. This vibrational technique utilizes resonance enhancement of both the peptide backbone and aromatic residue Raman modes to provide information about the ensemble hydration status of the backbone (amide I and II modes) as well as more regiospecific information via aromatic residue environment-derived changes. Here we focus in particular on the phenyalanine modes F_{1a} and F_{2a}, as this is the only aromatic residue type in the protein core – the sequences contain no tyrosine side chains and the tryptophan rings are purposely placed only at the solvent-exposed C-position termini of each helix for the purpose of protein concentration determination (25). Both the F_{1a} and F_{2a} mode intensities, as well as the intensity ratio of the amide I and amide II modes, have been shown in other proteins and model systems to change as a function of regional side chain dehydration (43-45).

As Figure 2C and 2D demonstrates, the magnitudes of the phenylalanine (F_{1a} and F_{2a}) and amide II mode intensities in both the ferric and ferrous heme complexes of the two proteins are markedly different. These differences are much larger than we have observed for the association of the helical pH low insertion peptide (pHLIP) with membranes and are closer to those we see for membrane insertion of pHLIP (46). These data clearly demonstrate that the protein interiors of both the ferric holo and ferrous holo states of FHFH are dehydrated in comparison to the same states of HP7-H7F, further evidence that the addition of the loop indeed does restrict the size of the conformational ensemble.

Ferric and ferrous heme binding. Given the change in core hydration without a significant

change in protein structure, it is important to determine the effect the introduction of this loop has on the binding of the heme cofactor. As has been observed for HP7-H7F (5), ferric heme binding to FHFH is too tight to reliably extract a dissociation constant. This enabled us to perform endpoint titrations, confirming that FHFH binds a single heme cofactor as well as calculate the extinction coefficient of the oxidized complex (Figure 3A).

Ferrous heme binding, however, is weak enough that binding energies can be quantitated under reducing conditions (Figure 3B). Ferrous heme binding is more than five-fold weaker in FHFH than HP7-H7F (Table I). The reduction potential of the bound heme in FHFH is 17 mV lower than that of HP7-H7F (Figure 3A), and in combination with the free heme reduction potential of -63 mV vs. NHE enables the calculation of the ferric heme binding affinity using a thermodynamic cycle (47). The calculated ferric heme affinity, 900 pM, is itself three-fold weaker than that of HP7-H7F.

Distal histidine affinity. A possible reason for the observed decrease in ferrous heme affinity in FHFH is a decrease in distal histidine affinity caused by the decreased dielectric in the protein core. We have previously shown that the overall heme binding affinity to hexacoordinate hemoglobins is (5):

$$(10) \qquad K_d = K_{d,pent}\left(\frac{k_{-His}}{k_{+His}+k_{-His}}\right) = \frac{K_{d,pent}}{(1+K_{A,his})}$$

Where $K_{d,pent}$ is the dissociation constant for the pentacoordinate heme complex and $K_{A,his}$ is the distal histidine association constant – the ratio of the distal histidine on- (k_{+His}) and off-rates (k_{-}

$_{His}$). k_{-His} were determined using CO binding: as the protein is primarily in the hexacoordinate bis-histidine ligation state in solution (Table 2) and the rate-limiting step for CO binding at high ligand concentrations is the detachment of the distal histidine from the heme iron (see eq 6).(5) Figure 4A depicts the concentration dependence of the CO binding rate to both proteins. Comparison of the asymptotic limits demonstrates that the histidine off-rate is four-fold slower in FHFH than HP7-H7F while k_{+His} rates, determined using the laser flash method of Hargrove (Figure 4C) (48), were within error of each other (Table I). Thus the decreased binding energy indeed seems to primarily result from decreased distal histidine affinity, which is itself likely a result of the increase in the energy required to move polar glutamic acid side chains into a lower dielectric environment.

Gaseous ligand binding- Figures 4B and 4C depict the experimental determination of the rates of CO binding (k_{+CO}) and dissociation (k_{-CO}) to both proteins. Both rates and the derived pentacoordinate binding constants, K_{d,CO_pent}, are within error of each other (Table II). One significant difference is seen in the ferricyanide trapping analysis of CO release (Figure 4B) – although the intercepts and thus the dissociation rates are equivalent, the slopes of the replots are significantly different, perhaps another result of restricted access to the heme in FHFH. The true binding constants involve competition between gaseous ligand binding and distal histidine coordination (27):

(11) $$K_{d,gas_actual} = K_{d,gas_penta}\left(1 + K_{A,His}\right)$$

Therefore the actual binding CO dissociation constants differ by a factor of four.

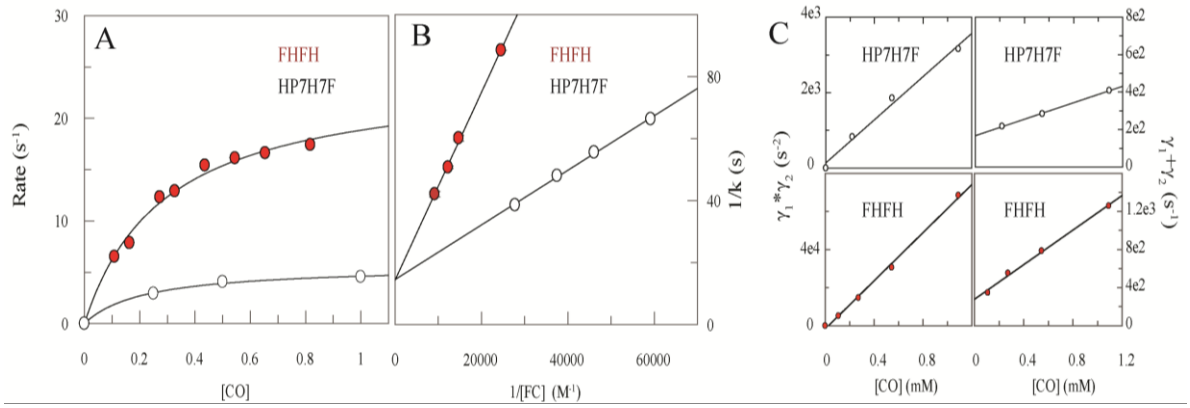

Figure 4 Kinetic analysis of CO and histidine binding to HP7-H7F (white circles) and FHFH (red circles). (A) Stopped-flow analysis of the rates of binding of CO to reduced heme proteins as a function of ligand concentration. Lines are fits with Eqn. 6. (B) Ferricyanide trapping analysis of CO release from FHFH and HP7H7F. Double-reciprocal plots of the rates of oxidation vs the concentration of ferricyanide extrapolated to infinite ferricyanide give the CO dissociation rate as shown in eq 7. (C) Laser flash kinetic analysis of histidine and CO rebinding. Flash data were fit with three exponentials, the first of which represents a CO concentration-independent relaxation process (see ref. 5). Replots of the products (left) and sums (right) of the second two exponentials. Fits with Eqns. 4 and 5 give the k_{+His} and k_{+CO}.

Similar behavior is observed with oxygen: Figure 5 depicts the stopped-flow analysis of the O_2 on- and off-rates of both proteins. Both pairs of rates are identical within experimental error (Table II), but faster than what is observed for CO. Likewise, both pentacoordinate binding constants, $K_{d,O2_pent}$, are within error of each other, and the difference in $K_{A,His}$ is a factor of three. Thus loop addition has no measurable effect on the binding of either ligand.

Oxyferrous state lifetime. Under standard conditions (1 µM heme-protein complex, 21% O_2) FHFH has an oxyferrous state lifetime more than threefold longer than HP7-H7F (Figure 6A). Figures 6B depict the O_2 dependence of the lifetime. Fitted O_2 dissociation constants agree well with those determined using the ratio of on and off rates (see Table III). FHFH has an oxidation rate two-fold slower than that of HP7-H7F.

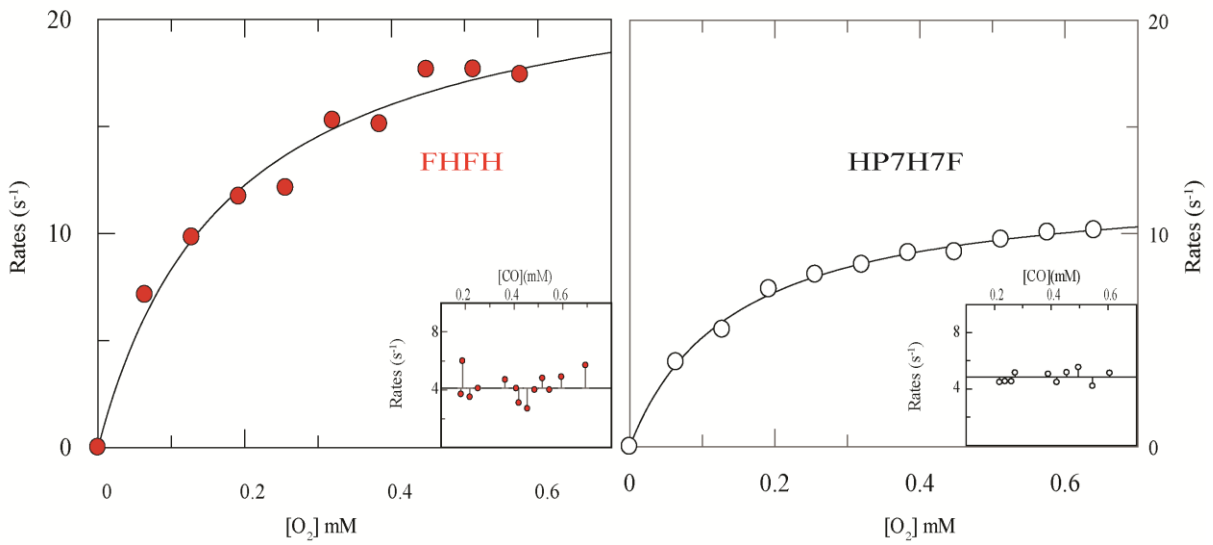

Figure 5. Kinetic analysis of O_2 binding to FHFH (red circles) and HP7-H7F (white circles). On-rates as a function of oxygen concentration. Reduced protein−heme complex (20 μM), prepared by careful titration with a slight excess of dithionite, was mixed with oxygen in a stopped-flow apparatus, and the rate of oxyferrous state formation was followed spectroscopically. Lines shown are fits to the data with eq 1. (Insets) Off rates as a function of carbon monoxide concentration. Lines are the average rates of oxygen displacement.

Discussion

The identical pentacoordinate binding constants of both ligands to each protein in concert with the indistinguishable spectra of the bound states suggests that the bound ligand environments are identical in both proteins. Therefore the addition of the connecting loop significantly reduces water access, and thus the internal dielectric, of the protein core without measurably altering the heme binding site. As one would expect, this has a small effect on the binding of the heme cofactor but no measurable effect on the binding of uncharged gaseous ligands. It does have a significant effect on protein function – the lifetime of the oxyferrous state increases by a factor of more than three. Two parameters play a role in oxyferrous state lifetimes, the O_2 binding constant and the second order rate constant for the oxidation of ferrous heme by unbound

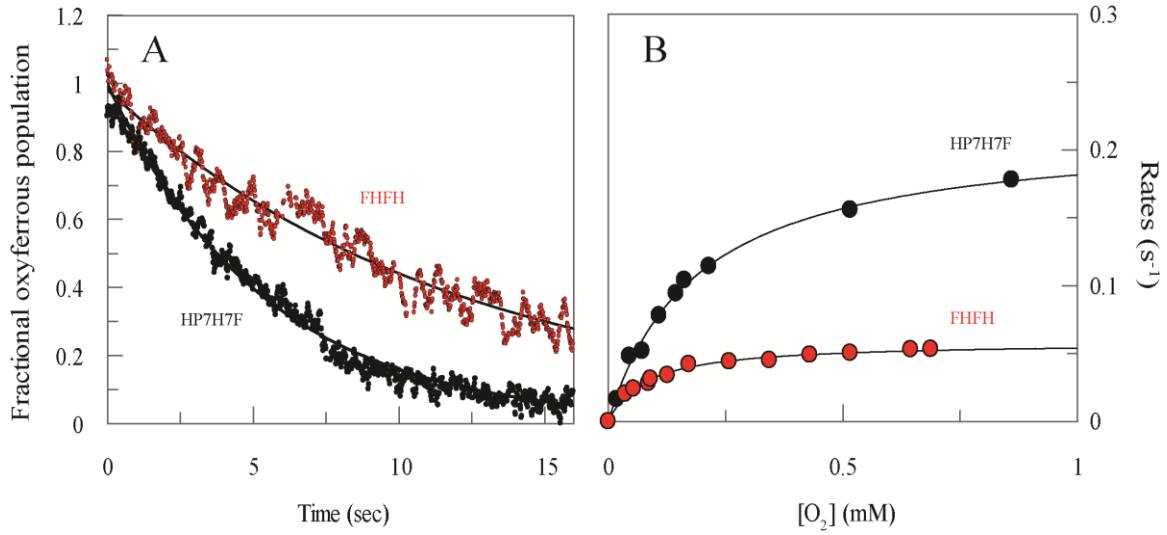

Figure 6. Rates of oxyferrous state breakdown in FHFH (red circles) and HP7-H7F (black circles). Reduced protein−heme complex (1 μM), prepared by careful titration with a slight excess of dithionite, was mixed with oxygenated buffer and the rate of protein oxidation followed spectroscopically. (A) Direct comparison of the breakdown of 1 μM oxyferrous state in air (21% saturated O_2). Lines are fits with a single exponential function with rates of 0.041 s^{-1} and 0.13 s^{-1}. (B) Oxygen concentration dependence of the oxyferrous state breakdown of FHFH (red circles) and HP7-H7F (black circles) respectively.

oxygen.

The slower oxidation rate, which correlates with a decrease in the rate of water penetration, is the principal basis for the increase in oxyferrous state lifetime. The distal histidine dissociation rate, which is also altered by the loop addition, will not affect the lifetime as both proteins have oxygen on and off-rates that are more than an order of magnitude faster than the histidine on-rate. Therefore once oxygen is bound, oxidation of the reduced heme occurs only after oxygen transiently dissociates, allowing the pentacoordinate heme to be oxidized by a molecule of oxygen (49). The increased distal histidine off-rate will in theory result in an increased yield of oxyferrous complex upon mixing ferrous protein with oxygen, but HP7-H7F already quantitatively forms the oxyferrous state within experimental error. Furthermore the lower reduction potential of FHFH should serve to destabilize the oxyferrous state, making oxidation of the heme more favorable. Therefore, it is unlikely to be responsible

for the increase in lifetime and we attribute the lengthened lifetime directly to the decreased water penetration into the protein core. To our knowledge, this the first direct demonstration that dynamically-driven water penetration can be a rate-limiting step in the oxidation of protein-heme complexes.

This relationship between protein dynamics, water penetration, and the stability of metalloprotein intermediate states is instructive as to the reasons for the failure of earlier designed heme proteins to stably bind oxygen: early minimalist peptide-protein complexes directly expose the heme to solvent (50,51). The first designed heme proteins large enough to entirely occlude the heme surface were composed of helical pairs, crosslinked by disulfide bond formation between terminal cysteine residues, which tetramerized via hydrophobic sequestration (52-54). However, these proteins are molten globules with the attendant rapid water penetration. The next generation of designed proteins formed using this topology are uniquely structured in the oxidized state (25,55), but the disulfide bonds which hold the fold together have reduction potentials higher than the bound heme cofactors (56). This results in the formation of a molten globular state upon heme reduction and the concomitant reduction of the protein disulfides, and explains why these reduced proteins cannot stably bind oxygen. The candelabra fold, which is the scaffold for HP7, consists of two helix-loop-helix chains which bind two hemes at the interface between the two monomers (23). The effect of this is to add sufficient rigidity to restrict water penetration such that there is measurable amide proton hydrogen exchange protection (4) and stable formation of an oxyferrous state. The increased rigidity imparted by the addition of the connecting loop in FHFH further restricts water penetration and thus further increases the stability and lifetime of the oxyferrous state. We note that another designed helical heme protein, C45, displays an increased catalytic activity when the helix stabilizing agent

trifluoroethanol is added (57). This enzyme does not bind molecular oxygen, rather the ferric heme reacts directly with hydrogen peroxide, creating compound I, another high energy intermediate susceptible to reaction with water (58).

In conclusion, these data are not only instructive as to the engineering requirements for natural oxygen transport and oxygen activation proteins; they make it clear that an important design principle in the construction of functional metalloproteins and metalloenzymes is the restriction of water from the active site.

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Table I. Heme binding parameters for two artificial heme proteins

Protein	$K^{*}_{d\text{-}red}$ (μM)	$E^{\dagger}_{mid}$ (mV)	$K^{\ddagger}_{d\text{-}ox}$ (nM)	k_{+his} (s^{-1})	k_{-his} (s^{-1})	$K_{A,his}$[§]	K_{d_pen}[‖] (μM)	pKa[¶]
HP7-H7F	0.6±0.2	-260±6	0.3±0.1	160±20	5.6±0.3	29±5	18±9	7.3±0.2
FHFH	3.4 ±0.5	-277±1	0.9±0.1	210±40	24±1	9±2	34±8	5.9±0.2

[*] hexacoordinate dissociation constant in the reduced state
[†] midpoint reduction potential vs NHE
[‡] ferric dissociation constant calculated using thermodynamic cycles (see ref 4)
[§] association constant for distal histidine ligation
[‖] ferrous pentacoordinate dissociation constant
[¶] apoprotein ligand histidine pK_a

Table II Gaseous ligand on- and off-rates and equilibrium constants at pH 9

Heme Protein	Gas ligand type	k_{+gas}[*] ($mM^{-1}s^{-1}$)	k_{-gas}[†] (s^{-1})	$K_{d\ gas_pen}$[‡] (μM)	$K_{d,actual}$[¶] (μM)	k_{+his} (s^{-1})	k_{-his} (s^{-1})	$K_{A,his}$[§]
HP7-H7F[↕]	CO	220±10	0.067±0.002	0.30±0.01	9±2	160±20	5.6±0.3	29±5
	O_2	1200±200	4.8±0.4	3.9±0.9	120±50			
FHFH	CO	300±90	0.071±0.007	0.24±0.07	2±1	210±40	24±1	9±2
	O_2	1200±300	4.1±0.3	3±1	40±20			

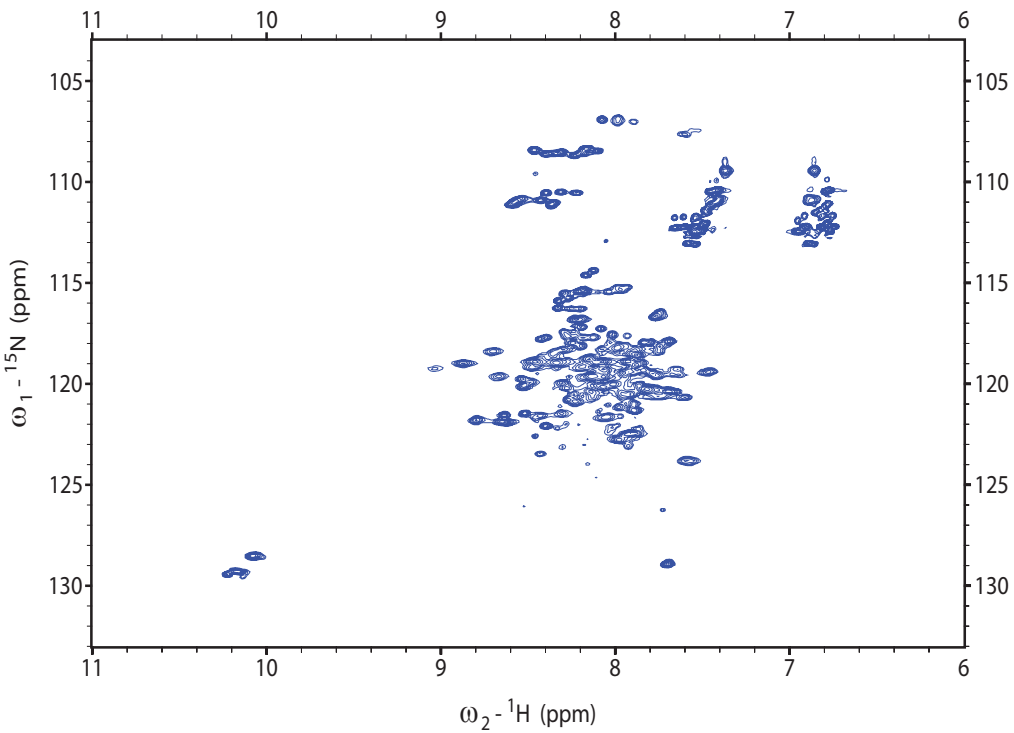
[*]On-rate for gaseous ligand binding to the pentacoordinate heme
[†]Rate of gaseous ligand dissociation
[‡]Gaseous ligand dissociation constant to the pentacoordinate heme
[¶]Gaseous ligand heme dissociation constant
[§]Association constant for distal histidine ligation
[↕]Data for HP7-H7F taken from reference 4.

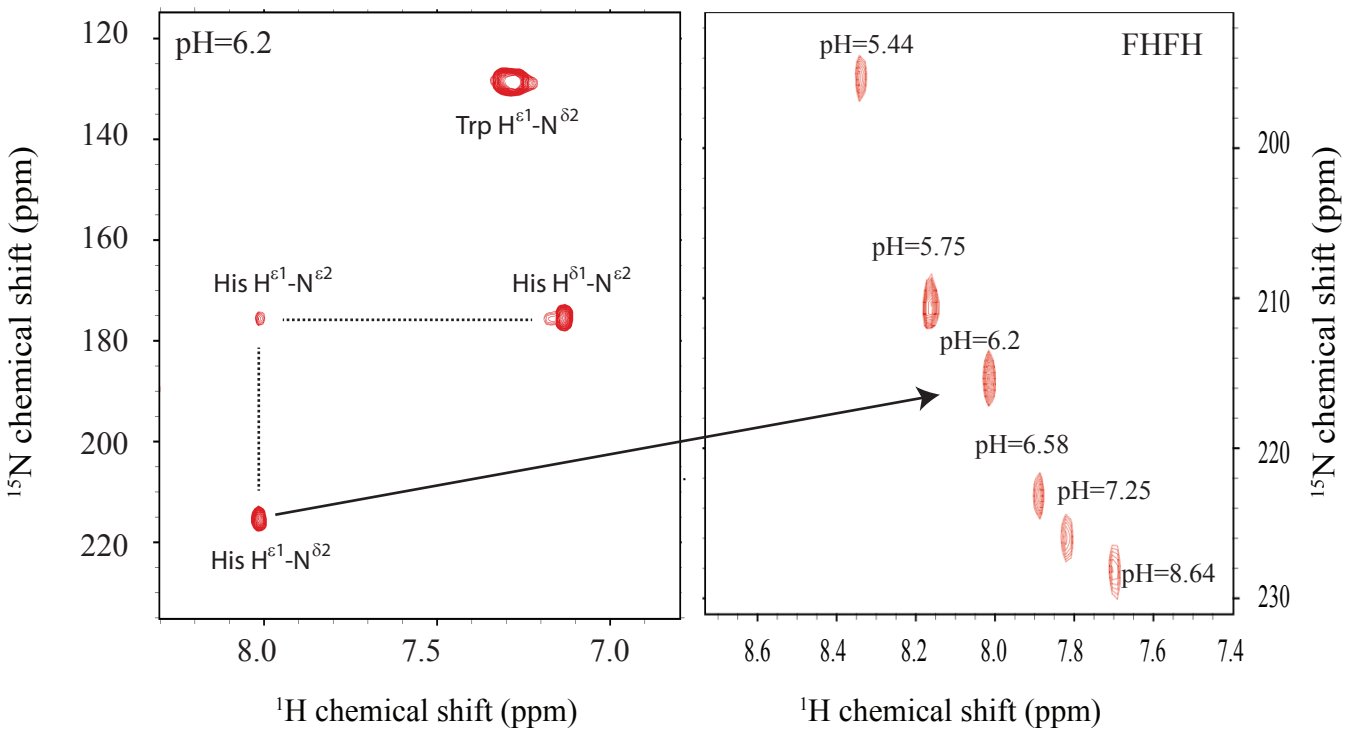
Table III. Oxyferrous state lifetimes

protein	$K_{d,O2}$ (μM)	k_{ox} ($s^{-1}mM^{-2}$)	R_{air} (s^{-1})[†]	$t_{1/2}$ (s)[*]
HP7-H7F	150±30	1200±40	0.13±0.01	5.2±0.7
FHFH	80±10	650±70	0.041±0.009	17±4

[†]Rate of oxyferrous state oxidation in air (21% O_2)
[*]Oxyferrous state half-life in air (21% O_2) with 1μM protein complex.

Supplementary Figure 1.





Supplementary Figure 2. pH Titration of histidine pK_as. (A) Multiple bond correlation spectrum of FHFH at pH 6.2 (B) Overlay of multiple-bond correlated $H^{\varepsilon1}$-$N^{\delta2}$ signals as a function of pH. The ^{1}H chemical shifts of this resonance were used to generate the plot in Figure 2A. Fits of the ^{15}N shift data give the same result, within error.