# Quantifying Amide-Aromatic Interactions at Molecular and Atomic Levels: Experimentally-determined Enthalpic and Entropic Contributions to Interactions of Amide $sp^2$O, N, C and $sp^3$C Unified Atoms with Naphthalene $sp^2$C Atoms in Water

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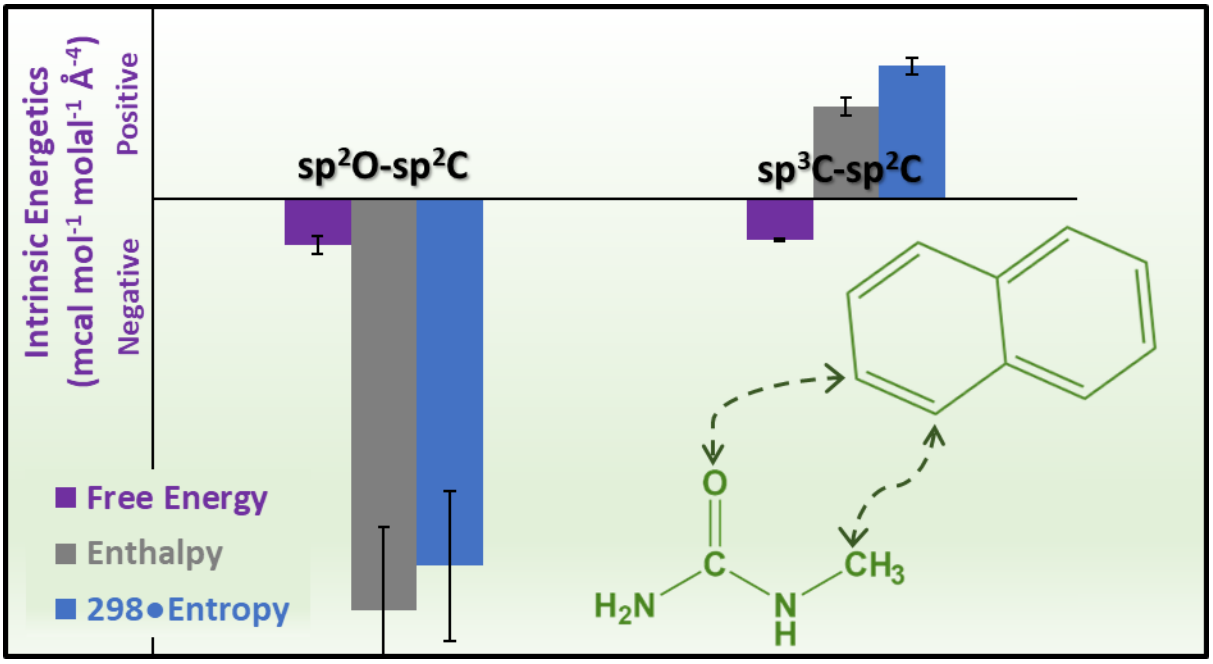
Table of Contents Graphic



**Abstract**

In addition to amide hydrogen bonds and the hydrophobic effect, interactions involving π-bonded $sp^2$ atoms of amides, aromatics and other groups occur in protein self-assembly processes including folding, oligomerization and condensate formation. These interactions also occur in aqueous solutions of amide and aromatic compounds, where they can be quantified. Previous analysis of thermodynamic coefficients quantifying net-favorable interactions of amide compounds with other amides and aromatics revealed that interactions of amide $sp^2O$ with amide $sp^2N$ unified atoms (presumably C=O···H-N hydrogen bonds) and amide/aromatic $sp^2C$ (lone pair-π, n-π*) are particularly favorable. $Sp^3C$-$sp^3C$ (hydrophobic), $sp^3C$-$sp^2C$ (hydrophobic, CH-π), $sp^2C$-$sp^2C$ (hydrophobic, π-π) and $sp^3C$-$sp^2N$ interactions are favorable, $sp^2C$-$sp^2N$ interactions are neutral, while $sp^2O$-$sp^2O$ and $sp^2N$-$sp^2N$ self-interactions and $sp^2O$-$sp^3C$ interactions are unfavorable. Here, from determinations of favorable effects of fourteen amides on naphthalene solubility at 10, 25 and 45 °C, we dissect amide-aromatic interaction free energies into enthalpic and entropic contributions and find these vary systematically with amide composition. Analysis of these results yields enthalpic and entropic contributions to intrinsic strengths of interactions of amide $sp^2O$, $sp^2N$, $sp^2C$ and $sp^3C$ unified atoms with aromatic $sp^2C$ atoms. For each interaction, enthalpic and entropic contributions have the same sign and are much larger in magnitude than the interaction free energy itself. The amide $sp^2O$-aromatic $sp^2C$ interaction is enthalpy-driven and entropically unfavorable, consistent with direct chemical interaction (e.g. lone pair-π) while amide $sp^3C$- and $sp^2C$-aromatic $sp^2C$ interactions are entropy-driven and enthalpically unfavorable, consistent with hydrophobic effects. These findings are relevant for interactions involving π-bonded $sp^2$ atoms in protein processes.

## Introduction

Self-assembly and binding interactions of biopolymers involve large numbers of relatively weak favorable (and unfavorable) interactions between individual atoms and/or groups. Because these individual interactions are comparable in strength to interactions of the atoms or groups with water, they have been difficult to quantify, resulting in ambiguity with regard to their contributions to the thermodynamics of forming biopolymer complexes and assemblies. These weak interactions also determine the effects of stabilizing or destabilizing biochemical solutes like glycine betaine or urea on biopolymer processes, and the strengths of interaction of these small solutes with one another in aqueous solution, where they are responsible for the deviations from a random solute mixture that are quantified as solute nonideality. From measurements of the effect of changing the concentration of one solute on the activity coefficient of a second solute, we determine information about the solute-solute interaction free energy. The quantity obtained is the chemical potential derivative $(\partial\mu_2/\partial m_3)_{T,P,m2} = \mu_{23}$, a model-independent fundamental thermodynamic coefficient which can be interpreted as a transfer free energy and which quantifies the free energy of the preferential interaction of the two solutes, relative to their interactions with water.[1-3]

For interactions of two small solutes, solubility or osmometry assays are used to determine $\mu_{23}$.[3-14] Solubility assays are used in the present work because one solute (naphthalene) is only sparingly soluble and because data as a function of temperature are needed. Where both solutes are sufficently soluble, osmometry is the assay of choice but only a few temperatures are readily available (e.g. 0 °C by freezing point osmometry; room temperature by vapor pressure osmometry). Analysis of these $\mu_{23}$ values demonstrates that interactions of individual groups or types of unified atoms on the surfaces of one or both of the solutes make additive contributions to $\mu_{23}$.[2,3,7-14] Chemically-reasonable interaction parameters (called α-values) are obtained that quantify these contributions. An earlier approach determined

the effect of urea and other solutes on the solubility of amino acids and dipeptides, and assumed additivity to obtain contributions of the peptide backbone and the nineteen amino acid side chains to a transfer free energy.[4-6] The two approaches have been compared, and advantages of the approach used here have been discussed.[3]

Recent research determined preferential interaction free energies ($\mu_{23}$ values) for 105 amide-amide and amide-aromatic interactions[12,14] and analyzed them as additive contributions from interactions of aliphatic $sp^3C$ and amide $sp^2O$, N and C unified atoms of the amide compound with one another and with the $sp^2C$ of the aromatic compound (naphthalene, anthracene).[14] From this analysis we obtained ten intrinsic interaction strengths (two-way $\alpha$-values) for the contributions to $\mu_{23}$ from each pairwise interaction between these unified atoms, expressed per unit of water-accessible surface area (ASA). Intrinsic strengths of interactions of aromatic $sp^2C$ and amide $sp^2C$ with other atom types were found to be sufficiently similar that they could be analyzed together.[12,14]

Interactions of amide $sp^2O$ with amide $sp^2N$ unified atoms ($sp^2O$-$sp^2N$; presumably C=O···H-N hydrogen bonds[15-19]) and with amide/aromatic $sp^2C$ ($sp^2O$-$sp^2C$; presumably a lone pair-π[20-22] or n-π* [18,23-25] interaction) are found to be particularly favorable.[14] All pairwise C-C interactions investigated are favorable. These include interactions between $sp^3C$ unified atoms of two amide compounds ($sp^3C$-$sp^3C$), interactions between $sp^3C$ atoms of amide compounds and amide or aromatic $sp^2C$ unified atoms ($sp^3C$-$sp^2C$) and interactions of amide $sp^2C$ with aromatic $sp^2C$ ($sp^2C$-$sp^2C$).[14] $Sp^3C$-$sp^3C$ interactions are presumably hydrophobic.[26-30] $Sp^3C$-$sp^2C$ and $sp^2C$-$sp^2C$ interactions may also involve the π system(s) (e.g. CH-π[31-35] and π-π[36-38] interactions).

Interactions between amide $sp^2N$ and $sp^3C$ unified atoms of amide compounds ($sp^2N$-$sp^3C$) are modestly favorable while amide $sp^2N$-$sp^2N$ self-interactions are unfavorable,[14] consistent with the observation that amide $sp^2N$-$sp^2N$ hydrogen bonds are uncommon in

proteins.[39] Self-interactions between amide $sp^2O$ unified atoms ($sp^2O$-$sp^2O$) are particularly unfavorable, as are $sp^2O$-$sp^3C$ interactions.[14] These unfavorable interactions of amide $sp^2O$ unified atoms with other amide $sp^2O$ unified atoms and with $sp^3C$ unified atoms of amide compounds indicate the strong preference of amide $sp^2O$ to interact with water than with itself or $sp^3C$. Likewise amide $sp^2N$ unified atoms evidently prefer to interact with water than with themselves.

From these $\alpha$-values and ASA information, one can predict the interaction of any amide solute with another amide or aromatic solute. Because the protein surface exposed in unfolding is composed mostly of these hydrocarbon and amide unified atoms, effects of any amide solute on the stability of a protein to unfolding can also be predicted from information about the amount and composition of the $\Delta$ASA of unfolding.[14]

Here solubility assays at 10, 25 and 45 °C are used to determine effects of temperature on free energies of preferential interactions ($\mu_{23}$ values) of fourteen amide compounds with naphthalene. Global analysis of these data for each amide compound yields $s_{23}$, the entropic contribution to $\mu_{23}$, which is the negative of the derivative of $\mu_{23}$ with respect to temperature. From $\mu_{23}$ and $s_{23}$ we obtain enthalpic ($h_{23}$) contributions to each amide-naphthalene preferential interaction ($h_{23} = \mu_{23} + Ts_{23}$). We interpret these results to obtain the four two-way $\alpha$-values and their entropic and enthalpic components for interactions of the $sp^2O$, $sp^2N$, $sp^2C$ and $sp^3C$ unified atoms of amide compounds with aromatic $sp^2C$ atoms. These $\alpha$-values make good chemical sense and, moreover, are useful to predict the $\mu_{23}$ value and its entropic ($s_{23}$) and enthalpic ($h_{23}$) components for any amide-aromatic interaction.

To examine whether an atom-atom (four-parameter) analysis of the free energy, enthalpy and entropy of amide-naphthalene interactions is more justifiable than alternatives, we also interpret these data using two interaction parameters, one quantifying amide π-aromatic π interactions, the other for amide $sp^3C$-aromatic π (i.e. CH–π) interactions. We compare the

statistical and chemical significance of free energy $\alpha$-values and their enthalpic and entropic components obtained using these two analyses of amide-naphthalene interactions, and conclude that the atom-atom (four-parameter) analysis is statistically better and yields interaction parameters (two-way $\alpha$-values) that are more chemically meaningful.

Combining the 25 °C amide-naphthalene data obtained here with that reported previously for amide-amide and amide-aromatic interactions,[14] we analyze this large (109 member) data set to refine the ten two-way $\alpha$-values obtained previously that quantify the interactions of amide $sp^2O$, amide $sp^2N$, aliphatic $sp^3C$ and amide/aromatic $sp^2C$ atoms. Additionally, as a test of whether an atom-atom (ten-parameter) analysis of the free energy of amide-amide and amide-aromatic interactions is more justifiable than alternatives, we also analyze the 109-member set of amide-amide and amide-aromatic $\mu_{23}$ values as π-π and CH-π interactions. This analyses requires only four two-way $\alpha$-values, less than half the number needed for an atom-atom analysis (ten two-way $\alpha$-values). Although π-π and CH-π $\alpha$-values obtained from this fit provide moderately good predictions of $\mu_{23}$ values from ASA information, neither the statistics of the fit nor the chemical significance of the parameters is of the same quality as when interpreted using ten atom-atom interactions.

## Background on Determinations of Molecular and Atom-Atom Amide-Naphthalene Interaction Free Energies; Extension to Interaction Entropies and Enthalpies

### *Solubility Determinations of Naphthalene Transfer Free Energies from Water to Amide Solutions ($\Delta G^o_{obs,tr}$) and their Entropic ($\Delta S^o_{obs,tr}$) and Enthalpic ($\Delta H^o_{obs,tr}$) Components*

Early research investigated interactions of amino acids (solute species 2) with urea or osmolytes (solute species 3) by solubility measurements. In these studies the effect of solute 3

on the solubility of solute 2 was determined and used to calculate an observed standard free energy of transfer ($\Delta G^{o}_{obs,tr}$) of solute 2 from water to a specified concentration of solute 3:[4-6]

$$\Delta G^{o}_{obs,tr} = - RT \ln(m_2^{ss,m3}/m_2^{ss,0}) \qquad \text{Eq. 1}$$

where $m_2^{ss,m3}$ is the molal scale solubility of solute 2 at molality $m_3$ of solute 3 and $m_2^{ss,0}$ is the solubility in absence of solute 3. In the studies reported here naphthalene is solute 2 and a set of 14 amides, including urea, are investigated as solute 3. Values of $\Delta G^{o}_{obs,tr}$ determined as a function of amide concentration $m_3$ provide information about the change in the activity coefficient of naphthalene as a function of $m_3$, and therefore about the amide-naphthalene preferential interaction free energy $\mu_{23}$ (see SI Eqs. S1-3).

The entropy of transfer $\Delta S^{o}_{obs,tr}$ is obtained from the temperature dependence of $\Delta G^{o}_{obs,tr}$

$$\Delta S^{o}_{obs,tr} = - (\partial \Delta G^{o}_{obs,tr} / \partial T)_P \qquad \text{Eq. 2}$$

and the enthalpy of transfer $\Delta H^{o}_{obs,tr} = \Delta G^{o}_{obs,tr} + T\Delta S^{o}_{obs,tr}$. Eq. 3

No systematic studies of $\Delta H^{o}_{obs,tr}$ or $\Delta S^{o}_{obs,tr}$ have been reported previously.

*Free Energies of Amide-Naphthalene Preferential Interaction ($\mu_{23}$) and their Entropic ($s_{23}$) and Enthalpic ($h_{23}$) Components, Determined from Transfer Free Energies*

The observed transfer free energy $\Delta G^{o}_{obs,tr}$ depends on the concentration of solute 3. Expanding $\Delta G^{o}_{obs,tr}$ in a Taylor-Maclaurin series in powers of $m_3$ yields:

$$\Delta G^{o}_{obs,tr} = \mu_{23}^{m3=0}\, m_3 + 0.5\, \mu_{233}^{m3=0} m_3^2 + \ldots\ldots \qquad \text{Eq. 4}$$

In Eq. 4, $\mu_{23}$ and $\mu_{233}$ are defined as

$$\mu_{23} = (\partial \mu_2 / \partial m_3)_{T,P,m_2} = \left(\partial \Delta G^{o}_{obs,tr} / \partial m_3\right)_{T,P,m_2} \qquad \text{Eq. 5}$$

$$\mu_{233} = (\partial^2 \mu_2 / \partial m_3^2)_{T,P,m_2} = (\partial^2 \Delta G^{o}_{obs,tr} / \partial m_3^2)_{T,P,m_2} \qquad \text{Eq. 6}$$

(see SI Eqs. S2-3), and $\mu_{23}^{m3=0}$ and $\mu_{233}^{m3=0}$ are limiting values of $\mu_{23}$ and $\mu_{233}$ as $m_3$ approaches 0.

The chemical potential derivative $\mu_{23}$ quantifies the preferential interaction of the amide compound with naphthalene. Since $\mu_{23}$ is the initial slope of a plot of $\Delta G^{o}_{obs,tr}$ vs. $m_3$, it is the

solubility *m*-value. Since all $\mu_{23}$ and $\mu_{233}$ obtained here are $\mu_{23}^{m3=0}$ and $\mu_{233}^{m3=0}$ we omit this superscript hereafter. Values of $\mu_{23}$ at temperature T are related to those at 298 K by

$$\mu_{23}^{T} = \mu_{23}^{298K} - °C\, s_{23}\,(T - 298\ K) \qquad \text{Eq. 7}$$

where $s_{23} = -(\partial\mu_{23}/\partial T)_P$ is the entropy of the naphthalene-amide preferential interaction, assumed independent of T. Since $\mu_{23}$ refers to $\mu_{23}^{m3\rightarrow 0}$, similarly $s_{23}$ is the entropy of the preferential interaction in the limit as $m_3 \rightarrow 0$.

From Eqs. 4 and 7,

$$\Delta G^{o}_{obs,tr} = \mu_{23}^{298K} m_3 - s_{23}\, m_3 (T - 298) + 0.5\, \mu_{233} m_3^2 \qquad \text{Eq. 8}$$

Naphthalene solubility data as a function of concentration of each amide compound at 10, 25 and 45 °C are fit globally to Eq. 8 to obtain $\mu_{23}^{298K}$, $s_{23}$ and $\mu_{233}$, where $\mu_{233}$ is also assumed to be independent of T. The enthalpy of preferential interaction (also assumed independent of T) is obtained from $\mu_{23}^{298}$ and $s_{23}$.

$$h_{23} = \mu_{23}^{298K} + 298\, s_{23} \qquad \text{Eq. 9}$$

Alternatively, data for $\Delta G^{o}_{obs,tr}/T$ as a function of $m_3$ and T may be analyzed globally using Eq. S4 (paralleling Eq. 8) to obtain $h_{23}$ directly. Values of $h_{23}$, $s_{23}$ and $\mu_{23}^{298K}$ obtained from Eq. S4 and Eq. 9 are the same within uncertainty as those obtained from Eqs. 8-9.

*Amide Atom-Naphthalene Atom Contributions ($\alpha_{ij}$-values) to Preferential Interaction Free Energies; Proposed Extension to Preferential Interaction Enthalpies and Entropies*

Values of $\mu_{23}$ for amide-naphthalene interactions are dissected into interactions of naphthalene and its $sp^2C$ atoms with the four types of amide atoms (amide $sp^2O$, $sp^2N$ and $sp^2C$; methyl or methylene $sp^3C$), listed by the index i ($1 \leq i \leq 4$), using the relationship[14]

$$\mu_{23} = \sum_i \alpha_{i,naph}^{\mu 23} ASA_i^{amide} = (\sum_i \alpha_{i,sp2C}^{\mu 23} ASA_i^{amide}) ASA_{sp2C}^{naph} \qquad \text{Eq. 10}$$

where $\alpha_{i,naph}^{\mu 23}$ and $\alpha_{i,sp2C}^{\mu 23}$ are the intrinsic strengths of interaction of amide atom type i with naphthalene and with its $sp^2C$ atoms, both expressed per unit of water-accessible surface area (ASA) of amide atom i (cf. Table S1 for ASA values). In Eq. 10, $\alpha_{i,sp2C}^{\mu 23} = \alpha_{i,naph}^{\mu 23}/273$ where the $ASA_{sp2C}^{naph}$ is 273 Å$^2$. The assumptions of additivity in the one-way and two-way analyses of Eq. 10 have been tested and justified previously[3,7-14] and are also tested here, because the size of the experimental data set (14 $\mu_{23}$ values) exceeds the number of fitting parameters (four different $\alpha_{i,sp2C}^{\mu 23}$).

Here (in Discussion) we apply the two way analysis of Eq. 10 to amide-aromatic intereactions and extend this analysis to $h_{23}$ and $s_{23}$. We introduce and test dissections of $s_{23}$ and $h_{23}$ based on additivity and ASA that are analogous to Eq. 10 for $\mu_{23}^{298K}$. From this analysis we obtain enthalpic ($\alpha_{i,sp2C}^{h23}$) and entropic (298 $\alpha_{i,sp2C}^{s23}$) contributions to free energy $\alpha$-values ($\alpha_{i,sp2C}^{\mu 23}$) from interactions of naphthalene $sp^2C$ unified atoms with amide $sp^2O$, $sp^2N$ and $sp^2C$ and methyl or methylene $sp^3C$ atoms of the amide compounds. We also use a similar approach to examine whether interactions with the amide group can be successfully modeled as interactions with the amide π system.

## Results

*Large Effect of Increasing Hydrocarbon (sp$^3$C) ASA of Amide Compounds on the Temperature-Dependence of the Free Energy ($\Delta G^o_{tr}$) of Transfer of Naphthalene from Water to Amide Solutions*

To quantify the thermodynamics of amide-naphthalene interactions, naphthlene solubilities were measured as a function of molal concentration $m_3$ of fourteen different amide compounds at three temperatures (10, 25, 45 °C). Naphthalene transfer free energies $\Delta G^o_{obs,tr}$ (Eq. 1) calculated from these solubilities are plotted vs $m_3$ in Figures 1 and S1.

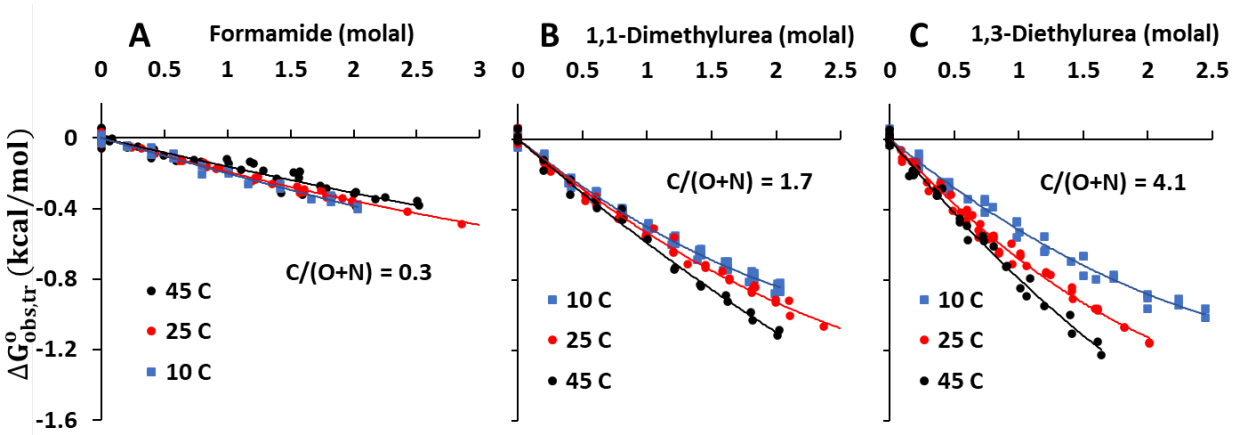



**Figure 1. Naphthalene transfer free energies ($\Delta G^o_{obs,tr}$) from water to amide solutions at 10, 25, and 45 °C: different temperature effects for different amide nonpolar (C) : polar (O, N) ASA ratios.** Values of $\Delta G^o_{obs,tr}$, obtained from solubilities (Eq. 1) at 10, 25, and 45 °C are plotted vs. molal concentration of three amides (panel A formamide, panel B 1,1-dimethylurea, and panel C 1,3-diethylurea) selected for their widely different ratios of nonpolar (C) to polar (O+N) surface area, indicated as C/(O+N) in the figure panels. Initial slopes of quadratic fits of the data of each panel to Eq. 4 (plotted curves) are $\mu_{23}$ values at 10, 25, and 45 °C (see Table 1). Analogous plots for 11 other amide compounds are given in Figure S1.

Results for three amide compounds (formamide, 1,1-dimethylurea (1,1-dmu) and 1,3-diethylurea (1,3-deu)), selected for their very different amounts of hydrocarbon (methyl and methylene sp$^3$C) and amide (sp$^2$O, sp$^2$N, sp$^2$C) surface area (ASA) and their different ratios of nonpolar (C) to polar (O+N) ASA, are shown in Figure 1. These are plotted using the same free energy and concentration scales so the magnitudes and temperature dependences of $\Delta G^o_{obs,tr}$

for the three amides can be visually compared. Corresponding plots for eleven other amide compounds, together with an expanded free energy scale plot for formamide, are given in the panels of Figure S1.

All fourteen amide compounds increase the solubility of naphthalene at all three temperatures, indicating that all of them interact favorably with naphthalene. Values of $\Delta G^{o}_{obs,tr}$ for transfer of naphthalene from water to these amide solutions are negative and increase with increasing amide concentration. All plots in Figures 1 and S1 are linear within uncertainty at amide concentrations below 0.5 molal. From Eq. 4-5 the initial slopes of these plots are free energies of amide-naphthalene preferential interactions ($\mu_{23}$ values). Except for the smallest amides (formamide, urea), curvature is observed at amide concentrations above 0.5 - 1.0 molal. This curvature, quantified by the derivative of $\mu_{23}$ with respect to $m_3$ ($\mu_{233}$; Eq. 6) generally increases with the size of the amide compound.

Formamide is the smallest amide compound, lacks alkyl ($sp^3C$) hydrocarbon, and has a relatively small C/(O+N) ASA ratio of 0.3 (Table 1). Figure 1 shows that the transfer free energy $\Delta G^{o}_{tr}$ of naphthalene from water to a specified formamide concentration is much smaller in magnitude (less favorable) and much less temperature-dependent than for the other amides in Figure 1. From Figure 1A (expanded in Figure S1H) the transfer free energy $\Delta G^{o}_{obs,tr}$ of naphthalene from water to a specified formamide concentration is slightly more favorable at 10 °C than at 25 °C, and least favorable at 45 °C. Because $\Delta G^{o}_{obs,tr}$ increases with increasing temperature, the corresponding entropy of transfer $\Delta S^{o}_{obs,tr}$ of naphthalene from water to formamide is negative (unfavorable) and the favorable $\Delta G^{o}_{obs,tr}$ therefore results entirely from a favorable (negative) enthalpy of transfer $\Delta H^{o}_{obs,tr}$.

Formamide is one of five amide compounds in Figure S1 for which naphthalene transfer free energies are only weakly temperature dependent. Of the other four (urea, acetamide, malonamide, methylformamide), all but methylformamide also have relatively small C/(O+N)

ASA ratios (Table 1). Unlike formamide, $\Delta G^{o}_{obs,tr}$ for each of these four amides, compared at any amide concentration, decreases (becomes slightly more negative) with increasing temperature, revealing that the corresponding entropies of transfer $\Delta S^{o}_{obs,tr}$ of naphthalene from water to these amides are positive (favorable). Values of $\Delta G^{o}_{obs,tr}$ for these four amides are 1.5- to 2-fold larger in magnitude than those of formamide at any amide concentration.

Alkylated ureas with $sp^3C$ (methyl, methylene) ASA have much larger C/(O+N) ASA ratios than formamide and, compared at the same concentration, are much better solubilizing agents for naphthalene than formamide. Two examples are shown in Figure 1B-C. Values for $\Delta G^{o}_{obs,tr}$ for transfer of naphthalene from water to 1,1-dmu (C/(O+N) = 1.7) and to 1,3-deu (C/(O+N) = 4.1) are three to four times more favorable than for formamide or urea at the same temperature and $m_3$. For these and other alkylureas in Figure S1, $\Delta G^{o}_{obs,tr}$ decreases strongly with increasing temperature at any fixed $m_3$, and the effect of temperature is larger for the larger alkylureas. Hence values of $\Delta S^{o}_{obs,tr}$ are positive and increase with increases in the C/(O+N) ASA ratio for these alkylureas.

Taken together, the results of Figures 1 and S1 indicate that the favorable $\Delta G^{o}_{obs,tr}$ of naphthalene from water to amide solutions is the result of favorable interactions of naphthalene with both the polar (O+N) and nonpolar (C) portions of these amides. The favorable contribution from the nonpolar (C) portion of the amide compound to naphthalene $\Delta G^{o}_{obs,tr}$ results from a very favorable contribution to $\Delta S^{o}_{obs,tr}$, partially countered by an unfavorable $\Delta H^{o}_{obs,tr}$. On the other hand, the origin of the favorable contribution from the polar (O+N) portion of the amide to naphthalene $\Delta G^{o}_{tr}$ results from a favorable contribution to $\Delta H^{o}_{obs,tr}$.

*Quantifying Free Energies of Amide Compound-Naphthalene Preferential Interactions ($\mu_{23}$) and their Entropic and Enthalpic Parts ($s_{23}$, $h_{23}$) from* $\Delta G^{o}_{obs,tr}$

Free energies ($\mu_{23}^{298K}$) and entropies ($s_{23}$) of the preferential interaction of each of fourteen amide compounds with naphthalene interactions are obtained from global fitting of naphthalene $\Delta G^{o}_{obs,tr}$ data (Figure 1 and S1) at 10, 25 and 45 °C to Eq. 8. Values of $\mu_{23}^{298K}$ and 298 $s_{23}$ are listed in Table 1, together with C/(O+N) ASA ratios for these amide compounds. Enthalpies ($h_{23}$) of the preferential interactions of each amide compound with naphthalene, calculated from $\mu_{23}^{298K}$ and 298 $s_{23}$ by Eq. 9, are also listed in Table 1.

**Table 1. Thermodynamics of Amide-Naphthalene Preferential Interactions**

| | | (kcal mol$^{-1}$ molal$^{-1}$) | | |
|---|---|---|---|---|
| **Solute** | **C/(O+N)** [a] | $\mu_{23}^{298K}$ [b] | **298 $s_{23}$** [b,c] | **$h_{23}$** [d] |
| urea | 0.04 | -0.15 ± 0.01 | 0.14 ± 0.06 | -0.01 ± 0.07 |
| methylurea | 0.75 | -0.37 ± 0.01 | 0.43 ± 0.14 | 0.06 ± 0.15 |
| ethylurea | 1.09 | -0.49 ± 0.01 | 0.56 ± 0.13 | 0.07 ± 0.14 |
| 1,1-dmu [e] | 1.67 | -0.61 ± 0.02 | 0.96 ± 0.10 | 0.35 ± 0.11 |
| 1,3-dmu [e] | 2.48 | -0.57 ± 0.01 | 1.08 ± 0.09 | 0.51 ± 0.10 |
| 1,1-deu [e] | 2.48 | -0.80 ± 0.02 | 2.24 ± 0.12 | 1.45 ± 0.14 |
| 1,3-deu [e] | 4.08 | -0.75 ± 0.02 | 2.29 ± 0.18 | 1.55 ± 0.20 |
| malonamide | 0.30 | -0.36 ± 0.02 | 0.22 ± 0.14 | -0.14 ± 0.16 |
| propionamide | 1.31 | -0.45 ± 0.02 | 0.55 ± 0.12 | 0.11 ± 0.14 |
| formamide | 0.33 | -0.20 ± 0.01 | -0.32 ± 0.09 | -0.52 ± 0.10 |
| methylformamide | 1.84 | -0.43 ± 0.01 | 0.07 ± 0.10 | -0.36 ± 0.11 |
| dimethylformamide | 4.44 | -0.67 ± 0.01 | 0.96 ± 0.13 | 0.30 ± 0.15 |
| acetamide | 0.88 | -0.31 ± 0.01 | 0.16 ± 0.14 | -0.15 ± 0.15 |
| methylacetamide | 3.34 | -0.63 ± 0.02 | 0.48 ± 0.14 | -0.15 ± 0.16 |

[a] C/(O+N) is the ratio of combined $sp^3C$ and $sp^2C$ ASA to combined $sp^2O$ and $sp^2N$ ASA of amide compounds. ASA information is listed in SI Table S1.
[b] Obtained by global fitting of data of panels of Figures 1 and S1 to Eq. 8. Uncertainties are reported as twice the fitting error.
[c] Assumed independent of temperature in range 10 - 45 °C.
[d] Calculated from Eq. 9; uncertainty propagated from uncertainties in $\mu_{23}^{298K}$ and 298 $s_{23}$.[11]
[e] Abbreviations: dmu, dimethylurea; deu, diethylurea

Alternatively, values of $\Delta G^{o}_{obs,tr}/T$ are analyzed as a function of 1/T using Eq. S4 to obtain $\mu_{23}^{298K}$ and $h_{23}$, from which $s_{23}$ is calculated by Eq. 9. Identical values of $\mu_{23}^{298K}$ for interactions of each of the 14 amides with naphthalene are obtained by fitting to either Eq. 8 or

S4. Similar but not identical values and uncertainties are obtained for $h_{23}$ and $s_{23}$ by these two approaches; these are compared in SI Tables S2 (for 298 $s_{23}$) and S3 (for $h_{23}$). Values of $\mu_{233}$ obtained from fitting the $\Delta G^{o}_{obs,tr}$ values to Eq. 8 or S4 are compared in Table S4. No significant differences in $\mu_{233}$ from these two fitting approaches are observed. For the alkylurea series, $\mu_{233}$ increases from approximately zero for urea to 0.19 for 1,3-diethylurea. Figure 2 and Table 1 group the fourteen amides into ureas, formamides, acetamides, and other amides, and list the compounds in each group in order of increasing C/(O+N) ASA ratio. Values of $\mu_{23}^{298K}$, $h_{23}$ and 298 $s_{23}$ for interactions of each amide with naphthalene are plotted on the bar graph in Figure 2, which shows three patterns in the thermodynamics of these amide-naphthalene preferential interactions.

i) Free energies ($\mu_{23}^{298K}$) of these interactions, always favorable ($\mu_{23}^{298K} < 0$), generally become more favorable with increasing C/(O+N) ASA ratio in each group. Minor exceptions are observed for the 1,3-dialkylureas, for which $\mu_{23}^{298K}$ values are somewhat less favorable than for the corresponding 1,1-dialkylureas, even though the C/(O+N) ASA ratios are larger for 1,3- than for 1,1-dialkylureas. Within each group of amides, the changes in the C/(O+N) ASA ratio directly reflect an increase in aliphatic $sp^3C$ ASA and a reduction in amide $sp^2N$ ASA, with smaller reductions in the ASA of amide $sp^2O$ and $sp^2C$ (Table S1). The trends in $\mu_{23}$ in all series are therefore consistent with our previous findings that interaction of amide $sp^3C$ with aromatic $sp^2C$ is favorable while interaction of amide $sp^2N$ with aromatic $sp^2C$ is marginally unfavorable.[14]

ii) For the series of ureas and formamides and the two acetamides in Figure 2, enthalpic ($h_{23}$) and entropic (298 $s_{23}$) contributions to the amide-naphthalene preferential interaction free energy $\mu_{23}^{298K}$ increase with increasing C/(O+N) ASA ratio. Enthalpies $h_{23}$ are zero or negative (favorable) for the member of each group with the smallest C/(O+N) ASA ratio (urea, formamide, methylacetamide) and increase monotonically with increasing C/(O+N) ASA ratio.

These interaction enthalpies are unfavorable for all alkylureas and for two other amides with moderately large C/(O+N) ratios (propionamide, dimethylformamide).

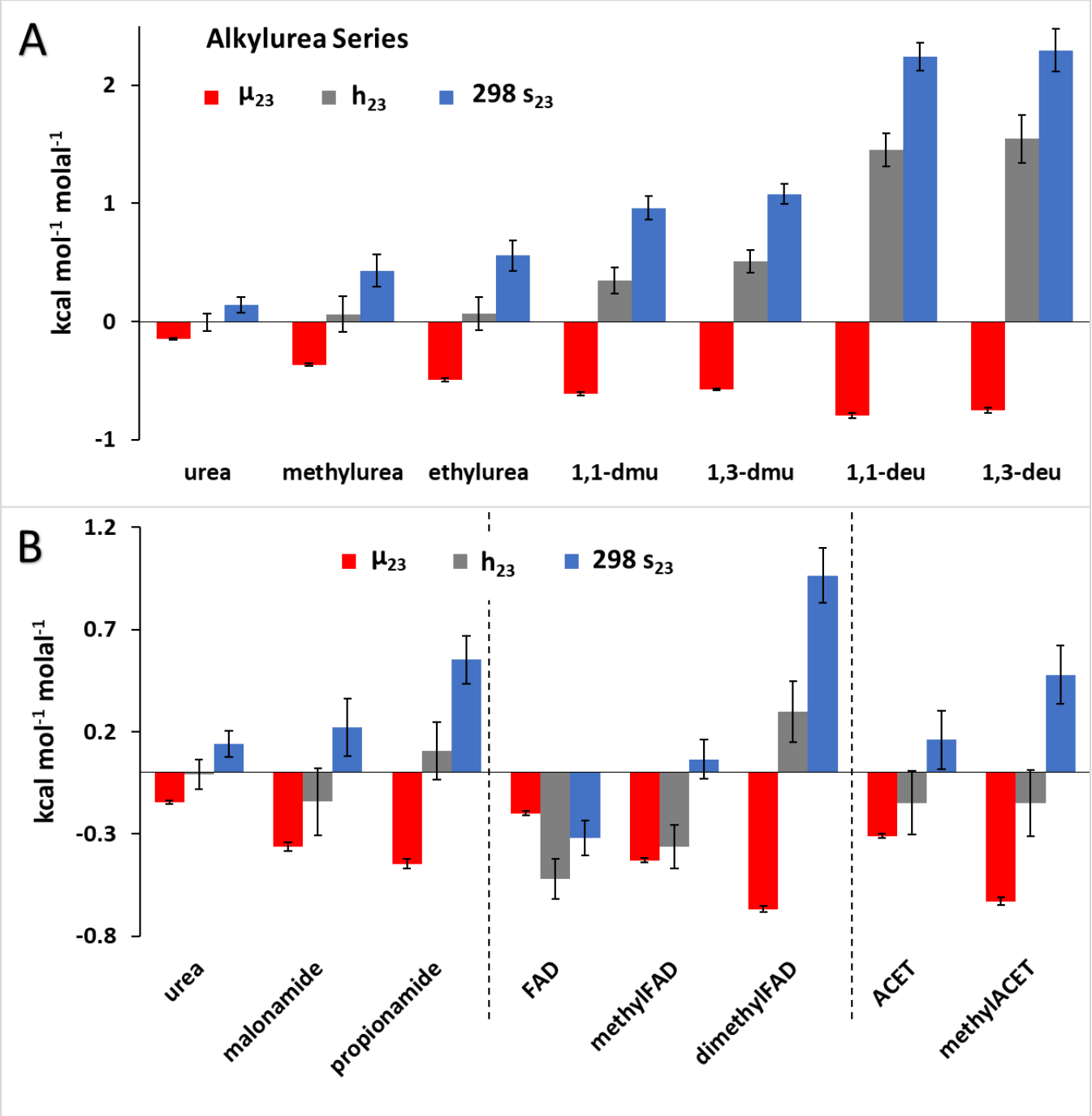



**Figure 2. Dissection of $\mu_{23}$ values for amide-naphthalene preferential interactions at 25 °C into enthalpic ($h_{23}$) and entropic (298 $s_{23}$) contributions.** Values of $\mu_{23}^{298K}$, $h_{23}$ and 298 $s_{23}$ are listed in Table 1). Panel A shows the series of alkylureas in order of increasing C/(O+N) ASA ratio. Panel B shows three different groups of amides, also ordered within each group by increasing C/(O+N) ASA ratio. Abbreviations: FAD, formamide; ACET, acetamide.

Figure 2 reveals several general trends characteristic of these groups of amides. Entropic contributions to $\mu_{23}^{298K}$ (i.e. 298 $s_{23}$) are favorable ($s_{23} > 0$) for all amides except formamide, and increases monotonically with increasing C/(O+N) ASA ratio. These positive 298 $s_{23}$ and the large increases in 298 $s_{23}$ with increasing C/(O+N) ASA ratio are consistent with a hydrophobic origin of the favorable interaction of amide $sp^3C$ with aromatic $sp^2C$ because (at least for the liquid hydrocarbon model) the hydrophobic effect is predominantly entropic near room temperature.[26-28]

iii) The favorable preferential interactions of all alkylureas with naphthalene at 25 °C ($\mu_{23}^{298K} < 0$) are entropy-driven (298 $s_{23} > h_{23} > 0$). (This situation also holds for alkylurea-naphthalene interactions at 10 and 45 °C.) Favorable preferential interactions between naphthalene and the most nonpolar amides in other groups in Figure 2 (dimethylformamide, propionamide, methylacetamide) are also entropy-driven with unfavorable (or modestly favorable) enthalpies. Favorable preferential interactions between naphthalene and the most polar amides in these groups (formamide, methylformamide, acetamide) are entirely or primarily enthalpy-driven, with an entropic contribution that is unfavorable for formamide and slightly favorable for methylformamide and acetamide.

## Analysis and Discussion

### *Atom-atom Interpretation of Amide - Naphthalene Interaction Free Energies*

As in a previous analysis of amide interactions,[14] we use Eq. 10 to interpret the fourteen amide-naphthalene interaction free energies at 25 °C ($\mu_{23}^{298K}$; Table 1) as sums of contributions from interactions of the $sp^2O$, $sp^2N$, $sp^2C$ and $sp^3C$ unified atoms displayed on each amide compound with naphthalene $sp^2C$ unified atoms. In this analysis, Eq. 10 is written for each individual amide-naphthalene $\mu_{23}^{298K}$, incorporating the appropriate ASA of $sp^2O$, $sp^2N$, $sp^2C$ and $sp^3C$ unified atoms of an amide compound (Table S1), the ASA of naphthalene $sp^2C$ (273 $Å^2$;

Table S1) and the four unknown amide atom-napthalene atom interaction strengths $\alpha_{i,sp2C}^{\mu 23}$. These $\alpha_{i,sp2C}^{\mu 23}$ are contributions to $\mu_{23}^{298K}$ from interactions of 1 Å$^2$ of amide compound atom-type i (where the index i represents sp$^2$O, sp$^2$N, sp$^2$C, or sp$^3$C unified atoms) with 1 Å$^2$ of naphthalene sp$^2$C ASA. Results are shown in Figure 3 and in the first column of Table 2. Very similar results are obtained from analysis of the combined set of 24 amide-aromatic (naphthalene, anthracene) $\mu_{23}^{298K}$ values using Eq. 10 (see Table S9 column 1).

**Table 2. Thermodynamics of amide atom–naphthalene sp$^2$C interactions: free energy α-values ($\alpha_{i,sp2C}^{\mu 23}$) and their enthalpic ($\alpha_{i,sp2C}^{h23}$) and entropic ($298\ \alpha_{i,sp2C}^{s23}$) components [a,b]**

| Amide atom | $\alpha_{i,sp2C}^{\mu 23}$ [c] | $\alpha_{i,sp2C}^{h23}$ [d] | $298\ \alpha_{i,sp2C}^{s23}$ [e] |
|---|---|---|---|
| **sp$^2$O** | -11.9 ± 2.3 | -108 ± 22 | -97 ± 20 |
| **sp$^2$N** | 0.7 ± 0.8 | 34 ± 8 | 33 ± 7 |
| **sp$^2$C** | -5.3 ± 1.8 | 25 ± 17 | 30 ± 15 |
| **sp$^3$C** | -10.6 ± 0.3 | 25 ± 2 | 35 ± 2 |
| **RMSE[f]** | 0.051 (in $\mu_{23}$) | 0.37 (in $h_{23}$) | 0.39 (in 298 $s_{23}$) |

[a] Units of $\alpha_{i,sp2C}^{\mu 23}$, $\alpha_{i,sp2C}^{h23}$ and 298 $\alpha_{i,sp2C}^{s23}$ are millical mol$^{-1}$ molal$^{-1}$ Å$^{-4}$
[b] Uncertainties are propagated as described previously.[11]
[c] Obtained from Eq. 10; values reported are at 25 °C.
[d] Obtained from Eq. 12; assumed independent of temperature
[e] Obtained from Eq. 11; assumed independent of temperature
[f] Root mean square errors (RMSE, calculated by Eq. S10) in comparisons of predicted and observed values of $\mu_{23}$ (see Figure 4B below) and of $h_{23}$ and 298 $s_{23}$ (see Figure 4C-D below) for the set of 14 amide-naphthalene interactions. Units of RMSE are kcal mol$^{-1}$ molal$^{-1}$.

This α-value analysis reveals that amide-naphthalene interactions are favorable ($\mu_{23}$ < 0; Figure 2) in the temperature range investigated because interactions of three of the four atom types of amide compounds with naphthalene sp$^2$C are favorable at these temperatures. Intrinsic strengths ($\alpha_{i,sp2C}^{\mu 23}$) of interactions of amide sp$^2$O and methyl/methylene sp$^3$C with naphthalene sp$^2$C are both very favorable, ranking with the amide sp$^2$N-amide sp$^2$O interaction (the amide NH···O=C hydrogen bond[14]) as the intrinsically most favorable amide interactions

investigated. We previously proposed that the very favorable interaction of amide $sp^2O$ with aromatic $sp^2C$ is a lone pair-π or n-π* interaction, and that the similarly-favorable interaction of amide $sp^3C$ with aromatic $sp^2C$ is a hydrophobic effect and/or a CH-π interaction. The amide $sp^2N$-aromatic $sp^2C$ interaction is marginally unfavorable at 25 °C, as previously determined.[14]

The amide $sp^2C$-aromatic $sp^2C$ interaction (a hydrophobic effect and/or π-π interaction) is favorable, but less so than amide $sp^2O$-aromatic $sp^2C$ and amide $sp^3C$-aromatic $sp^2C$ interactions. The amide $sp^2C$-aromatic $sp^2C$ α-value obtained here (cf. Table 2) is smaller in magnitude than that obtained previously for a combined data set of 85 amide-amide and 20 amide-aromatic $\mu_{23}$ value.[15] This interaction is more challenging to determine than others because of the small amounts of $sp^2C$ ASA for most amides. Incorporation of the amide-naphthalene data determined here into the previously-analyzed set of amide-amide and amide-aromatic $\mu_{23}$ values affects the value of $\alpha^{\mu 23}_{sp2C,sp2C}$. The newly-calculated $\alpha^{\mu 23}_{sp2C,sp2C}$ = -5.0 ± 3.2 millical $mol^{-1}$ $molal^{-1}$ $Å^{-4}$ for that 109-member set (see SI text and Table S9) agrees within uncertainty with that reported here for the 14-member amide-naphthalene set.

From fitted $\mu_{23}$ values at 10 and 45 °C (Table S5), free energy $\alpha^{\mu 23}_{i,sp2C}$ values for interactions of amide compound $sp^2O$, $sp^2N$, $sp^2C$, and $sp^3C$ atom-types with naphthalene $sp^2C$ were calculated at these temperatures using Eq. 10. These are summarized together with 25 °C $\alpha^{\mu 23}_{i,sp2C}$ values in the bar graphs of Figure 3A. The favorable interaction of amide $sp^2O$ with naphthalene $sp^2C$ becomes less favorable (less negative $\alpha^{\mu 23}_{sp2O,sp2C}$) with increasing temperature, demonstrating that the entropic contribution to the favorable $sp^2O$-$sp^2C$ interaction is negative and hence unfavorable. Since the free energy contribution of this amide $sp^2O$ - naphthalene $sp^2C$ interaction is favorable ($\alpha^{\mu 23}_{sp2O,sp2C} < 0$), its enthalpic component must be very favorable ($\alpha^{h23}_{sp2O,sp2C} \ll 0$).

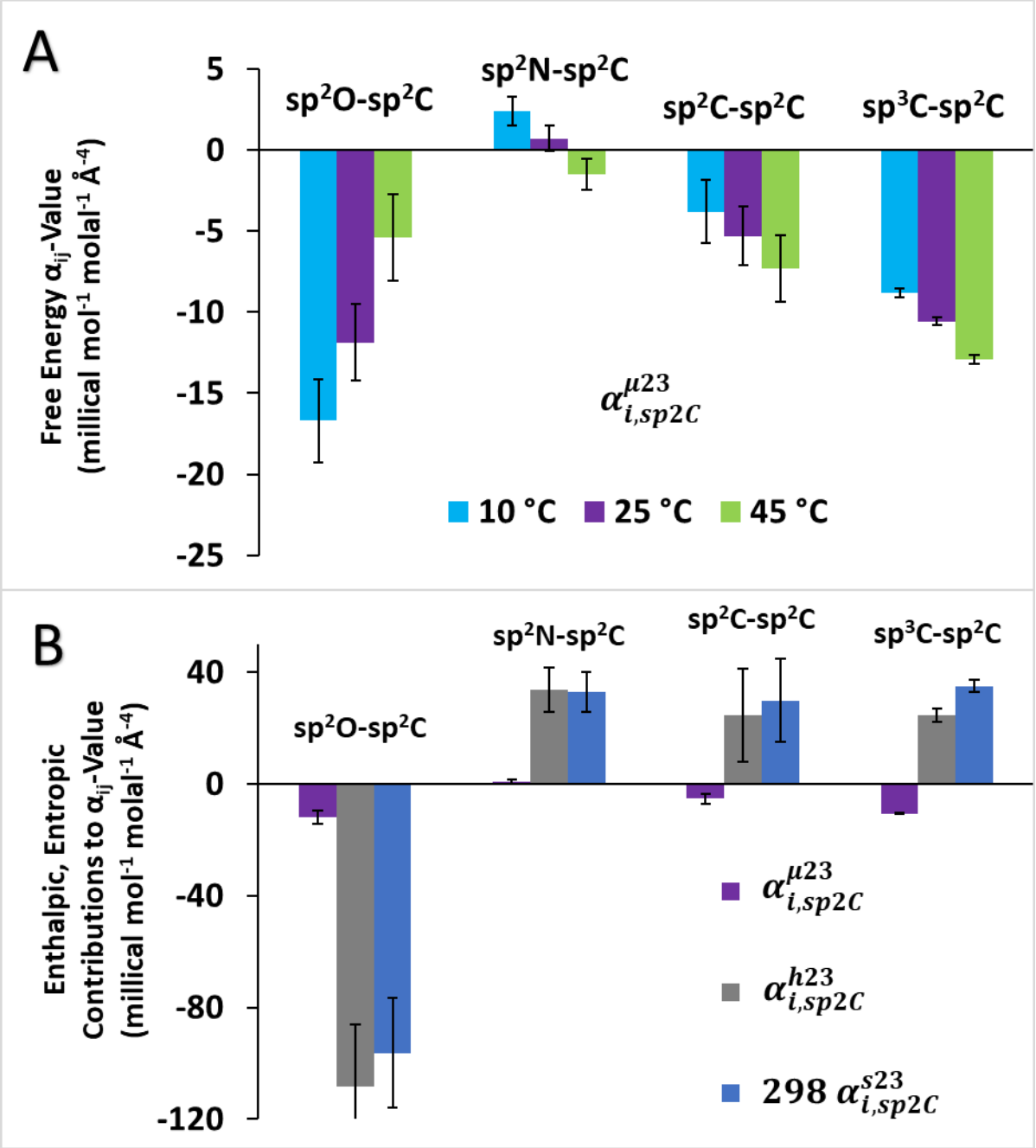



**Figure 3. Two-way α-values for amide atom-naphthalene atom interactions: Temperature dependence (panel A) and enthalpic-entropic dissection (panel B).** Panel A displays the four two-way α-values ($\alpha^{\mu 23}_{i,sp2C}$) for amide atom-naphthalene sp$^2$C interactions at 10, 25, and 45 °C. Panel B gives the dissections of these 25 °C $\alpha^{\mu 23}_{i,sp2C}$ into enthalpic ($\alpha^{h23}_{i,sp2C}$) and entropic (298 $\alpha^{s23}_{i,sp2C}$ contributions. All $\alpha_{i,sp2C}$ values and uncertainties are from Table 2.

In contrast to the amide sp$^2$O-naphthalene sp$^2$C interaction, $\alpha_{i,sp2C}^{\mu 23}$ values for interactions of amide sp$^2$C and methyl/methylene sp$^3$C with naphthalene sp$^2$C become more favorable at higher temperature, demonstrating that the entropic contribution to the these interactions is positive and hence favorable. Interactions of amide sp$^2$N with naphthalene sp$^2$C, slightly unfavorable at 10 °C, also appear to become less unfavorable with increasing temperature, though the uncertainty is large, indicating that here also the intrinsic entropic contribution is positive ($\alpha_{i,sp2C}^{s23} > 0$). At 45 °C this sp$^2$N - sp$^2$C interaction appears to be slightly favorable.

*Dissection of Interaction Enthalpies ($h_{23}$) and Entropies ($s_{23}$) Into Atom-atom Contributions*

Here we propose and test a dissection of $h_{23}$ and $s_{23}$ values for interactions of 14 amides with naphthalene in order to obtain atom-atom enthalpic and entropic information about interactions of the various unified atoms of the amide compounds investigated (sp$^2$O, sp$^2$N, sp$^2$C, sp$^3$C) with naphthalene sp$^2$C unified atoms. By analogy with the dissection of $\mu_{23}^{298\mathrm{K}}$ values (Eq. 10), we propose Eqs. 11-12 to interpret 298 $s_{23}$ and $h_{23}$.

$$298\ s_{23} = 298 \sum_i \alpha_{i,naph}^{s23} \mathrm{ASA}_i^{\mathrm{amide}} = (298 \sum_i \alpha_{i,sp2C}^{s23} \mathrm{ASA}_i^{\mathrm{amide}}) \mathrm{ASA}_{sp2C}^{\mathrm{naph}} \qquad \text{Eq. 11}$$

$$h_{23} = \sum_i \alpha_{i,naph}^{h23} \mathrm{ASA}_i^{\mathrm{amide}} = (\sum_i \alpha_{i,sp2C}^{h23} \mathrm{ASA}_i^{\mathrm{amide}}) \mathrm{ASA}_{sp2C}^{\mathrm{naph}} \qquad \text{Eq. 12}$$

Eqs. 11 and 12 provide both one-way (298 $\alpha_{i,naph}^{s23}$, $\alpha_{i,naph}^{h23}$) and two-way (298 $\alpha_{i,sp2C}^{s23}$, $\alpha_{i,sp2C}^{h23}$) entropic and enthalpic contributions to the corresponding free energy α-values ($\alpha_{i,naph}^{\mu 23}$, $\alpha_{i,sp2C}^{\mu 23}$) for each amide atom – naphthalene and amide atom – naphthalene atom interaction, where $\alpha_{i,naph}^{\mu 23}$ = $\alpha_{i,naph}^{h23}$ - 298 $\alpha_{i,naph}^{s23}$ and $\alpha_{i,sp2C}^{\mu 23}$ = $\alpha_{i,sp2C}^{h23}$ - 298 $\alpha_{i,sp2C}^{s23}$). As with Eq. 10, the assumption of additivity is tested in the analysis because the size of the experimental data sets (14 $s_{23}$, 14 $h_{23}$ values) is much greater than the number of fitting parameters (four different $\alpha_{i,sp2C}^{s23}$, four different $\alpha_{i,sp2C}^{h23}$). Enthalpic ($\alpha_{i,sp2C}^{h23}$) and entropic (298 $\alpha_{i,sp2C}^{s23}$) contributions to each 25 °C $\alpha_{i,sp2C}^{\mu 23}$

value obtained by application of Eqs. 11-12 are summarized in the bar graphs of Figure 3B and in Table 2.

Figure 3B reveals the very different thermodynamic signature of the amide $sp^2O$-aromatic $sp^2C$ interaction than those of the other atom-atom interactions. The $sp^2O$-$sp^2C$ interaction is strongly enthalpy-driven ($\alpha^{h23}_{sp2O,sp2C} << 0$), with a large opposing unfavorable entropic contribution (298 $\alpha^{s23}_{sp2O,sp2C} << 0$). This provides thermodynamic support for a direct chemical interaction of amide $sp^2O$ atoms with aromatic $sp^2C$ atoms (e.g. a lone pair-π[20-22] or n-π*[18,23-25] interaction).

On the other hand, the interaction of methyl/methylene $sp^3C$ with naphthalene $sp^2C$ is entropy-driven ($298\ \alpha^{s23}_{sp3C,sp2C} \gg 0$), with a large opposing unfavorable enthalpic contribution ($\alpha^{h23}_{sp3C,sp2C} \gg 0$). This is the thermodynamic signature of the hydrophobic effect,[26-29] and indicates that the driving force for this $sp^3C$-$sp^2C$ interaction is the favorable result of removing hydrocarbon surface from exposure to water.

Figure 3B also indicates why the free energy α-value ($\alpha^{\mu 23}_{sp2N,sp2C}$) for interaction of amide $sp^2N$ with aromatic $sp^2C$ is significantly less favorable than $\alpha^{\mu 23}_{sp3C,sp2C}$ for interaction of methyl/methylene $sp^3C$ with aromatic $sp^2C$. The entropic term is similarly favorable for both these interactions, but is completely compensated by an unfavorable enthalpic term for the amide $sp^2N$-aromatic $sp^2C$ interaction. These $sp^2N$-$sp^2C$ thermodyamics are consistent with a hydrophobic effect (entropic) from reducing the exposure of aromatic $sp^2C$ surface to water and an offsetting enthalpic cost of disrupting interactions of water with amide $sp^2N$.

The opposite signs of the entropic contributions ($\alpha^{s23}_{sp2O,sp2C} < 0;\ \alpha^{s23}_{sp3C,sp2C} > 0$)) to the intrinsic $sp^2O$-$sp^2C$ and $sp^3C$-$sp^2C$ interaction free energies ($\alpha^{\mu 23}_{sp2O,sp2C}, \alpha^{\mu 23}_{sp3C,sp2C}$) explain the opposite effects of temperature on $\mu_{23}$ of interactions of formamide and 1,3-deu with naphthalene in Figure 1. From the values of $\alpha^{\mu 23}_{i,sp2C}$ at 25 °C in Table 2 and the ASA information

in Table S1 it is straightforward to calculate that the favorable $sp^2O$-$sp^2C$ interaction makes the dominant contribution (~75%) to the formamide naphthalene $\mu_{23}$, while the favorable $sp^3C$-$sp^2C$ interaction makes the dominant contribution (~90%) to the 1,3-deu-naphthalene $\mu_{23}$ (cf. Table S7). The temperature dependences of these dominant atom-atom interactions determine the temperature dependences of the $\mu_{23}$ values.

*Enthalpy-Entropy Compensation in Atom-Atom and Solute-Solute Interactions*

Strikingly, Figure 3B reveals that all four atom-atom preferential interactions between $sp^2C$ atoms of naphthalene and $sp^3C$, $sp^2O$, $sp^2N$, and $sp^2C$ atoms of amide compounds exhibit enthalpy-entropy compensation. Enthalpy-entropy compensation in protein processes has been recently reviewed.[40] For the atom-atom interactions studied here, enthalpic ($\alpha_{i,sp2C}^{h23}$) and entropic (298 $\alpha_{i,sp2C}^{s23}$) contributions to the intrinsic interaction free energy ($\alpha_{i,sp2C}^{\mu 23}$) at 25 °C have the same sign and are both much larger in magnitude than $\alpha_{i,sp2C}^{\mu 23}$ = $\alpha_{i,sp2C}^{h23}$ - 298 $\alpha_{i,sp2C}^{s23}$. These are relatively weak favorable or neutral atom-atom preferential interactions and do not involve conventional site-binding of the amide compound to naphthalene. The observed enthalpy-entropy compensation may originate in the atom-atom interactions themselves and/or in the interactions with water that are perturbed when an amide atom interacts with a naphthalene atom. For the enthalpically-driven $sp^2O$-$sp^2C$ interaction, magnitudes of $\alpha_{i,sp2C}^{h23}$ and 298 $\alpha_{i,sp2C}^{s23}$ are about ten-fold larger than $\alpha_{i,sp2C}^{\mu 23}$. For the equally favorable, entropically-driven $sp^3C$-$sp^2C$ interaction, magnitudes of $\alpha_{i,sp2C}^{h23}$ and 298 $\alpha_{i,sp2C}^{s23}$ are four and three times larger than $\alpha_{i,sp2C}^{\mu 23}$. For the less-favorable, entropically-driven $sp^2C$-$sp^2C$ interaction, values of $\alpha_{i,sp2C}^{h23}$ and 298 $\alpha_{i,sp2C}^{s23}$ are approximately five and six times larger than $\alpha_{i,sp2C}^{\mu 23}$, with large uncertainties. For the neutral or slightly unfavorable $sp^2N$-$sp^2C$ interaction, $\alpha_{i,sp2C}^{h23}$ and 298 $\alpha_{i,sp2C}^{s23}$ are as large or larger than for $sp^3C$-$sp^2C$ and $sp^2C$-$sp^2C$ interactions and exceed $\alpha_{i,sp2C}^{\mu 23}$ by more than ten-fold.

While all amide atom–naphthalene atom interactions exhibit enthalpy-entropy compensation,Table 1 reveals that only two or three of the fourteen amide molecule – naphthalene molecule interactions investigated here exhibit it to any degree. Interactions of the two diethylureas with naphthalene have positive enthalpic ($h_{23}$) and entropic (298 $s_{23}$) contributions to the interaction free energy ($\mu_{23}^{298K}$) which are at least 1.5 times as large in magnitude as $\mu_{23}^{298K}$. The interaction of formamide with naphthalene has negative $h_{23}$ and 298 $s_{23}$ contributions to $\mu_{23}^{298K}$ which also are more than 1.5 times as large in magnitude as $\mu_{23}^{298K}$. The explanation for these behaviors is provided in Table S7, where the atom-atom $\alpha_{i,sp2C}^{h23}$ and 298 $\alpha_{i,sp2C}^{s23}$ values from Table 2 are used to predict the contributions of each atom-atom interaction to $h_{23}$ and 298 $s_{23}$. For 1,1-deu and 1,3-deu, as a result of the relatively large $sp^3C$ ASA and relatively small $sp^2O$ ASA, positive contributions to $h_{23}$ and 298 $s_{23}$ from the $sp^3C$-$sp^2C$, $sp^2C$-$sp^2C$, and $sp^2N$-$sp^2C$ interactions sufficiently outweigh the negative contributions to $h_{23}$ and 298 $s_{23}$ from the $sp^2O$-$sp^2C$ interaction to give rise to compensation. Conversely for formamide, which lacks $sp^3C$, the negative contributions to $h_{23}$ and 298 $s_{23}$ from the $sp^2O$-$sp^2C$ interaction dominate and give rise to compensation.

Thermodynamic studies of solute effects on protein stability as a function of solute concentration and temperature provide examples of enthalpy-entropy compensation, and of the lack of enthalpy-entropy compensation, at the level of solute molecule – protein molecule preferential interactions. Results for urea, glycine betaine, ethylene glycol and triethylene glycol, obtained from published solute *m*-values,[41-44] are summarized in Table S10. For unfolding, the solute *m*-value = $d\Delta G^o_{obs}/dm_{solute} = \Delta\mu_{23} = \Delta h_{23} - T\Delta s_{23}$. For solute effects that do not involve site binding, the differences in solute interactions with the unfolded protein vs the folded protein represented by $\Delta\mu_{23}$ , $\Delta h_{23}$ and $\Delta s_{23}$ are justifiably interpreted as solute interactions with the protein surface exposed in unfolding (i.e. the ΔASA).[2,45] This ΔASA of unfolding is typically about 2/3 C (mostly $sp^3C$) and 1/3 N,O, the majority of which is amide $sp^2N$ and $sp^2O$.

Table S10 reveals that interactions of urea and glycine betaine with the ΔASA of unfolding lacDBD and interactions of urea with the ΔASA of unfolding HPr do not exhibit enthalpy-entropy compensation. Interactions of ethylene glycol and oligomers of polyethylene glycol (e.g.triethylene glycol) with the ΔASA of unfolding SH3 do exhibit enthalpy-entropy compensation (Table S10).[45] Atom-atom thermodynamic information like that obtained here for interactions of naphthalene $sp^2C$ atoms with the $sp^3C$ and $sp^2O$, N and C atoms of amide compounds, is needed for interactions of the atoms of these solutes with the $sp^3C$, $sp^2N$ and $sp^2O$ atoms exposed in unfolding. If the results obtained here for interactions of the atoms of amide compounds with naphthalene atoms (Table 2) are general and the individual atom-atom interactions between the protein ASA exposed in unfolding and urea, GB, EG or PEG polyols in water also exhibit enthalpy-entropy compensation, then these effects must reinforce in the case of EG and small PEGs, but offset one another for urea and GB, like they do for the interaction of urea with naphthalene (Table 1 and Table S7) which does not exhibit enthalpy-entropy compensation.

*Comparison of Experimental Amide-Aromatic $\mu_{23}^{298K}$, $h_{23}$ and 298 $s_{23}$ with Predictions from Two-Way Atom-Atom $\alpha_{ij}$-Values*

Panels A-C of Figure 4 compare observed $\mu_{23}$-values at 45, 25 and 10 °C with those predicted from Eq. 10 using the appropriate $\alpha_{i,sp2C}^{\mu 23}$ (Tables 2 and S6) and ASA information (Table S1) for each amide. Dotted lines represent perfect agreement of predicted and observed $\mu_{23}$-values. Estimated uncertainties in both observed and predicted $\mu_{23}$-values are shown; in some cases these are the size of the data points or smaller. Figure 4A-C show that more than 70% of observed $\mu_{23}$-values at each temperature agree with predictions from Eq. 10 within the estimated uncertainty.

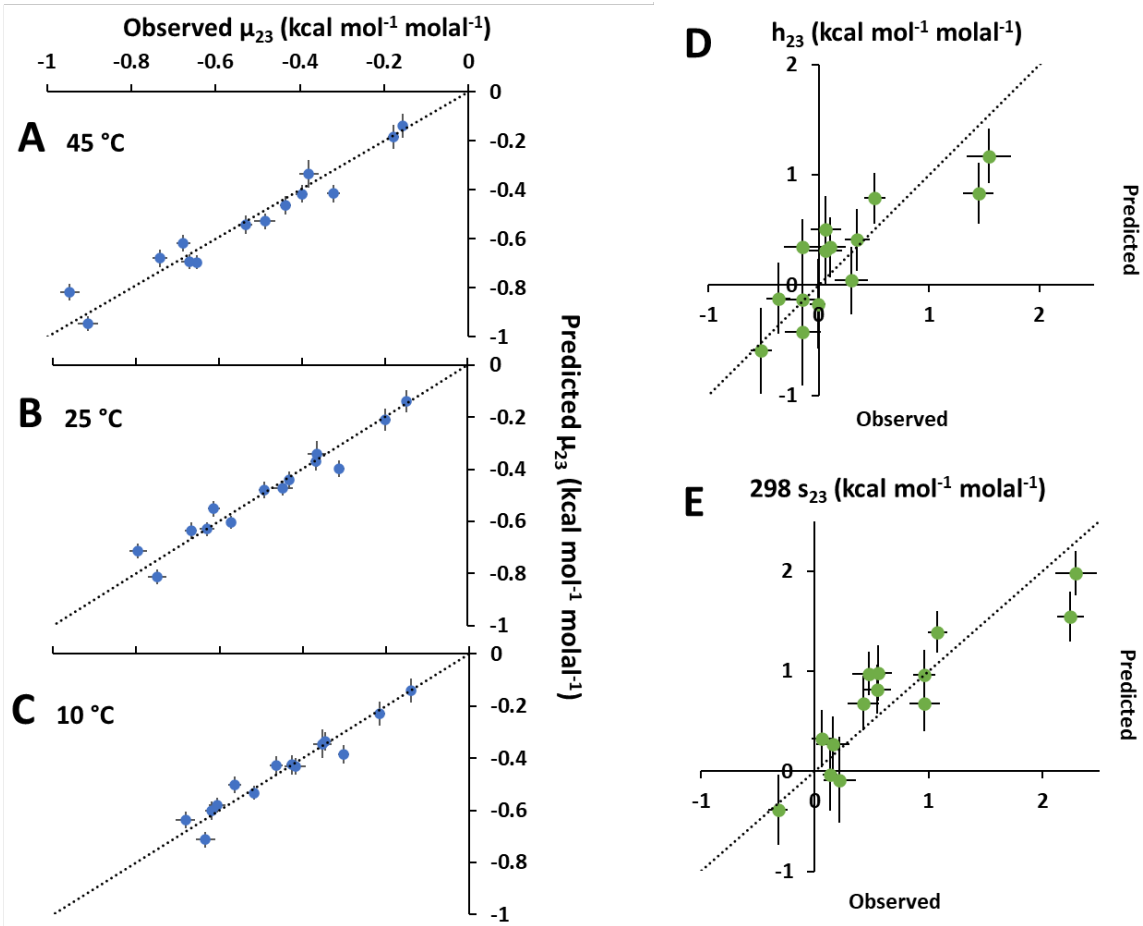



**Figure 4. Comparison of predicted and observed $\mu_{23}$ ,$h_{23}$ and 298 $s_{23}$ values.** Panels A-C compare predicted and observed $\mu_{23}$ values at 45, 25 and 10 °C. Panels D-E compare predicted and observed enthalpic ($h_{23}$) and entropic (298 $s_{23}$) contributions to 25 °C $\mu_{23}$ values. Observed $\mu_{23}$ values and uncertainties are from Table 1. Predictions of $\mu_{23}$ values at each temperature are made using Eq. 10 with the four $\alpha_{i,sp2C}^{\mu 23}$ obtained at that temperature (Tables 2, S6) and the solute ASA (Table S1). Predictions of 298 $s_{23}$ and $h_{23}$ are made using Eqs. 11 and 12, the four $\alpha_{i,sp2C}^{h23}$ and $\alpha_{i,sp2C}^{s23}$ (Table 2) and the solute ASA (Table S1). Uncertainties in all predicted values are propagated as described previously.[11]

Figure 4 D-E compare predicted and observed enthalpic ($h_{23}$) and entropic (298 $s_{23}$) contributions to $\mu_{23}^{298K}$. Observed values are reported in Tables S2 and S3, obtained as described above. Predictions are made using Eqs. 11-12, the appropriate $\alpha_{i,sp2C}^{h23}$ and $\alpha_{i,sp2C}^{s23}$ from Table S6, and ASA information (Table S1), and are listed in Tables S2-S3. As expected because of the greater uncertainties in $\alpha_{i,sp2C}^{h23}$ and $\alpha_{i,sp2C}^{s23}$ than in $\alpha_{i,sp2C}^{\mu 23}$, the predicted percentage uncertainties in $h_{23}$ and $s_{23}$ are larger than for $\mu_{23}^{298K}$. Agreement between predicted

and observed values of 298 $s_{23}$ (also $h_{23}$) within the estimated uncertainties is obtained for more than 70% of the amides investigated.

*Intrinsic Strengths of Atom – Atom Interactions of Amides with Amides and Aromatics*

The 25 °C amide-naphthalene interaction results ($\mu_{23}$ values) reported here supplement and update the set of amide-aromatic $\mu_{23}$ values determined previously.[12,14] Four of the fourteen amide compounds studied here (FAD, mACET, ppa, mad) were not investigated previously, and additional naphthalene solubility data as a function of amide concentration were obtained for the other ten amide compounds. Values of $\mu_{23}$ for these fourteen amide-naphthalene interactions were analyzed together with ten $\mu_{23}$ values for other amide-aromatic interactions (including amide-anthracene interactions) and eighty-five osmometrically-determined $\mu_{23}$ values quantifying interactions of alkylureas and other amides with one another.[14] This analysis yields an updated set of ten two-way free energy $\alpha$-values quantifying interactions of $sp^2O$, $sp^2N$, $sp^2C$ and $sp^3C$ atoms of amide and aromatic compounds with one another, shown in Figure 5A and listed in Table S9. In this analysis, as examined and justified previously, amide $sp^2C$ and aromatic $sp^2C$ atoms are analyzed together.[12,14]

All but one of the ten $\alpha_{ij}$-values calculated here (Figure 5) that quantify the different amide atom-amide/aromatic atom interactions are the same within uncertainty as those reported previously.[14] The exception is $\alpha^{\mu 23}_{sp2C,sp2C}$, quantifying the intrinsic strength of interactions of amide $sp^2C$ with amide $sp^2C$ or aromatic $sp^2C$. The value of $\alpha^{\mu 23}_{sp2C,sp2C}$ reported in Figure 5 (-5.0 $\pm$ 3.1 mcal mol$^{-1}$ molal$^{-1}$ Å$^{-4}$ ) is the same within uncertainty as that reported in Table 2 from analysis of the amide-naphthalene set of fourteen $\mu_{23}$ values (-5.3 $\pm$ 1.8 mcal mol$^{-1}$ molal$^{-1}$ Å$^{-4}$), and replaces that determined previously (-11 $\pm$ 3 mcal mol$^{-1}$ molal$^{-1}$ Å$^{-4}$).[14] These values differ by slightly more than the sum of the estimated uncertainties. Because only the three formamides have significant $sp^2C$ ASA, $\alpha^{\mu 23}_{sp2C,sp2C}$ is more difficult to determine than other $\alpha_{ij}$-values. Uncertainties in the other C-C $\alpha_{ij}$-values are much smaller and agreement with those

determined previously[14] is much better. The $\alpha_{ij}$-value for the $sp^3C$-$sp^2C$ interaction is significantly more favorable (-10.8 $\pm$ 0.4 mcal $mol^{-1}$ $molal^{-1}$ $Å^{-4}$) than for $sp^2C$-$sp^2C$, while that for $sp^3C$-$sp^3C$ interaction (-4.0 $\pm$ 0.1 mcal $mol^{-1}$ $molal^{-1}$ $Å^{-4}$) is similar in strength to $sp^2C$-$sp^2C$.

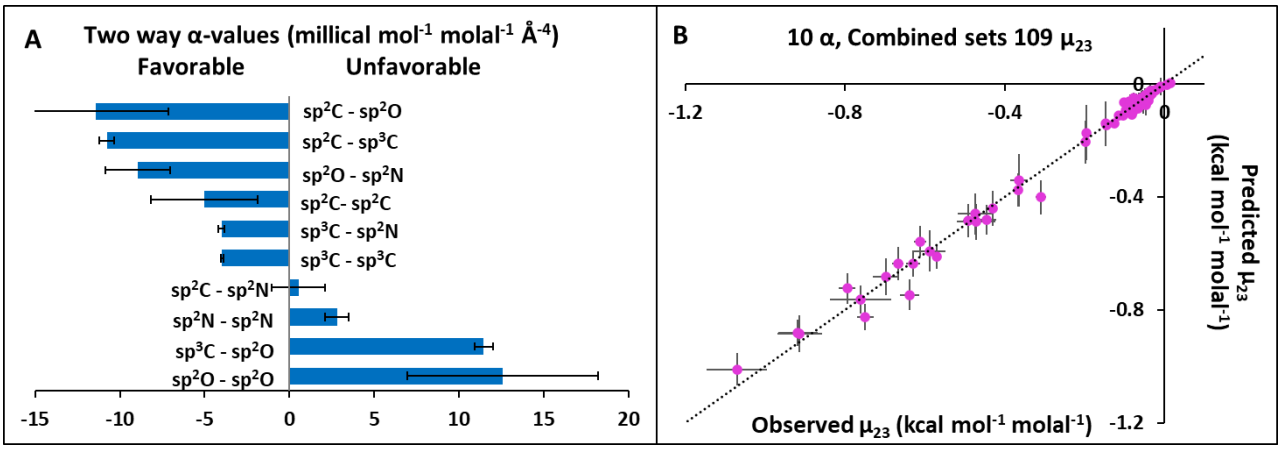



**Figure 5. Two-way $\alpha_{ij}$-values quantifying favorable and unfavorable interactions of amide unified atoms with amide and aromatic unified atoms in water.**
Panel A: Ranking of $\alpha_{ij}$-values obtained from analysis of 85 amide-amide $\mu_{23}$ values (VPO, ~23 °C) and 24 amide-aromatic $\mu_{23}$ values (solubility, 25 °C). $\alpha_{ij}$-values and uncertainties are from Table S9. Panel B: Comparison of observed $\mu_{23}$ values (listed with uncertainties in Table 1) with those predicted from two-way $\alpha_{ij}$-values (panel A) and ASA information Table S1). See Figure S8D for an expanded scale plot of the amide-amide portion of this figure. Uncertainties in predicted $\mu_{23}$ values are propagated as described previously.[11]

*Investigation of a Chemical Group Alternative (π-π, CH-π) to an Atom-Atom Analysis of Amide-Naphthalene Interactions*

A potential alternative to the above atom-atom interpretation of amide - naphthalene interactions is to propose that the amide group and the aromatic ring interact with one another and/or with methyl/methylene $sp^3C$ atoms primarily as π systems rather than as individual $sp^2C$, $sp^2O$ and $sp^2N$ atoms. We therefore asked whether the $\mu_{23}^{298K}$, $h_{23}$ and 298 $s_{23}$ results reported here for amide-naphthalene interactions could be analyzed as amide π-aromatic π and methyl/methylene $sp^3C$-aromatic π examples of π-π and CH-π interactions, respectively.

Here we test this alternative interpretation by determining the free energy, enthalpy and entropic alpha values for π-π and CH-π interactions and assessing their chemical significance

as well as their ability to predict amide-naphthalene $\mu_{23}$ values, using Eqs. S5-8. Only two $\alpha$-values are needed if the amide group is considered as an amide π system with an ASA that is the sum of the amide $sp^2O$, $sp^2N$ and $sp^2C$ ASA values in Table S1, whereas four $\alpha$-values are needed when interactions of the $sp^2O$, N and C amide atoms are analyzed individually. The two $\alpha$-value analysis would be warranted if the interaction of the amide group with the aromatic ring primarily involves the π electrons of the amide group and not the individual $sp^2O$, N and C atoms of the amide group, which contribute differently to the solvent accessibility of the entire amide group in different amide compounds.

Results of the two-parameter (aromatic π-amide π, CH-aromatic π) analysis are shown in the panels of Figure S7 and in Table S9. Panel A (Figure S7) shows the favorable amide π-aromatic π and CH-aromatic π $\alpha$-values obtained from analysis of the 14 amide-naphthalene 25 °C $\mu_{23}$ values in Table 1 using Eq. S5. Panel A also compares these amide π-aromatic π and CH-aromatic π $\alpha$-values with the atom-atom $\alpha$-values (for $sp^2O$–$sp^2C$, $sp^2N$–$sp^2C$, $sp^2C$–$sp^2C$ and $sp^3C$–$sp^2C$ interactions; Table 2) obtained from analysis of these $\mu_{23}$ values using Eq. 10. The favorable CH-aromatic π $\alpha_{ij}$-value from the two-parameter analysis is almost the same as the $sp^3C$–$sp^2C$ $\alpha$-value from the four-parameter analysis. Panel A shows that the aromatic π-amide π $\alpha$-value is in the mid-range of the individual amide atom - aromatic atom $\alpha_{ij}$-values, and much less favorable than the CH-aromatic π $\alpha_{ij}$-value.

Panel B of Figure S7 compares experimentally determined 25 °C amide-naphthalene $\mu_{23}$ values with those predicted by the two $\alpha_{ij}$-value (aromatic π-amide π, CH-aromatic π) analysis with those from the atom-atom, four $\alpha_{ij}$-value analysis. Visually it is clear from panel B that, while both analyses predict the observed $\mu_{23}$ values quite well, the atom-atom analysis is better than the aromatic π-amide π, CH-π analysis, a conclusion confirmed by the statistical analysis inTable S9. Comparison of root mean square errors (RMSE), calculated by Eq. S10 using differences between observed and predicted $\mu_{23}^{298K}$ (Table S9), reveals that RMSE is ~20%

larger for the analysis using two $\alpha$-values. Expanding the $\mu_{23}^{298K}$ data set to 24 amide-aromatic interactions using anthracene and additional naphthalene results obtained only at 298 K yields the same conclusion. These $\mu_{23}^{298K}$ are quite well-fit using two $\alpha_{ij}$-values which are the same within uncertainly as for the 14 amide set, but from both a chemical and statistical perspective the analysis using four $\alpha$-values is superior (Table S9; Figure S6 vs S8A). Likewise the analysis of $\mu_{23}$ values for amide-naphthalene interactions at 10 and 45 °C (Figure S7 panels C, D) reveals that atom-atom $\alpha_{ij}$-values provide a better interpretation both statistically and chemically.

We conclude that the aromatic π-amide π $\alpha$-values ($\alpha_{ij}^{\mu 23}$, $\alpha_{ij}^{h23}$, $\alpha_{ij}^{s23}$) are better interpreted as weighted-average approximations to the atom-atom $\alpha$-values for the different amides in the data set, rather than as π-π interactions. The ability of the aromatic π-amide π, CH-π analysis to fit the data with this average $\alpha_{ij}$-value for the range of amide groups on the 14 amide compounds investigated shows why the uncertainties in the individual atom-atom $\alpha_{ij}$-values for the amide group are relatively large (Tables 2, S9), because of residual correlations between the atom-atom $\alpha_{ij}$-values for amide $sp^2$O, N and C atoms.

Two and four $\alpha$-value analyses of 298 $s_{23}$ and $h_{23}$ values (Eq. S7-8; Figure S7 panels F,G) give similar results. While predictions of 298 $s_{23}$ and $h_{23}$ using only two $\alpha$-values (π-π, CH-π, Table S8) are almost as good as those obtained using four atom-atom $\alpha$-values (Table S6), in all cases the quality of the fit as assessed by RMSE is better using the four atom-atom $\alpha$-values.

In summary, although the two $\alpha$-value analysis of these data as π interactions described in this section is moderately successful from a statistical perspective, the four parameter analysis as atom-atom interactions is in all cases ($\alpha_{i,sp2C}^{\mu 23}$ at 10, 25 and 45 °C; $\alpha_{i,sp2C}^{s23}$, $\alpha_{i,sp2C}^{h23}$) a better statistical fit. Also, the thermodynamics that we determine from both analyses for the very favorable interaction of methyl and methylene portions of the amide compound with

naphthalene (entropy driven, enthalpically unfavorable) are more consistent with this being a hydrophobic effect, driven by the entropic effect of releasing local water from these hydrocarbon surfaces when they interact, than what would be expected for a direct CH-π interaction.

We also tested whether the combined 109 member set of amide-amide and amide-aromatic $\mu_{23}$ values could be successfully interpreted if interactions with the amide group are analyzed as interactions with the amide π system rather than as interactions with amide $sp^2O$, $sp^2N$ and $sp^2C$ atoms, using Eqs. S10A and B. Table S9 summarizes the α-values obtained from this analysis and compares RMSE values for 4 and 10-parameter analyses. Figures S8 compares predicted and observed $\mu_{23}^{298K}$ values using these different fitting analyses. Figure S8B shows the full range of the 109 $\mu_{23}^{298K}$ data set, which is dominated by the amide-aromatic portion, while Figure S8C focus on the amide-amide portion of that large $\mu_{23}^{298K}$ data set. The four parameter analysis predicts a subset of the $\mu_{23}^{298K}$ values quite well, as shown in Figure S8B, but statistical analysis (RMSE; Table S9) reveals that it is not as good as the fit involving ten $\alpha$-values. More importantly, $\alpha$-values for the combined amide group lack the molecular information provided by individual amide atom $\alpha$-values, which make good chemical sense as discussed above and previously.[12,14]

In summary, while the reduced parameter fits do quite well, the atom-atom fits are better. Where $\alpha$-values in the reduced parameter fits have an atom-atom interpretation, like CH-aromatic π as $sp^3C$-$sp^2C$, these $\alpha$-values are very similar to the atom-atom $\alpha$-values. The amide π $\alpha$-values behave like rough averages of the atom-atom alpha values for amide $sp^2O$, $sp^2N$ and $sp^2C$ and provides no support for the idea that amide interations are primarily π system interactions as opposed to individual atom-atom interactions. Atom-atom $\alpha$-values provide a better analysis of all these amide interaction data, both statistically and chemically.

## Conclusions

We conclude from measurements of effects of 14 amide compounds on solubility of naphthalene at 10, 25 and 45 °C that the thermodynamics of amide compound - naphthalene interactions ($\mu_{23}$, $h_{23}$, $s_{23}$) are well-analyzed by the ASA-based approach previously applied to $\mu_{23}$ values for amide interactions at 25 °C.[14] This analysis makes and tests the assumption of additivity of the different atom-atom interactions to molecule-molecule interactions ($\mu_{23}$, $h_{23}$, $s_{23}$). In this analysis, the intrinsic contribution of each atom-atom interaction (expressed per unit of accessible surface area of each atom) is weighted by the product of ASA of these interacting atoms on the two molecules. This use of surface area as the fundamental variable is justified in previous SPM analyses of solute-protein interactions[46] and solute and salt effects on surface tension[47-49]. For amide-amide and amide-aromatic interactions, use of ASA as the fundamental variable was shown to be superior to an analysis based only on the number of each type of atom and not its accessibility.[12,14]

From the temperature dependences of $\mu_{23}$ values for these 14 amide compounds with naphthalene, we dissect $\mu_{23}$ values into their enthalpic ($h_{23}$) and entropic (298 $s_{23}$) contributions at 25 °C. For the series of alkylureas and formamides, $\mu_{23}$, $h_{23}$ and 298 $s_{23}$ all vary systematically with amide composition. Analysis of these results yields enthalpic ($\alpha_{ij}^{h23}$) and entropic (298 $\alpha_{ij}^{s23}$) contributions to intrinsic strengths of interactions ($\alpha_{ij}^{\mu 23}$) of amide $sp^2O$, $sp^2N$, $sp^2C$ and $sp^3C$ unified atoms with aromatic $sp^2C$ atoms. These enthalpic and entropic contributions have the same sign and are much larger in magnitude than the interaction free energy itself. The amide $sp^2O$-aromatic $sp^2C$ interaction is enthalpy-driven and entropically unfavorable, consistent with direct chemical interaction (e.g. lone pair-π). On the other hand, amide $sp^3C$- and $sp^2C$-aromatic $sp^2C$ interactions are entropy-driven and enthalpically unfavorable, consistent with hydrophobic effects, and the amide $sp^2N$-aromatic $sp^2C$ interaction

is entropically favorable, as if hydrophobic, with an unfavorable enthalpic term of equal magnitude so the net free energy of this interaction is zero within uncertainty. These findings will be relevant to interpret interactions involving π-bonded $sp^2$ atoms in protein processes.

## Methods

### *Experimental Methods*

Solubility assays are performed and analyzed as previously described.[10-12] Additional details of these assays, and information about the naphthalene and amide samples used in this research, are given in SI Methods.

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