Hybrid Synthesis of bioplastics polyhydroxybutyrate from carbon dioxide

Jie Zhang¹,²,³,#, Dingyu Liu²,³,⁴#, Yuwan Liu²,³,⁴, Huanyu Chu²,³, Jian Cheng²,³, Haodong Zhao¹,²,³, Shaoping Fu²,³, Huihong, Liu⁵, YuE Fu⁵, Yanhe Ma²,³*, Huifneg Jiang²,³*

Affiliations:

¹School of Life Sciences, Division of Life Sciences and Medicine, University of Science and Technology of China, Hefei 230027, China.
²Tianjin Institute of Industrial Biotechnology, Chinese Academy of Sciences, Tianjin 300308, China.
³National Center of Technology Innovation for Synthetic Biology, Tianjin, 300308, China.
⁴Haihe Laboratory of Synthetic Biology, Tianjin 300308, China.
⁵China BlueChemical Ltd., Beijing 100029, China.

#These authors contributed equally to this work.
*Corresponding author. Email: Huifeng Jiang, jiang_hf@tib.cas.cn; Yanhe Ma, ma_yh@tib.cas.cn.
Abstract

The accelerating environmental crisis has intensified the demand for switching from traditional economy to a renewable one with a reduced carbon footprint. Here we reported a hybrid system, coupling chemical process of CO₂ hydrogen reduction and biological process for polyhydroxybutyrate (PHB) synthesis, that utilized CO₂ as a raw material to produce PHB in vitro. The synthetic pathway of PHB was optimized by screening more efficient methanol oxidases, high activity mutants of glycolaldehyde synthase and coordinating enzyme dosages in the pathway, which achieved the carbon yield of 93.6% for producing PHB from methanol. Finally, by combining with the chemical process from CO₂ to methanol, a scaling-up bio-system was performed to convert CO₂ into PHB, yielding 5.8 g/L with the productivity of 1.06 g/L⋅h⁻¹. This approach represents a promising carbon-neutral way to produce biodegradable plastics.
Introduction

The atmospheric CO₂ concentration are increasing at an alarming rate, resulting in dire climate change effects. The use of CO₂ as feedstock delivers environmentally sustainable, carbon-negative manufacturing of chemicals and materials.[1,2]. Polyhydroxybutyrate (PHB) is a class of biodegradable plastics, a well-recognized renewable and petroleum-based plastics substitute. Over the last 30 years, advances in metabolic engineering have enabled construction of microbial cell factories and cell-free systems for PHB production from sugar-based feedstock[3]. Although great success has been achieved, large-scale production and commercial applications remain limited due to manufacturing expense, approximately 50% of which are generated from raw materials[4]. Therefore, the utilization of CO₂ as a source material in the synthesis of biodegradable plastics PHB offers a win-win strategy to both decrease the CO₂ emissions and solve plastic pollution.

Photosynthetic carbon fixation is the main natural CO₂ fixation process[5]. However, the titers and productivities of PHB in the engineered cyanobacteria were relatively low to unable economic feasibility and suffers from very low solar energy efficiency and slow growth rate[6, 7]. The Cupriavidus necator is also a very attractive autotroph platform with half a century of research history, and it can produce PHB from CO₂, O₂ and H₂ at a rate of up to 1.55 g/L/h[8]. However, the large-scale production of PHB by mixed-gas fermentation presents a serious explosive risk. Despite significant efforts in chemistry, biology, materials, and engineering to develop CO₂ fixation and utilization, comprehensive solutions remain to be explored.

Here we designed a hybrid system, coupling photovoltaic hydrogen production, CO₂ hydrogenation with chemoenzymatic PHB synthesis, that drove an artificial environmentally friendly carbon-negative process for carbon fixation into PHB. A space-coupled system with excellent carbon fixation rate and molar conversion efficiency were obtained. In a chemoenzymatic process, carbon molar yield of 93.6% was achieved by a two-step pathway for producing PHB from methanol. Finally, a scaling-up hybrid system was performed to convert CO₂ into PHB, yielding 5.8 g/L.
with the productivity of 1.06 g\textsuperscript{-1}L\textsuperscript{-1}h\textsuperscript{-1}.

**Results**

**Construction of hybrid system for PHB synthesis from CO\textsubscript{2}**

Combining the captured CO\textsubscript{2} with hydrogen produced from renewable energy to synthesize green methanol\textsuperscript{[9]}, has great potential to provide precursors for one carbon bio-manufacturing. The Synthetic Acetyl-CoA (SACA) pathway, a carbon-conserved and ATP-independent acetyl-CoA synthesis pathway, was applied to convert one carbon to acetyl-CoA\textsuperscript{[10]}. In fact, the carbon yield of PHB production is determined by the yield of acetyl-CoA in the synthetic pathway, which usually was limited by inherent carbon loss during the natural CO\textsubscript{2} assimilation pathway. Integrating with the synthetic pathway from acetyl-CoA to PHB, a powerful driving system for energy conversion and material metabolism is promised to produce PHB from CO\textsubscript{2} with high efficient and carbon yield (Fig. 1). The technology of CO\textsubscript{2} hydrogenation was employed to drive CO\textsubscript{2} fixation to provide one-carbon (C1) unit\textsuperscript{[11]}. Methanol was served as intermediate to bridge chemocatalysis and biosynthesis in this work. We coupled methanol oxidation reaction with the SACA pathway to convert C1 unit into acetyl-CoA\textsuperscript{[10]}, which is then converted into PHB by the synthesis pathway of PHB\textsuperscript{[12]}. All reactions and the standard Gibbs free energy change ($\Delta G'$) of each enzymatic reaction are listed in Table S2. This pathway efficiently converted CO\textsubscript{2} to PHB with theoretical carbon yield of 100%, which displayed a significant atom-economic advantage compared to the natural metabolic pathway.

To prove this design, we used 20 mM methanol as substrate to synthesize PHB. However, one-pot enzymatic synthesis enabled inefficient production of PHB, only representing approximately 10% of the theoretical molar conversion efficiency of carbon (Supplementary Fig. 1). Acetate produced by the spontaneous hydrolysis of acetyl-phosphate was the main by-product in the pathway\textsuperscript{[13]}, since the carbon flux was kinetically trapped at acetyl-phosphate due to subsequent thermodynamically unfavourable reaction. At the same time, NADPH, the downstream driving force of
thermodynamically unfavourable reaction, was also oxidized by catalase (CAT)[14]
and methanol oxidase (AOX) (Supplementary Fig. 2). Thus, the acetyl-phosphate
preferred to be converted into acetate. In addition, formaldehyde is toxic for many
enzymes[15]. In order to run well the desired enzymatic cascade, we proposed to
redivide the pathway into three modules, the module I containing all chemical
reactions from CO₂ to methanol; the module II including biological reactions from
methanol to glycolaldehyde and the module III involving enzymic cascade from
glycolaldehyde to PHB.

**Dynamic modulation of module II**

Methanol was converted into glycolaldehyde via condensation of formaldehyde,
an enzymatic cascade reaction containing AOX, CAT and glycolaldehyde synthase
(GALS) (Fig.2a). In exception to the physiological substrate methanol, the AOX can
also oxidize formaldehyde into formate due to the hydration of formaldehyde in
aqueous solution[16]. Considering the principle of atomic economy, we proposed to
employ formate to regenerate NADPH by formate dehydrogenase (FDH) for
module III. In fact, if the ratio of formaldehyde/formate is 4:1, the generated NADPH
is perfect for the synthesis of PHB. Thus, we screened AOX genes with significant
formaldehyde specificity based on product distribution in the case of methanol as a
substrate (Fig.2b). The PcAOX (from *Phanerochaete chrysosporium*) was finally
selected for glycolaldehyde synthesis module, which was capable of balancing carbon
fluxes of NADPH regeneration and PHB synthesis via formaldehyde[17]. The
percentages of formaldehyde and formate were 84% and 16%, which almost achieved
the theoretical stoichiometry for converting methanol to PHB (Fig. 2b). Finally, the
glycolaldehyde synthesis module enabled 8 mM glycolaldehyde and 3.5 mM formate
production from 20 mM methanol in 1.5 h, containing 0.2 g/L PcAOX, 10 g/L GALS
and 300 U/mL CAT.

The GALS is a key element in the hybrid system, however it accounted for 90%
of the total protein dosage in the module II. In order to cut down the concentration of
total protein, we determined to improve the catalytic activity of GALS. We proposed
to screen significant active site residues around the active center, where 14 previously
identified positions were selected to do single-point saturation mutagenesis[10]. The
sub-saturating concentrations of formaldehyde (30 mM) was employed to screen for
GALS variants with higher affinity for substrate and improved activity for
glycolaldehyde production. After screening, libraries of positions N27,E28,F397 and
C398 contained more variants with significantly increased activities (Supplementary
Fig. 3). Subsequently, we introduced a four-site random combination of mutations
into GALS and selected the highest active mutant. After screening of more than 5000
clones the beneficial combinations F397Y and C398M were identified and the double
mutant (GALS\textsubscript{F397YC398M}) enabled the substrate affinity and catalytic efficiency to
reach 117 mM and 146 M\textsuperscript{-1}\textsuperscript{S}\textsuperscript{-1}, respectively, where the kcat of variant was improved
approximately 10.8-fold than GALS (Fig. 2c). The mutations of F397Y and C398M
are all located in the monomer-monomer interface, and the substitutions shrinked the
volume of substrate binding pocket and may enhanced the interaction between ThDP
and substrate (Fig. 2d). Finally, 2 g/L GALS\textsubscript{F397YC398M} produced the similar yield of
glycolaldehyde with 10 g/L GALS (Fig. 2e). The total protein dosage was reduced by
approximately 5-fold, which was superior for industrial scale-up and
commercialization.

**Precise optimization of module III**

Subsequently, the module III from glycolaldehyde to PHB was optimized
containing six enzymes: acetyl-phosphate synthase (ACPS), phosphate
acetyltransferase (PTA), Acetyl-CoA acetyltransferase (PhaA), Acetoacetyl-CoA
reductase (PhaB), PHB synthase (PhaC) and formate dehydrogenase (FDH) (Fig. 3a).
In order to eliminate by-product acetate, we balanced the reaction rates by fine-tuning
the ratio of individual enzyme dosages to prevent the accumulation of intermediates.
The increases in all enzyme loadings by 2-fold did not significantly enhance PHB
molar yield of carbon, indicating that these enzyme loadings was sufficient
To avoid the accumulation of acetyl-phosphate, the dosages of ACPS and PTA were gradually regulated to balance the flux ratio of acetyl-phosphate synthesis and consumption. The enzyme dosages of the PHB synthesis pathway were also adjusted to pull fluxes into the downstream reactions. By rerouting the carbon metabolic flux, enzymatic reaction rates in the cascade reached an equilibrium with only slight amounts of acetate (Fig. 3b).

Finally, 20 mM methanol was used as a substrate for PHB synthesis by combing module II and III. Firstly, the methanol was converted into 8.46 mM glycolaldehyde and 3.5 mM formate in 1.5 hour. After removing enzymes from this system, approximately 4 mM of PHB was accumulated in the subsequent 5 hours by supplementing the remaining six enzymes and auxiliary components (Fig. 3c). This system achieved a carbon molar yield of approximately 93.6%, exceeding all PHB biosynthetic pathways reported to date [18, 19].

**PHB synthesis via hybrid system from CO2**

Based on the above optimization, we attempted to *de novo* synthesize PHB from CO2 and hydrogen by coupling the enzymatic processes with CO2 reduction (Fig. 4a). In fact, as increasing substrate concentration, module III was capable of maintaining high carbon molar yield up to 300 mM glycolaldehyde (Supplementary Fig. 5a). Unfortunately, the yield of glycolaldehyde began to decrease when over 20 mM methanol was added in module II (Supplementary Fig. 5b). It should be ascribed to the specificity of the AOX enzyme for glycolaldehyde and insufficient dissolved oxygen in the system. To further test the scale up potential of this system, scaling-up technology of one-pot concentration was introduced in subsequent projects.

Firstly, the chemical reaction unit was operated at 250°C, 5 MPa and CO2 was chemically hydrogenated to methanol at a rate of 18 mM h⁻¹ g⁻¹ Cu-based catalyst with molar conversion of carbon of 85.2% and hydrogen of 84.8% (Fig. 4b)[20]. The produced methanol was constantly condensed and fed into the enzymatic unit to a final concentration of 20 mM. In the enzymatic reaction unit, a 1 L volume of
glycolaldehyde synthesis module was carried out, yielding approximately the equivalent glycolaldehyde titer and efficiency as the micro-scale system system in 1.5 hours. After stopping the reaction and removing enzyme preparation, the reaction solution was concentrated by vacuum enrichment at ambient temperature and pressure. Finally, the volume of 1 L is concentrated to 58.8 ml and a 5% loss of target product was detected in the physical treatment process (Fig.4c).

The processed solution was added as a substrate into subsequent module as well as supplementing the remaining enzymes and cofactors. The last enzymatic reaction produced 5.8 g/L PHB in total in the subsequent 5 hours, corresponding to a slightly dropped carbon yield (85%) (Fig.4d). Compared to other in vitro biosynthetic systems for PHB production, this artificial hybrid system features higher efficiencies in enzymatic reaction. The excellent molar conversion efficiency of carbon was benefited from design of atom-economic pathway and efficient engineered enzymes. By using spatial and temporal segregation of steps, the hybrid system achieved a PHB productivity of 1.06 g L⁻¹h⁻¹ from CO₂ with 73.6% of molar utilization efficiency of CO₂, exceeding that of other PHB synthesis biosystems from CO₂.

Discussion
Key challenges of synthesis of PHB based on CO₂ are powerful energy efficiency and efficient carbon-conserving pathway for carbon-negative biosynthesis. The hybrid system is the most appealing strategy due to its high energy efficiency and stability[21]. Recently, chemoautotrophic photo-electrosynthesis have been gaining prominence for fixing CO₂ and some typical electro-biological models are reported. A hybrid inorganic-biological system in combination with the Ralstonia eutropha drove CO₂ direct fixation into PHB, resulting in a titer of 700 mg/l and efficiency (ηₑlec) of 36%[22]. Another integrated two-step process was developed, coupling the hybrid CO₂ electrosynthesis with acetate fermentation by Sporomusa ovata and Cupriavidus basilensis, which converted CO₂ to PHB with 11.06% of overall carbon molar yield[23]. The solar energy conversion rate of the chemoautotrophic
photo-electrosynthesis system is about 7-8%, which is a significant increase compared to the photosynthetic system[24]. However, commercial scale is still not available, as electron transfer mechanism and electrode-biological interface remain issues to be resolved. Moreover, the CO₂ assimilation pathway in autotrophic systems were mainly dependent on the CBB cycle and Wood-Ljungdahl pathway, so the carbon yield of PHB was limited by CO₂ emission via natural carbon metabolisms in host strain.

Over the recent years, CO₂ hydrogenation to methanol is one of the attractive and potentially profitable routes in CCSU (Carbon Capture, Storage, and Utilization)[25]. Bio-manufacturing based on artificial atomic economic pathway using green methanol as a substrate offers carbon-negative route to synthesize target products in environmentally friendly manner. Very recently, this approach has been demonstrated for starch synthesis from directly CO₂ via 11 core reactions.[26]. In this work, we described a bio-hybrid system, coupling energy capture with carbon fixation and conversion, that efficiently converts CO₂ to PHB, resulting in 5.8 g/L with 73.6% of CO₂ utilization yield and maximum 93.6% of carbon molar yield in biological process. This hybrid system displayed much higher carbon yield and conversion rate than the chemoenzymic system for starch. We believe that the integration of photovoltaic solar-harvesters, chemically reduced CO₂ and enzymatic biocatalysts would representing a promising carbon-neutral way to fundamentally solve the CO₂ utilization.

Plastic pollution worldwide has raised the demand for biodegradable bioplastics[27, 28]. PHAs is considered as a promising alternative to traditional chemical plastics because it rapidly and completely degrades in environment[29]. However, although PHAs have been extensively studied for 30 years, their commercialization is limited due to high cost of the production procedure[30]. The PHA produced by current microbiological fermentation has resulted in a high price of US$4-6/kg, 5-6 fold that of petroleum-based plastics[30]. Reducing the cost of raw materials is the essential key to overcome this issue. In our hybrid system, the CO₂ consumption for obtaining 1 kg of methanol is 1.4 kg and the hydrogen consumption
is 0.2 kg. The integrated costs of 1 kg of methanol is approximately US$0.37, as measured by the photovoltaic power guide price and CO₂ capture costs in CCUS (Carbon Capture, Storage, and Utilization). The raw material cost of 1 kg of PHB is approximately US$0.67, which is only 1/6-1/10 of the current price. This price is expected to be continuously reduced because photovoltaic electricity prices will continue to plummet with cost decrease in renewable technologies and local policy supports.

Current global CO₂ emissions are already 37 billion tonnes per year and global atmospheric CO₂ concentration is expected to reach 500 ppm by 2045[31]. The exacerbating climate crisis has accelerated the demand for carbon-negative chemical manufacturing and renewable economic models[32]. A carbon-negative manufacturing for PHB allows the establishment of a complete closed-loop PHB production from CO₂ and its degradation to CO₂. We envisioned a practical prospect of 3G bio-manufacturing platform that artificial hybrid system drove efficient carbon-neutral manufacturing. In the future, the key will be to keep enzyme costs low by finding stable enzymes so that they can be used for long periods of time, creating methods for recycling the enzymes, and developing inexpensive purification methods. These are largely technical rather than fundamental challenges. We therefore propose further development of hybrid system for PHB synthesis from CO₂.

**Conclusion**

In this study, we developed an artificial hybrid system to produce bio-degradable plastic PHB directly from CO₂. A space-coupled hybrid system with excellent carbon fixation rate and molar conversion efficiency was constructed. In the hybrid process, carbon molar yield of 93.6% was achieved for producing PHB from methanol. Overall, this study provides a feasible alternative solution to address plastic pollution and excessive CO₂ emissions simultaneously.
References


Acknowledgments: We thank the core facility center at Tianjin Institution of Industrial Biotechnology, CAS, for instrument and technology support.

Funding:

- National Key R&D Program of China Grant 2021YFC2103500 (YWL)
- National Natural Science Foundation of China NSFC-32001030
- Strategic Priority Research Program of the Chinese Academy of Sciences-Precision Seed Design and Breeding grant XDA24020103-3 (HFJ)
- Tianjin Synthetic Biotechnology Innovation Capacity Improvement Project Grant TSBICIP-KJGG-007 (HFJ)

Author contributions:

- Conceptualization: HFJ, YHM, JZ, DYL
- Methodology: HFJ, JZ, DYL
- Investigation: JZ, DYL, YWL, HYC, JC, HDZ, SPF
- Visualization: HFJ, JZ, DYL
- Funding acquisition: YWL, DYL, HFJ, YHM
- Project administration: HFJ
- Supervision: HFJ, YHM
- Writing – original draft: DYL, JZ, HFJ

Competing interests: There is no competing financial interest

Data and materials availability:

All data are available in the main text or the supplementary materials.
Supplementary Materials

Materials and Methods

Figs. S1 to S5

Tables S1 to S4

References 1-5
**Fig. 1.** Design of hybrid system for producing PHB from CO$_2$. Inner circle: A complete closed-loop carbon cycle for PHB synthesis to degradation. The SACA pathway indicates Synthetic Acetyl-CoA pathway. Green labelled H$_2$O and H$_2$ indicates the process of hydrogen production by water electrolysis. Outer circle: schematic of hybrid system, coupling photovoltaic hydrogen production, CO$_2$ hydrogenation with chemoenzymatic PHB synthesis and environmental degradation. The individual modules were colored. Auxiliary enzymes and chemicals are indicated.
**Fig. 2. Dynamic modulation of interface from methanol to glycolaldehyde.** (A) Schematic of module II from CO\(_2\) to glycolaldehyde and formate, with individual modules colored. All enzymes and chemicals are indicated. Red asterisk indicates key elements to be improved. (B) Screen of methanol oxidase (AOX). The reaction mixture (0.2 mL) was performed under the condition of 20 mM methanol, at 37°C, 1.5 h. (C) Enzyme kinetics of GALS and mutant. The main figure is enzyme kinetics of GALS and mutant. Enlarged view show enzyme kinetics of GALS more accurately. Enzyme kinetics were determined with 0.05 mg mL\(^{-1}\) enzyme. The concentration of formaldehyde ranged from 0 to 160 mM. (D) The overview of the selected two mutations in active center. The yellow dotted lines indicate distances between the ThaP and two mutations. (E) Time profiles of glycolaldehyde and formate production under the condition of 10 g/L GALS or 2 g/L GALS variant. The reaction mixture (0.5 mL) was performed under the condition of 20 mM methanol, at 37°C, 1.5 h. All values shown are means of triplicate measurements. The error bars represent standard deviations.
Fig. 3. Optimization of pathway from glycolaldehyde to PHB. (A) Schematic of module II from CO₂ to glycolaldehyde and formate, with individual modules colored. All enzymes and chemicals are indicated. The grey indicates products from module II. The red cross indicates deletion of by-product pathway. (B) Optimization of enzyme loadings for balancing the flux ratio of pathway. The forms described main enzyme loadings in this pathway. The reaction mixture (0.5 mL) was performed under the condition of 20 mM glycolaldehyde, at 37 °C, 5 h. (C) Demonstration of the integrated pathway from methanol to PHB. The reaction mixture (0.5 mL) was performed under the condition of 20 mM methanol as initiated substrate, at 37 °C, 5 h. All values shown are means of triplicate measurements. The error bars represent standard deviations.
Fig. 4. PHB synthesis via hybrid system from CO$_2$. (A) The schematic procedure of PHB synthesis from CO$_2$ and hydrogen. Yield of module I indicates molar conversion efficiency of CO$_2$. Yield of module II indicates carbon yield of enzymatic reaction and loss efficiency of product in the physical enrichment process. Yield of module III indicates carbon yield of last enzymatic reaction. (B) Time profiles of methanol production. The reaction was operated at 250°C, 5 MPa. The airspeed was 4000 h$^{-1}$. (C) Effect of physical enrichment process. The reaction solution was collected centrally and then was concentrated by Vacuum Concentrator at ambient temperature and pressure. (D) Time profiles of PHB production and glycolaldehyde and formate consumption in the last enzymatic process. The reaction mixture (1 mL) was performed under the condition of processed solution, at 37°C, 6 h. All values shown are means of triplicate measurements. The error bars represent standard deviations.